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**SAMPLE PAPER-03 (solved)**  
**CHEMISTRY (Theory)**  
**Class - XI**

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Time allowed: 3 hours

Maximum Marks: 70

**General Instructions:**

- a) All the questions are compulsory.
- b) There are **26** questions in total.
- c) Questions **1** to **5** are very short answer type questions and carry **one** mark each.
- d) Questions **6** to **10** carry **two** marks each.
- e) Questions **11** to **22** carry **three** marks each.
- f) Questions **23** is value based question carrying **four** marks.
- g) Questions **24** to **26** carry **five** marks each.
- h) There is no overall choice. However, an internal choice has been provided in one question of two marks, one question of three marks and all three questions in five marks each. You have to attempt only one of the choices in such questions.
- i) Use of calculators is **not** permitted. However, you may use log tables if necessary.

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1. Why the molar enthalpy of vapourisation of acetone is less than that of water?
  2. How many electrons are present in 16g of methane?
  3. How many grams of  $\text{Na}_2\text{CO}_3$  should be dissolved to make 100 cc of 0.15M  $\text{Na}_2\text{CO}_3$ ?
  4. What will be the volume of ammonia formed if 2L of nitrogen is mixed with 2L of hydrogen at constant temperature and pressure?
  5. Give two examples of state functions.
  6.
    - a) Is it possible to achieve equilibrium between water and its vapour in an open vessel?
    - b) Explain your answer and say what happens eventually.
  7. How does electronegativity vary i) down the group and ii) across the period?
  8. The standard solution of NaOH cannot be prepared by weighing. Why?
  9. Give reason: "Although geometries of ammonia and water molecules are distorted tetrahedral, bond angle in water is less than that of ammonia".

Or

Why is benzene extra-ordinary stable though it contains three double bonds?

10. In the estimation of sulphur by Carius method, 0.468 g of an organic sulphur compound afforded 0.668 g of barium sulphate. Find out the percentage of sulphur in the given compound.
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11. What happens when alkali metals reacts in air?
12. Give the Lewis representation of:
- Nitric acid
  - Ammonia
  - Ozone molecule
13. In the reaction:  $2\text{SO}_2(\text{g}) + \text{O}_2 \rightleftharpoons 2\text{SO}_3(\text{g}) + 189.4 \text{ kJ}$
- Indicate the direction in which the equilibrium will shift when:
- Concentration of  $\text{SO}_2$  is increased.
  - Concentration of  $\text{SO}_3$  is increased.
  - Temperature is increased.
14. Calculate the wavelength in nm, of visible light having a frequency of  $4.37 \times 10^{14}/\text{s}$ .
15. Comment on each of the following observations:
- Lithium forms a nitride directly like magnesium. Give equation involved.
  - $\text{BaO}$  is soluble but  $\text{BaSO}_4$  is insoluble in water.
16. Give reasons: "The reaction  $2 \text{Na}(\text{s}) + \text{H}_2(\text{g}) \rightarrow 2 \text{NaH}(\text{s})$  is a redox change".
17. Give reason:
- Graphite is used as lubricant.
  - Diamond is used as an abrasive.
  - Aluminium alloys are used to make aircraft body.

Or

Explain isomerization in alkanes with examples.

18. An alkyl halide compound 'A' ( $\text{C}_5\text{H}_{11}\text{Br}$ ) reacts with ethanolic  $\text{KOH}$  to give compound 'B', an alkene. 'B' on reaction with bromine gives compound 'C'. 'C' on further dehydrobromination gives compound 'D'. When one mole of 'D' is treated with sodium metal in liquid ammonia, it gives one mole of sodium salt of 'D' and half a mole of hydrogen gas. On complete hydrogenation, 'D' gives a straight reactions involved.
19. Calculate the amount of ammonia formed when 50 kg of  $\text{N}_2(\text{g})$  and 10.0 kg  $\text{H}_2(\text{g})$  of are mixed to produce  $\text{NH}_3(\text{g})$ . Identify the limiting reagent.
20. Give reasons:
- Evaporation causes cooling.
  - Falling liquids drops are spherical.
  - Vapour pressure of acetone is less than that of ether at same temperature.
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21. Give a note on:
- Mist
  - Smoke
  - Fumes
  - Dust
22. Calculate the enthalpy change when 2.38g of CO vaporizes at its normal boiling point, if the enthalpy of vaporization of CO is 6.04 kJ/mol.
23. John takes snacks every day to school, but Mala takes vegetables, chapattis and curd. Chips and snacks packet are filled with nitrogen gas. If they are filled with oxygen, they will get rancid.
- What is meant by rancidity?
  - How do you preserve butter?
  - Why chips are packed with nitrogen gas?
24. Three students, A, B and C were asked to prepare the Lassaigne's extract independently by fusing the compound with sodium. Then, they added solid ferrous sulphate and dilute sulphuric acid to a part of Lassaigne's extract. Both A and B got Prussian blue colour but C got red colour. Can you help them with equations and reasons? Write the chemical equations to explain the formation of compounds of different colours.

Or

- What is the principle of chromatography?
  - How can forgery be detected with the help of chromatography?
  - Is it possible to separate components of orange ink by chromatography?
  - Name the stationary and mobile phase in paper chromatography.
  - What is the suitable adsorbent in the process of column chromatography?
25. Give reasons:
- Why silicones are used for nipples of feeding bottles?
  - Why are silicones used in cosmetic plants?
  - Why silicones are water-repellant?
  - Why are silicones thermally stable?
  - Are silicones safe for environment?

Or

- If a salt 'A' gives the following results:
    - Its aqueous solution is alkaline to litmus.
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- b) On strong heating, 'A' swells up to a glassy material 'B'.
- c) When concentrated HCl is added to a hot solution of 'A' white crystals of an acid 'C' separates out.

Write the chemical equations for the reactions and identify 'A', 'B' and 'C'.

ii) Complete the equations:



26.

- i) What is the change in internal energy in a process, 701 J of heat is absorbed by a system and 394 J of work is done by the system.
- ii) The equilibrium constant for the reaction is 10. Calculate the value of  $\Delta G^\theta$ . Given  $R = 8.0$  J/mol,  $T = 300$  K.

Or

Calculate the lattice energy for the change of  $\text{Li}^+(\text{g}) + \text{Cl}^-(\text{g}) \rightarrow \text{LiCl}(\text{s})$ ,

$\Delta_{\text{sub}} H^\theta$  of Li = 160.67 kJ/mol,  $\Delta_{\text{diss}} H^\theta$  of  $\text{Cl}_2 = 244.34$  kJ/mol,

$\Delta_{\text{le}} H^\theta$  of Li (g) = 520.07 kJ/mol,  $\Delta_{\text{eg}} H^\theta$  of Cl (g) = - 365.26 kJ/mol,

$\Delta_{\text{r}} H^\theta$  of LiCl (s) = - 401.66 kJ/mol.

Is the reaction spontaneous or not?

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**SAMPLE PAPER-03 (solved)**

CHEMISTRY (Theory)

Class - XI

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Time allowed: 3 hours

**Answer**

Maximum Marks: 70

1. It is because acetone has weak van der Waals' forces of attraction whereas water molecules have strong hydrogen bonding, therefore,  $\Delta_{\text{vap}}H^0$  of water is more.
  2. 1 molecule of methane =  $6 + 4 = 10$  electrons  
16 of methane contains  $10 \times 6.022 \times 10^{23}$  electrons =  $6.022 \times 10^{24}$  electrons.
  3.  $1000 \text{ cm}^3$  of  $0.15 \text{ M Na}_2\text{CO}_3$  contains  $0.15$  moles of  $\text{Na}_2\text{CO}_3$   
So,  $100 \text{ cm}^3$  of  $0.15 \text{ M Na}_2\text{CO}_3$  will contain = 
$$= \frac{0.15 \times 100}{1000}$$
$$= 0.015 \text{ moles of Na}_2\text{CO}_3$$
  
Number of moles = 
$$\frac{\text{Mass}}{\text{Molar mass}}$$
$$\therefore 0.015 = \frac{\text{Mass}}{106\text{g}}$$
Molar mass of  $\text{Na}_2\text{CO}_3 = 106 \text{ g}$ 
$$= 0.015 \times 106 = 1.59 \text{ g}$$
  4. 1L of nitrogen will react with 3L of hydrogen,  
2L of nitrogen will react with 6L of hydrogen, but we have 2L of hydrogen, therefore hydrogen is limiting reactant.  
So, 3L of hydrogen gives 2L of ammonia,  
2L of hydrogen gives  $\frac{2}{3} \times 2 = \frac{4}{3} = 1.33\text{L}$  of ammonia.
  5.  $\Delta H$ , the enthalpy change and  $\Delta U$ , the internal energy change are state functions as they depend on initial and final state and not on the path.
  6. No. Here, the rate of evaporation is still constant but since the molecules get dispersed into a large open volume, the rate of condensation from gas to liquid state can never be equal to the rate of evaporation eventually completely.
  - 7.
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iii. Ozone



13.

- If the concentration of SO<sub>2</sub> is increased the equilibrium will shift in the forward direction to consume the reactant SO<sub>2</sub>.
- If the concentration of SO<sub>3</sub> is increased the equilibrium will shift in the backward direction to consume the product SO<sub>3</sub>.
- If the temperature is increased, the equilibrium will shift in the backward direction as the increase in temperature will be compensated by absorbing heat

14. Since,  $\nu = \frac{c}{\lambda}$

Substituting the values,

$$\Lambda = 686 \text{ nm}$$

Here,  $n_1 = 6$  and  $n_2 = 1$ ,

The energy gap between two orbits for a hydrogen atom is given as,

$$\Delta E = 2.18 \times 10^{-18} \left( \frac{1}{n_1^2} - \frac{1}{n_2^2} \right)$$

$$= -2.11 \times 10^{-18} \text{ J}$$

Since  $\Delta E$  is negative energy, the frequency of photon is given by

$$\nu = \frac{\Delta E}{h}$$

Substituting, we get

$$\nu = 3.18 \times 10^{15} \text{ Hz}$$

15. :

- Lithium and magnesium follow diagonal relationship and so lithium like magnesium forms nitride while other alkali metals do not.
  - Size of O<sup>2-</sup> ion is smaller than SO<sub>4</sub><sup>2-</sup>. Since a bigger anions stabilizes bigger cation more than a smaller cation stabilizes a bigger anion, lattice enthalpy of BaO is smaller than BaSO<sub>4</sub>. BaO is soluble as hydration energy is more than lattice energy but BaSO<sub>4</sub> (as hydration energy is less than lattice energy) is insoluble in water.
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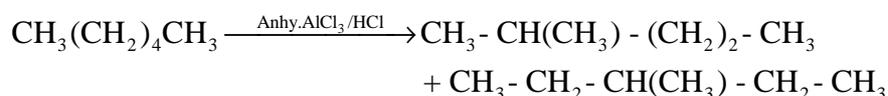
16. Since in the above reaction the compound formed is an ionic compound, which may also be represented as  $\text{Na}^+\text{H}^-$  (s), this suggests that one half reaction in this process is:  $2 \text{Na} (\text{s}) \rightarrow 2 \text{Na}^+(\text{g}) + 2\text{e}^-$  and the other half reaction is:  $\text{H}_2 (\text{g}) + 2\text{e}^- \rightarrow 2 \text{H}^-(\text{g})$ . This splitting of the reaction under examination into two half reactions automatically reveals that here sodium is oxidised and hydrogen is reduced, therefore, the complete reaction is a redox change.

17. :

- i) Graphite has layered structure in which the different layers are held together by weak Vander Waals forces and hence can be made to slip over one another. So, graphite acts as a lubricant.
- ii) Since diamond is very hard, it can be used as an abrasive.
- iii) Aluminium alloys are light, tough and resistant to corrosion and so are used to make aircraft body.

Or

The n-Alkanes on heating in the presence of anhydrous aluminium chloride and hydrogen chloride gas isomerize to branched chain alkanes. Major products are given below. Some minor products are also possible but are not reported in organic reactions.



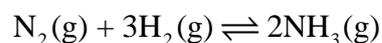
18. :

- a)  $\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CH}_2 - \text{CH}_2 - \text{Br} + \text{alc. KOH} \rightarrow \text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CH} = \text{CH}_2$
- b)  $\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CH} = \text{CH}_2 + \text{Br}_2 (\text{aq}) \rightarrow \text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CH}(\text{Br}) - \text{CH}_2 - \text{Br}$
- c)  $\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CH}(\text{Br}) - \text{CH}_3 + 2 \text{alc. KOH} \rightarrow \frac{1}{2} \text{H}_2(\text{g}) + \text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{C} \equiv \text{CNa}$
- d)  $\frac{1}{2} \text{H}_2(\text{g}) + \text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{C} \equiv \text{CNa} \xrightarrow{\text{Na}/\text{NH}_3} \text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{C} \equiv \text{CH}$
- e)  $\text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{C} \equiv \text{CH} \xrightarrow{\text{Ni}/2\text{H}_2} \text{CH}_3 - \text{CH}_2 - \text{CH}_2 - \text{CH}_2 - \text{CH}_3$

19. Moles of nitrogen =  $\frac{\text{Mass}}{\text{Molar mass}} = 1.786 \times 10^3 \text{ mol}$

Moles of hydrogen =  $\frac{\text{Mass}}{\text{Molar mass}} = 5.0 \times 10^3 \text{ mol}$

Then, to identify the limiting reagent



1 mole of nitrogen gas reacts with 3 moles of hydrogen gas

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Therefore  $1.786 \times 10^3$  mol of nitrogen will react with  $\frac{3 \times 1.786 \times 10^3}{1}$  moles of hydrogen =  $5.36 \times 10^3$  mol

Since, there is  $5.0 \times 10^3$  mol of hydrogen only, it is the limiting reagent.

To calculate the amount of ammonia formed,

3 moles of hydrogen gives 2 moles of ammonia,

Therefore  $5.0 \times 10^3$  moles of hydrogen gives  $\frac{2}{3} \times 5 \times 10^3$  of ammonia =  $3.3 \times 10^3$  moles of ammonia.

Mass of  $\text{NH}_3$  produced =  $3.3 \times 10^3 \times 17$  g of  $\text{NH}_3$  = 56.1 kg

20. :

- i) This is due to the reason that the molecules which undergo evaporation are high energy molecules and therefore, the kinetic energy of the remaining molecules becomes less. Since the remaining molecules have lower average kinetic energy, their temperature becomes low.
- ii) This is due to surface tension of liquids. Due to surface tension, the molecules of a liquid, try to make surface area to be minimum and for a given volume, sphere has the minimum surface area. Therefore the falling liquid drops are spherical.
- iii) Intermolecular forces are stronger in acetone than in ether. Thus the vapour pressure of acetone is less than ether.

21. :

- i. Mists are formed when certain herbicides and insecticides are sprayed in the liquid form over the plants.
- ii. Smokes are small particles of soot which are released in atmosphere in the form of oil, smoke etc by incomplete combustion of organic matter.
- iii. Fumes are released to atmosphere in metallurgical operations due to result of reactions in factories.
- iv. Dusts are released by grinding limestone, cement as fly ash etc.

22. Enthalpy of vaporization of CO = 6.04 kJ/mol

Molar mass of CO = 28 g/mol

Enthalpy change for vaporization of 28 g of CO at boiling point = 6.04 kJ

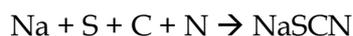
Therefore, the enthalpy change for vapourization of 2.38 g of CO at boiling point will be,

$$\frac{6.04 \times 2.38}{28} = 0.5134 \text{ kJ} = 513.4 \text{ J}$$

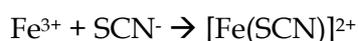
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23. :
- a) The spoilage of food due to oxidation is called rancidity.
  - b) We preserve butter at low temperature. Adding common salt and antioxidants to prevent it from getting spoiled.
  - c) It is done to prevent it from oxidation.

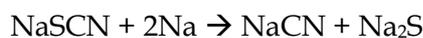
24. If S and N both are present, it result in the formation of NaSCN



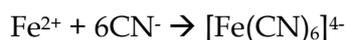
It also reacts with  $\text{Fe}^{3+}$  and gives red colour due to the formation of ferric thiocyanate.



This takes place when fusion is not carried out in excess of sodium. If excess of sodium is used, then thiocyanate ion will change to NaCN and sodium cyanide.



The Prussian blue colour is formed due to the formation of ferric ferro cyanide.



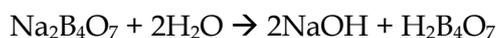
Or

- a) It is based on the principle of differential adsorption.
  - b) Different inks have different adsorbing power; therefore it can be detected by chromatography.
  - c) Yes.
  - d) Liquid on paper is stationary phase and the solution of substance to be separated is mobile phase.
  - e) Alumina.
25. :
- a) Because it is thermally stable.
  - b) Because they are biocompatible.
  - c) Since they are surrounded by non-polar alkyl groups which is water repelling.
  - d) Because they are highly polar and have strong forces of attraction.
  - e) Yes.

Or

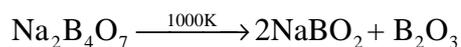
i) 'A' is borax ( $\text{Na}_2\text{B}_4\text{O}_7 \cdot 10\text{H}_2\text{O}$ )

a) The aqueous solution of borax is alkaline to litmus.



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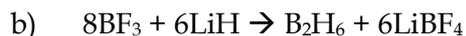
b) Borax swells to a glassy material on strong heating to give sodium metaborate, 'B'



c) On adding hot solution of 'A' to conc. HCl, it gives boric acid, 'C'.



ii)



26.

i) Heat absorbed by the system (q) = + 701 J

Work done by the system (w) = - 394 J

Change in the internal energy ( $\Delta U$ ) = q + w = + 701 + (- 394) = + 307 J

ii)  $\Delta G^\theta = - 2.303 RT \log K$

Substituting, the values, we get = - 5527.2 J/mol

Or

$$\Delta_f H^\theta = \Delta_{\text{sub}} H^\theta + \Delta_{\text{le}} H^\theta + \frac{1}{2} \Delta_{\text{diss}} H^\theta + \Delta_{\text{eg}} H^\theta + \Delta_{\text{lattice}} H^\theta$$

$\Delta_{\text{lattice}} H^\theta = - 839.31 \text{ kJ/mol}$ ,  $\Delta U^\theta = - 10.5 \text{ kJ}$ ,

$\Delta n (\text{g}) = - 1 \text{ mol}$ ,  $T = 298 \text{ K}$ ,  $R = 8.314 \times 10^{-3} \text{ kJ/K/mol}$

$\Delta H^\theta = \Delta U^\theta + \Delta n RT$

Substituting, we get

$\Delta H^\theta = - 12.978 \text{ kJ}$

Substituting the values for  $\Delta G^\theta = \Delta H^\theta - T\Delta S^\theta$ , we get

$\Delta G^\theta = - 2816.2 \text{ J}$

Since the value of  $\Delta G^\theta$  is negative, the reaction is spontaneous.

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