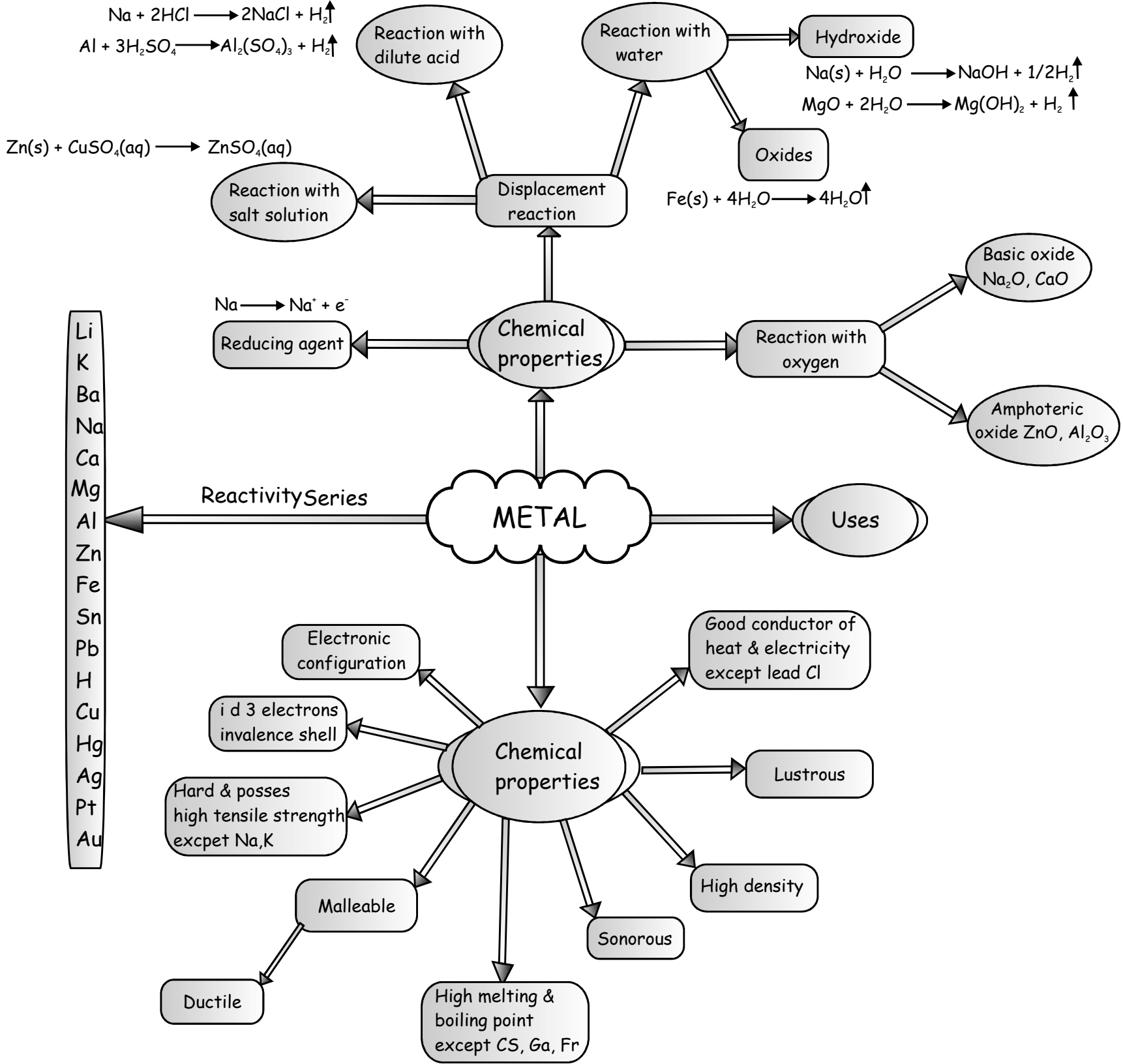
Na + 2HCl → 2NaCl + H2↑
Al + 3H2SO4 → Al2(SO4)3 + H2↑
Reaction with dilute acid
Reaction with water
Hydroxide
Na(s) + H2O → NaOH + 1/2H2↑
MgO + 2H2O → Mg(OH)2 + H2 ↑
Oxides
Fe(s) + 4H2O → 4H2O↑
Zn(s) + CuSO4(aq) → ZnSO4(aq)
Reaction with salt solution
Displacement reaction
Basic oxide Na2O, CaO
Na → Na+ + e-
Reducing agent
Chemical properties
Reaction with oxygen
Amphoteric oxide ZnO, Al2O3
Li
K
Ba
Na
Ca
Mg
Al
Zn
Fe
Sn
Pb
H
Cu
Hg
Ag
Pt
Au
Reactivity Series
METAL
Uses
Electronic configuration
Good conductor of heat & electricity except lead Cl
i d 3 electrons invalence shell
Chemical properties
Lustrous
Hard & posses high tensile strength excpet Na,K
High density
Malleable
Sonorous
Ductile
High melting & boiling point except CS, Ga, Fr