

Sample Question Paper - 6
Chemistry (043)
Class- XII, Session: 2021-22
TERM II

Time allowed : 2 hours

Maximum marks : 35

General Instructions :

Read the following instructions carefully.

1. There are 12 questions in this question paper with internal choice.
2. SECTION A - Q. No. 1 to 3 are very short answer questions carrying 2 marks each.
3. SECTION B - Q. No. 4 to 11 are short answer questions carrying 3 marks each.
4. SECTION C - Q. No. 12 is case based question carrying 5 marks.
5. All questions are compulsory.
6. Use of log tables and calculators is not allowed.

SECTION - A

1. Give the simple chemical tests to distinguish between the following pairs of compounds (*any two*) :
 - (a) Benzaldehyde and benzoic acid
 - (b) Propanal and propanone
 - (c) Benzaldehyde and acetophenone
2. Write the expression for the rate of reaction given below.
$$5\text{Br}^-_{(aq)} + \text{BrO}^-_{3(aq)} + 6\text{H}^+_{(aq)} \longrightarrow 3\text{Br}_{2(aq)} + 3\text{H}_2\text{O}_{(l)}$$
3. Name two reactions that can be used to increase number of carbon atoms in the chain of a carbonyl compound.

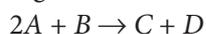
SECTION - B

4. Compare qualitatively the first and second ionisation potentials of copper and zinc. Explain the observation.

OR

How would you account for the following :

- (a) Of the d^4 species Cr^{2+} is strongly reducing while manganese (III) is strongly oxidising.
 - (b) Cobalt (II) is stable in aqueous solution but in the presence of complexing reagents it is easily oxidised.
 - (c) The d^1 configuration is very unstable in ions.
5. The following results have been obtained during the kinetic studies of the reaction :



Experiment	[A] / mol L ⁻¹	[B] / mol L ⁻¹	Initial rate of formation of D/mol L ⁻¹ min ⁻¹
I	0.1	0.1	6.0×10^{-3}
II	0.3	0.2	7.2×10^{-2}
III	0.3	0.4	2.88×10^{-1}
IV	0.4	0.1	2.40×10^{-2}

Determine the rate law and the rate constant for the reaction.

OR

While studying the decomposition of gaseous N_2O_5 it is observed that a plot of logarithm of its partial pressure versus time is linear. What kinetic parameters can be obtained from this observation?

6. Aqueous copper sulphate solution (blue in colour) gives:
- (a) a green precipitate with aqueous potassium fluoride and
 - (b) a bright green solution with aqueous potassium chloride.
- Explain these experimental results.
7. How will you convert :
- (a) Ethanoic acid into methanamine?
 - (b) Hexanenitrile into 1-aminopentane?
 - (c) Methanol to ethanoic acid?
8. What is an adsorption isotherm? Describe Freundlich adsorption isotherm.

OR

Explain the cleaning action of soap. Why do soaps not work in hard water?

9. Write short notes on the following :
- (a) Ammonolysis
 - (b) Gabriel phthalimide synthesis
10. Write the formulae of the following coordination compounds :
- (a) Tetraammineaquachloridocobalt (III) chloride.
 - (b) Potassium tetrahydroxozincate (II).
 - (c) Tetracabonylnickel (0).
11. Write structural formulae and names of four possible aldol condensation products from propanal and butanal. In each case, indicate which aldehyde acts as nucleophile and which as electrophile.

OR

An organic compound with the molecular formula $C_9H_{10}O$ forms 2,4-DNP derivative, reduces Tollens' reagent and undergoes Cannizzaro reaction. On vigorous oxidation, it gives 1,2-benzenedicarboxylic acid. Identify the compound.

SECTION - C

12. Read the passage given below and answer the questions that follow.
- All chemical reactions involve interaction of atoms and molecules. A large number of atoms/molecules are present in a few gram of any chemical compound varying with their atomic/molecular masses. To handle such large number conveniently, the mole concept was introduced. All electrochemical cell reactions are also based on mole concept. For example, a 4.0 molar aqueous solution of NaCl is prepared and 500 mL of this solution is electrolysed. This leads to the evolution of chlorine gas at one of the electrode. The amount of products formed can be calculated by using mole concept.
- (a) Calculate the total number of moles of chlorine gas evolved.
 - (b) Assuming Hg electrode as cathode, find out the maximum weight of amalgam formed by given solution.
 - (c) What is the total charge (coulomb) required for complete electrolysis?
 - (d) Find out the number of electrons involved in the complete electrolysis of NaCl.

OR

In electrolysis of NaCl, when Pt electrode is taken, then which gas is liberated at cathode?

Solution

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Class 12 - Chemistry

1. (a) Benzaldehyde and benzoic acid can be distinguished by sodium bicarbonate test.

Benzoic acid will give effervescence with NaHCO_3 but benzaldehyde will not react.

(b) Propanal and propanone can be distinguished by their reactions with Tollens' reagent.

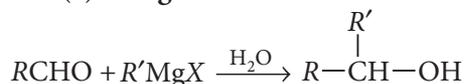
Propanal will form the silver mirror, but propanone does not react.

(c) Benzaldehyde and acetophenone can be distinguished by Tollens' test.

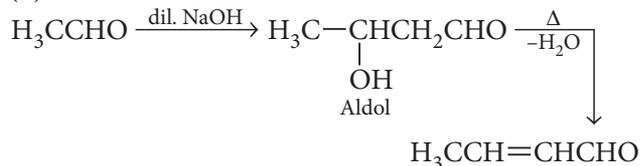
Benzaldehyde will form silver mirror, on treatment with Tollens' reagent whereas acetophenone will not show Tollens' test.

$$\begin{aligned} 2. \text{ Rate} &= -\frac{1}{5} \frac{\Delta[\text{Br}^-]}{\Delta t} = -\frac{\Delta[\text{BrO}_3^-]}{\Delta t} = -\frac{1}{6} \frac{\Delta[\text{H}^+]}{\Delta t} \\ &= \frac{1}{3} \frac{\Delta[\text{Br}_2]}{\Delta t} = \frac{1}{3} \frac{\Delta[\text{H}_2\text{O}]}{\Delta t} \end{aligned}$$

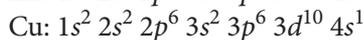
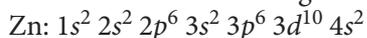
3. (a) **Grignard reaction :**



(b) **Aldol condensation :**



4. The electronic configuration of Zn and Cu are:

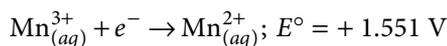
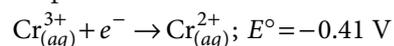


From the above configuration it is clear that first ionisation energy of Zn is greater than that of Cu (because of $4s^2$ and $4s^1$ configuration of Zn and Cu respectively). More energy is needed to remove an electron of $4s^2$ than that of $4s^1$.

The second *I.E.* of Cu is higher than that of Zn because for Cu^{2+} the configuration is $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10}$ and for Zn^+ the configuration is $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^1$, it is easier to remove $4s^1$ electron of Zn^+ than a $3d$ -electron from $3d^{10}$ (stable configuration).

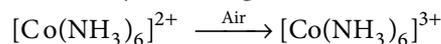
OR

(a) E° values for the $\text{Cr}^{3+}/\text{Cr}^{2+}$ and $\text{Mn}^{3+}/\text{Mn}^{2+}$ couples are



These E° values indicate that Cr^{2+} is strongly reducing while Mn^{3+} is strongly oxidising.

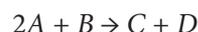
(b) The tendency to form complexes is high for Co(III) as compared to Co(II) . Co^{2+} ions are very stable and are difficult to oxidise. Co^{3+} ions are less stable and are reduced by water. In contrast many Co(II) complexes are readily oxidised to Co(III) complexes and Co(III) complexes are very stable, e.g.,



This happens because the crystal field stabilisation energy of Co(III) with a $d^6(t_{2g}^6)$ configuration is higher than for Co(II) with a $d^7(t_{2g}^6 e_g^1)$ arrangement.

(c) Transition metal ions with d^1 configuration have a tendency to lose this single electron and give ion with a noble gas configuration.

5. For the reaction



The rate law will be; $\text{Rate} = k[\text{A}]^x[\text{B}]^y$

From the given data

$$(r_0)_1 = 6.0 \times 10^{-3} = k_1 (0.1)^x (0.1)^y \quad \dots(1)$$

$$(r_0)_2 = 7.2 \times 10^{-2} = k_2 (0.3)^x (0.2)^y \quad \dots(2)$$

$$(r_0)_3 = 2.88 \times 10^{-1} = k_3 (0.3)^x (0.4)^y \quad \dots(3)$$

$$(r_0)_4 = 2.4 \times 10^{-2} = k_4 (0.4)^x (0.1)^y \quad \dots(4)$$

From equation (1) and (4)

$$\frac{(r_0)_4}{(r_0)_1} = \frac{2.4 \times 10^{-2}}{6.0 \times 10^{-3}} = \frac{k(0.4)^x (0.1)^y}{k(0.1)^x (0.1)^y}$$

$$4 = \left(\frac{0.4}{0.1}\right)^x = 4^x \quad \therefore x = 1$$

Thus, the order with respect to A is 1.

From equations (2) and (3)

$$\frac{(r_0)_3}{(r_0)_2} = \frac{2.88 \times 10^{-1}}{7.2 \times 10^{-2}} = \frac{k(0.3)^x (0.4)^y}{k(0.3)^x (0.2)^y}$$

$$4 = \left(\frac{0.4}{0.2}\right)^y \Rightarrow 4 = 2^y \Rightarrow y = 2$$

Thus, the order with respect to B is 2.

The rate law of the reaction is

$$\text{Rate} = k[\text{A}][\text{B}]^2$$

Calculation of rate constant

$$\text{From the rate law, } k = \frac{\text{Rate}}{[\text{A}][\text{B}]^2}$$

From equation (1) (substituting the values of x and y)

$$k_1 = \frac{6.0 \times 10^{-3}}{(0.1)(0.1)^2} = 6.0 \text{ mol}^{-2} \text{ L}^2 \text{ min}^{-1}$$

From equation (2)

$$k_2 = \frac{7.2 \times 10^{-2}}{(0.3)(0.2)^2} = 6.0 \text{ mol}^{-2} \text{ L}^2 \text{ min}^{-1}$$

From equation (3)

$$k_3 = \frac{2.88 \times 10^{-1}}{(0.3)(0.4)^2} = 6.0 \text{ mol}^{-2} \text{ L}^2 \text{ min}^{-1}$$

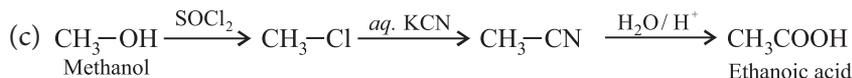
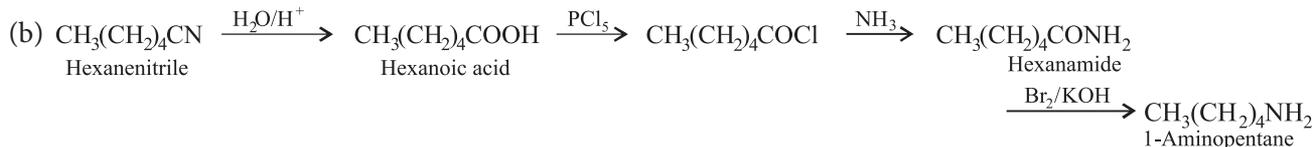
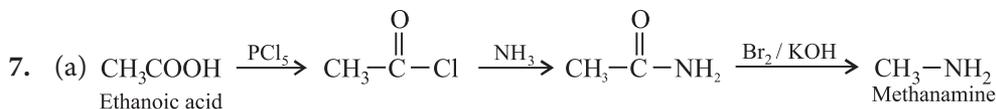
From equation (4)

$$k_4 = \frac{2.4 \times 10^{-2}}{(0.4)(0.1)^2} = 6.0 \text{ mol}^{-2} \text{ L}^2 \text{ min}^{-1}$$

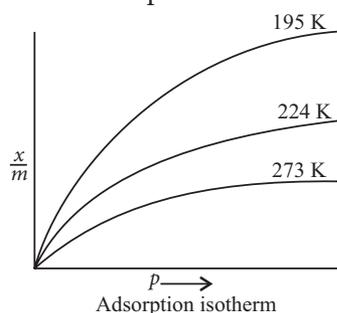
∴ Average $k = 6.0 \text{ mol}^{-2} \text{ L}^2 \text{ min}^{-1}$

OR

Let us assume the decomposition of N_2O_5 to be a first order reaction, then



8. Adsorption isotherm : It is the variation in the amount of gas adsorbed by the adsorbent with pressure at constant temperature.



These curves indicate that on increasing temperature, physical adsorption decreases at a fixed pressure.

Freundlich adsorption isotherm : It is an empirical relationship between the quantity of gas adsorbed by unit mass of solid adsorbent and pressure at a particular temperature.

$$\frac{x}{m} = kp^{1/n} \quad (n > 1) \quad \dots \text{(i)}$$

when, $n = 1, \Rightarrow \frac{x}{m} = kp$

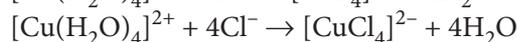
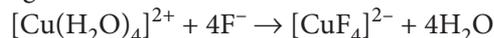
$$k = \frac{2.303}{t} \log \frac{a}{a-x} \quad \text{or} \quad k = \frac{2.303}{t} \log \frac{P_0}{P}$$

$$\text{or} \quad \log P = \frac{-kt}{2.303} + \log P_0$$

Thus a graph of $\log P$ vs t will be linear.

This is in accordance with given statement so the decomposition of N_2O_5 is a first order reaction.

6. An aqueous CuSO_4 solution has $[\text{Cu}(\text{H}_2\text{O})_4]^{2+}$ as the coordination entity which is of blue colour. The blue colour of $[\text{Cu}(\text{H}_2\text{O})_4]^{2+}$ ions changes to green by the addition of F^- or Cl^- ions because both F^- and Cl^- ions are weaker ligands than H_2O . These ions create small splitting (Δ), i.e., low energy light is required for small splitting. Thus light of red colour is absorbed and the $[\text{CuF}_4]^{2-}$ or $[\text{CuCl}_4]^{2-}$ complexes look green.



$$\text{or} \quad \frac{x}{m} \propto p$$

where, x is the mass of gas adsorbed on mass m of the adsorbent at pressure p . k and n are constants which depend on the nature of the adsorbent and the gas at the particular temperature.

Taking log in Eq. (i), gives

$$\log \frac{x}{m} = \log k + \frac{1}{n} \log p$$

The validity of Freundlich isotherm can be verified by plotting $\frac{x}{m}$ on y -axis and $\log p$ on x -axis.

If it comes to be a straight line, the Freundlich isotherm is valid.

