Chapter
12
Thermometry, Thermal Expansion and Calorimetry

Temperature

Temperature is defined as the degree of hotness or coldness of a body. The natural flow of heat is from higher temperature to lower temperature.

Two bodies are said to be in thermal equilibrium with each other, when no heat flows from one body to the other. That is when both the bodies are at the same temperature.

(1) Temperature is one of the seven fundamental quantities with dimension $[\theta]$. It is a scalar physical quantity with S.I. unit kelvin.

(2) When heat is given to a body and its state does not change, the temperature of the body rises and if heat is taken from a body its temperature falls *i.e.* temperature can be regarded as the effect of cause "heat".

(3) According to kinetic theory of gases, temperature (macroscopic physical quantity) is a measure of average translational kinetic energy of a molecule (microscopic physical quantity).

(4) Although the temperature of a body can to be raised without limit, it cannot be lowered without limit and theoretically limiting low temperature is taken to be zero of the kelvin scale.

(5) Highest possible temperature achieved in laboratory is about $10^8 K$ while lowest possible temperature attained is $10^{-8} K$.

(6) Temperature of the core of the sun is $10^7 K$ while that of its surface is 6000 K.

(7) Normal temperature of human body is 310. 15 K ($37°C = 98.6°F$).

(8) NTP or STP implies 273.15K ($0°C = 32°F$)

Scales of Temperature

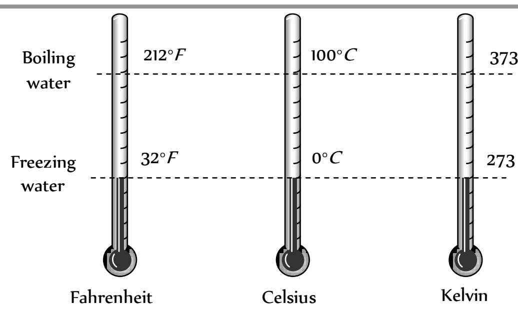


Fig. 12.1

The centigrade (°C), Farenheite (°F), Kelvin (K), Reaumer (R), Rankine (Ra) are commonly used temperature scales.

(1) To construct a scale of temperature, two fixed points are taken. First fixed point is the freezing point (ice point) of water, it is called lower fixed point (LFP). The second fixed point is the boiling point (steam point) of water, it is called upper fixed point (UFP).

(2) **Celsius scale :** In this scale LFP (ice point) is taken 0° and UFP (steam point) is taken 100°. The temperature measured on this scale all in degree Celsius (°C).

(3) **Farenheite scale :** This scale of temperature has LFP as 32°F and UFP as 212°F. The change in temperature of 1°F corresponds to a change of less than 1° on Celsius scale.

(4) **Kelvin scale :** The Kelvin temperature scale is also known as thermodynamic scale. The triple point of water is also selected to be the zero of scale of temperature. The temperature measured on this scale are in Kelvin (K).

The triple point of water is that point on a P-T diagram where the three phases of water, the solid, the liquid and the gas, can coexist in equilibrium.

Table 12.1 : Different measuring scales

Scale	Symbol for each degree	LFP	UFP	Number of divisions on the scale
Celsius	$°C$	$0°C$	$100°C$	100
Fahrenheit	$°F$	$32°F$	$212°F$	180
Reaumer	$°R$	$0°R$	$80°R$	80
Rankine	$°Ra$	460 Ra	672 Ra	212
Kelvin	K	273.15 K	373.15 K	100

(5) Temperature on one scale can be converted into other scale by using the following identity.

$$\frac{\text{Reading on any scale} - \text{LFP}}{\text{UFP} - \text{LFP}} = \text{Constant for all scales}$$

(6) All these temperatures are related to each other by the following relationship

$$\frac{C-0}{100} = \frac{F-32}{212-32} = \frac{K-273.15}{373.15-273.15} = \frac{R-0}{80-0} = \frac{Ra-460}{672-460}$$

or $$\frac{C}{5} = \frac{F-32}{9} = \frac{K-273}{5} = \frac{R}{4} = \frac{Ra-460}{10.6}$$

(7) The Celsius and Kelvin scales have different zero points but the same size degrees. Therefore any temperature difference is the same on the Celsius and Kelvin scales $(T_1 - T_2)°C = (T_1 - T_2)\,K$.

Thermometry

A branch of science which deals with the measurement of temperature of a substance is known as thermometry.

(1) The linear variation in some physical properties of a substance with change of temperature is the basic principle of thermometry and these properties are defined as thermometric property (x) of the substance.

(2) Thermometric properties (x) may be as follows

(i) Length of liquid in capillary

(ii) Pressure of gas at constant volume.

(iii) Volume of gas at constant pressure.

(iv) Resistance of a given platinum wire.

(3) In old thermometry, freezing point ($0°C$) and steam point ($100°C$) are taken to define the temperature scale. So if the thermometric property at temperature $0°C$, $100°C$ and $t°C$ are x_0, x_{100} and x respectively then

$$\frac{t-0}{100-0} = \frac{x-x_0}{x_{100}-x_0} \Rightarrow t°C = \frac{x-x_0}{x_{100}-x_0} \times 100°C$$

(4) In modern thermometry instead of two fixed points only one reference point is chosen (triple point of water 273.16 K) the other is itself 0 K where the value of thermometric property is assumed to be zero.

So if the value of thermometric property at 0 K, 273.16 K and TK are 0, x_{Tr} and x respectively then

$$\frac{T}{273.16} = \frac{x}{x_{Tr}} \Rightarrow T = 273.16\left[\frac{x}{x_{Tr}}\right]K$$

Thermometers

An instrument used to measure the temperature of a body is called a thermometer

It works by absorbing some heat from the body, so the temperature recorded by it is lesser than the actual value unless the body is at constant temperature. Some common types of thermometers are as follows

(1) **Liquid (mercury) thermometers :** In liquid thermometers mercury is preferred over other liquids as its expansion is large and uniform and it has high thermal conductivity and low specific heat.

(i) Range of temperature : $\underset{\text{(freezing point)}}{-50°C}$ to $\underset{\text{(boiling point)}}{350°C}$

(ii) Upper limit of range of mercury thermometer can be raised upto $550°C$ by filling nitrogen in space over mercury under pressure (which elevates boiling point of mercury).

(iii) Mercury thermometer with cylindrical bulbs are more sensitive than those with spherical bulbs.

(iv) If alcohol is used instead of mercury then range of temperature measurement becomes – $80°C$ to $350°C$

(v) Formula : $t = \frac{l-l_0}{l_{100}-l_0} \times 100°C$

(2) **Gas thermometers :** These are more sensitive and accurate than liquid thermometers as expansion of gases is more than that of liquids. The thermometers using a gas as thermoelectric substance are called ideal gas thermometers. These are of two types

(i) **Constant pressure gas thermometers**

(a) Principle $V \propto T$ (if P = constant)

(b) Formula : $t = \frac{V-V_0}{V_{100}-V_0} \times 100°C$ or $T = 273.16\frac{V}{V_{Tr}}K$

(ii) **Constant volume gas thermometers**

(a) Principle $P \propto T$ (if V = constant)

(b) Formula : $t = \frac{P-P_0}{P_{100}-P_0} \times 100°C$ or $T = 273.16\frac{P}{P_{Tr}}K$

(c) Range of temperature :

Hydrogen gas thermometer : – $200°C$ to $500°C$

Nitrogen gas thermometer : – $200°C$ to $1600°C$

Helium gas thermometer : – $268°C$ to $500°C$

(3) **Resistance thermometers :** Usually platinum is used in resistance thermometers due to high melting point and large value of temperature coefficient of resistance.

Resistance of metals varies with temperature according to relation. $R = R_0(1+\alpha t)$ where α is the temperature coefficient of resistance and t is change in temperature.

(i) Formula : $t = \frac{R-R_0}{R_{100}-R_0} \times 100°C$ or $T = 273.16\frac{R}{R_{Tr}}K$

(ii) Temperature range : For Platinum resistance thermometer it is – $200°C$ to $1200°C$

For Germanium resistance thermometer it is 4 to 77 K.

(4) **Thermoelectric thermometers :** These are based on "Seebeck effect" according to which when two distinct metals are joined to form a closed circuit called thermocouple and the difference in temperature is maintained between their junctions, an emf is developed. The emf is called thermo-emf and if one junction is at $0°C$, thermoelectric emf varies with temperature of hot junction (t) according to $e = at + bt$; where a and b are constants.

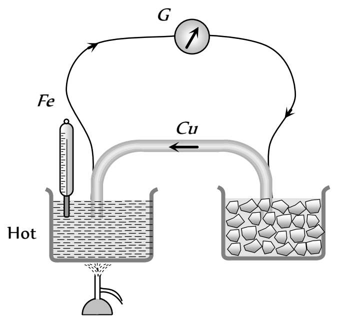


Fig. 12.2

Thermoelectric thermometers have low thermal capacity and high thermal conductivity, so can be used to measure quickly changing temperature

Table 12.2 : Different temperature range

Thermo couple	Temperature range
Copper-iron thermocouple	$0°C$ to $260°C$
Iron-constantan thermocouple	$0°C$ to $800°C$
Tungsten-molybdenum thermocouple	$2000°C$ to $3000°C$

(5) **Pyrometers :** These are the devices used to measure the temperature by measuring the intensity of radiations received from the body. They are based on the fact that the amount of radiations emitted from a body per unit area per second is directly proportional to the fourth power of temperature (Stefan's law).

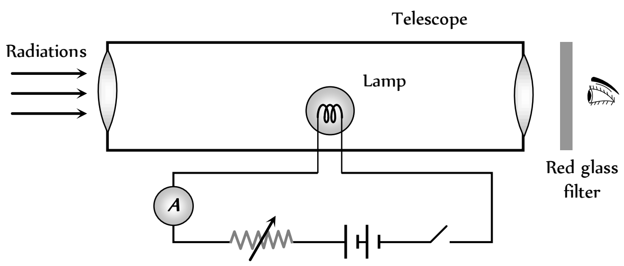


Fig. 12.3

(i) These can be used to measure temperatures ranging from $800°C$ to $6000°C$.

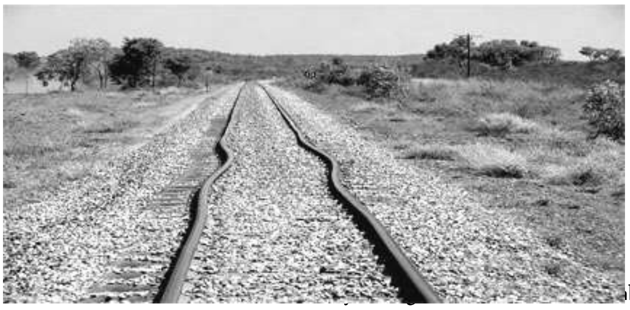
(ii) They cannot measure temperature below $800°C$ because the amount of radiations is too small to be measured.

(6) **Vapour pressure thermometer :** These are used to measure very low temperatures. They are based on the fact that saturated vapour pressure P of a liquid depends on the temperature according to the relation

$$\log P = a + bT + \frac{c}{T}$$

The range of these thermometers varies from 0.71 K to 120 K for different liquid vapours.

Thermal Expansion



lly expands. According to atomic theory of matter, a symmetry in potential energy curve is responsible for thermal expansion. As with rise in temperature the amplitude of vibration and hence energy of atoms increases, hence the average distance between the atoms increases. So the matter as a whole expands.

(1) Thermal expansion is minimum in case of solids but maximum in case of gases because intermolecular force is maximum in solids but minimum in gases.

(2) Solids can expand in one dimension (linear expansion), two dimension (superficial expansion) and three dimension (volume expansion) while liquids and gases usually suffers change in volume only.

(3) **Linear expansion :** When a solid is heated and it's length increases, then the expansion is called linear expansion.

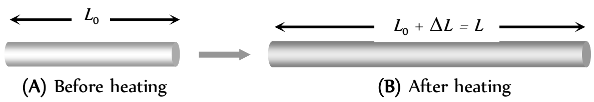


Fig. 12.4

(i) Change in length $\Delta L = L_0 \alpha \Delta T$

(L_0 = Original length, ΔT = Temperature change)

(ii) Final length $L = L_0 (1 + \alpha \Delta T)$

(iii) Co-efficient of linear expansion $\alpha = \dfrac{\Delta L}{L_0 \Delta T}$

(iv) Unit of α is $°C^{-1}$ or K^{-1}. It's dimension is $[\theta^{-1}]$

(4) **Superficial (areal) expansion :** When the temperature of a $2D$ object is changed, it's area changes, then the expansion is called superficial expansion.

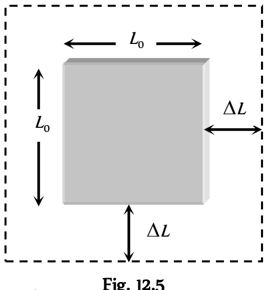


Fig. 12.5

(i) Change in area is $\Delta A = A_0 \beta \Delta T$

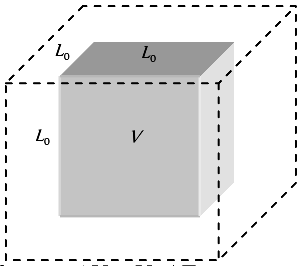
(A_0 = Original area, ΔT = Temperature change)

(ii) Final area $A = A_0 (1 + \beta \Delta T)$

(iii) Co-efficient of superficial expansion $\beta = \dfrac{\Delta A}{A_0 \Delta T}$

(iv) Unit of β is $°C$ or K.

(5) **Volume or cubical expansion :** When a solid is heated and it's volume increases, then the expansion is called volume or cubical expansion.



(i) Change in volume is $\Delta V = V_0 \gamma \Delta T$

Fig. 12.6

(V_0 = Original volume, ΔT = change in temperature)

(ii) Final volume $V = V_0(1 + \gamma \Delta T)$

(iii) Volume co-efficient of expansion $\gamma = \frac{\Delta V}{V_0 \Delta T}$

(iv) Unit of γ is $°C$ or K.

(6) **More about α, β and γ :** The co-efficient α, β and γ for a solid are related to each other as follows

$$\alpha = \frac{\beta}{2} = \frac{\gamma}{3} \Rightarrow \alpha : \beta : \gamma = 1 : 2 : 3$$

(i) Hence for the same rise in temperature

Percentage change in area = 2 × percentage change in length.

Percentage change in volume = 3 × percentage change in length.

(ii) The three coefficients of expansion are not constant for a given solid. Their values depend on the temperature range in which they are measured.

(iii) The values of α, β, γ are independent of the units of length, area and volume respectively.

(iv) For anisotropic solids $\gamma = \alpha_x + \alpha_y + \alpha_z$ where α_x, α_y and α_z represent the mean coefficients of linear expansion along three mutually perpendicular directions.

(7) **Contraction on heating :** Some rubber like substances contract with rising temperature, because transverse vibration of atoms of substance dominate over longitudinal vibration which is responsible for expansion.

Table 12.3 : α and γ for some materials

Material	$\alpha[K^{-1}$ or $(°C)^{-1}]$	$\gamma[K^{-1}$ or $(°C)^{-1}]$
Steel	1.2×10^{-5}	3.6×10^{-5}
Copper	1.7×10^{-5}	5.1×10^{-5}
Brass	2.0×10^{-5}	6.0×10^{-5}
Aluminium	2.4×10^{-5}	7.2×10^{-5}

Application of Thermal Expansion in Solids

(1) **Bi-metallic strip :** Two strips of equal lengths but of different materials (different coefficient of linear expansion) when join together, it is called "bi-metallic strip", and can be used in thermostat to break or make electrical contact. This strip has the characteristic property of bending on heating due to unequal linear expansion of the two metal. The strip will bend with metal of greater α on outer side *i.e.* convex side.

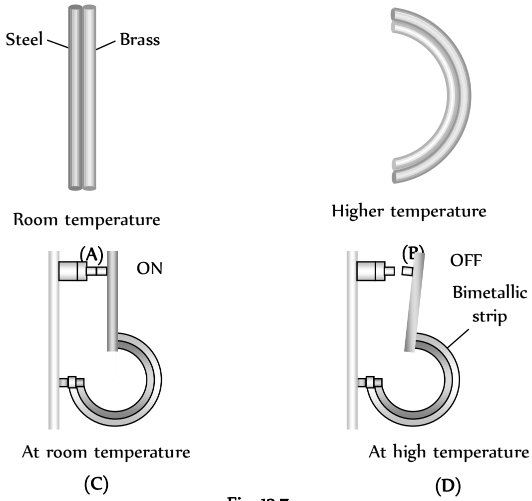


Fig. 12.7

(2) **Effect of temperature on the time period of a simple pendulum :** A pendulum clock keeps proper time at temperature θ. If temperature is increased to $\theta'(>\theta)$ then due to linear expansion, length of pendulum and hence its time period will increase.

Fractional change in time period $\frac{\Delta T}{T} = \frac{1}{2}\alpha\,\Delta\theta$

(i) Due to increment in its time period, a pendulum clock becomes slow in summer and will lose time.

Loss of time in a time period $\Delta T = \frac{1}{2}\alpha\ \Delta\theta\ T$

(ii) Time lost by the clock in a day (t = 86400 *sec*)

$$\Delta t = \frac{1}{2}\alpha\ \Delta\theta\ t = \frac{1}{2}\alpha\ \Delta\theta\ (86400) = 43200\,\alpha\ \Delta\theta\ sec$$

(iii) The clock will lose time *i.e.* will become slow if $\theta' > \theta$ (in summer)

and will gain time *i.e.* will become fast if $\theta' < \theta$ (in winter).

(iv) The gain or loss in time is independent of time period T and depends on the time interval t.

(v) Since coefficient of linear expansion (α) is very small for invar, hence pendulums are made of invar to show the correct time in all seasons.

(3) **Thermal stress in a rigidly fixed rod :** When a rod whose ends are rigidly fixed such as to prevent expansion or contraction, undergoes a change in temperature, due to thermal expansion or contraction, a compressive or tensile stress is developed in it. Due to this thermal stress the rod will exert a large force on the supports. If the change in temperature of a rod of length L is $\Delta\theta$ then

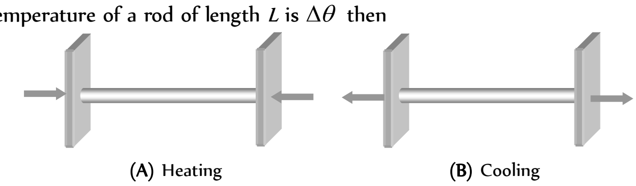


Fig. 12.8

Thermal strain $= \frac{\Delta L}{L} = \alpha\Delta\theta$ $\left[\text{As}\,\alpha = \frac{\Delta L}{L} \times \frac{1}{\Delta\theta}\right]$

So Thermal stress $= Y\alpha\Delta\theta$ $\left[\text{As } Y = \frac{\text{stress}}{\text{strain}}\right]$

or Force on the supports $F = YA\,\alpha\,\Delta\theta$

(4) **Error in scale reading due to expansion or contraction :** If a scale gives correct reading at temperature θ, at temperature $\theta'(>\theta)$ due to linear expansion of scale, the scale will expand and scale reading will be lesser than true value so that,

True value = Scale reading $[1+\alpha(\theta'-\theta)]$

i.e. $\text{TV} = SR[1+\alpha\,\Delta\theta]$ with $\Delta\theta = (\theta'-\theta)$

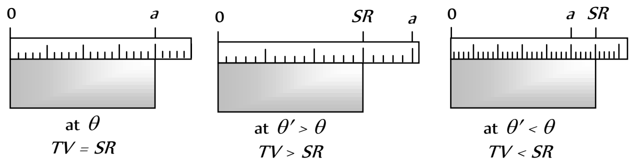
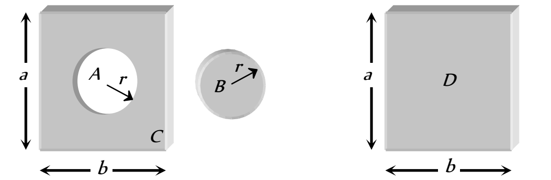


Fig. 12.9

However, if $\theta' < \theta$, due to contraction of scale, scale reading will be more than true value, so true value will be lesser than scale reading and will still be given by above equation with $\Delta\theta = (\theta'-\theta)$ negative.

(5) **Expansion of cavity :** Thermal expansion of an isotropic object may be imagined as a photographic enlargement. So if there is a hole *A* in a plate *C* (or cavity *A* inside a body *C*), the area of hole (or volume of cavity) will increase when body expands on heating, just as if the hole (or cavity) were solid *B* of the same material. Also the expansion of area (or volume) of the body *C* will be independent of shape and size of hole (or cavity), *i.e.*, will be equal to that of *D*.



Expansion of *A* = Expansion of *B* Expansion of *C* = Expansion of *D*

Fig. 12.10

(6) **Some other application**

(i) When rails are laid down on the ground, space is left between the ends of two rails.

(ii) The transmission cable are not tightly fixed to the poles.

(iii) Test tubes, beakers and crucibles are made of pyrex-glass or silica because they have very low value of coefficient of linear expansion.

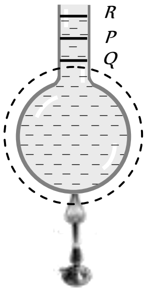
(iv) The iron rim to be put on a cart wheel is always of slightly smaller diameter than that of wheel.

(v) A glass stopper jammed in the neck of a glass bottle can be taken out by warming the neck of the bottle

Thermal Expansion in Liquids

(1) Liquids do not have linear and superficial expansion but these only have volume expansion.

(2) Since liquids are always to be heated along with a vessel which contains them so initially on heating the system (liquid + vessel), the level of liquid in vessel falls (as vessel expands more since it absorbs heat and liquid expands less) but later on, it starts rising due to faster expansion of the liquid.



$PQ \to$ represents expansion of vessel

$QR \to$ represents the real expansion of liquid

$PR \to$ Represent the apparent expansion of liquid

Fig. 12.11

(3) The actual increase in the volume of the liquid = The apparent increase in the volume of liquid + the increase in the volume of the vessel.

(4) Liquids have two coefficients of volume expansion.

(i) **Co-efficient of apparent expansion (γ) :** It is due to apparent (that appears to be, but is not) increase in the volume of liquid if expansion of vessel containing the liquid is not taken into account.

$$\gamma_a = \frac{\text{Apparent expansion in volume}}{\text{Initial vdume} \times \Delta\theta} = \frac{(\Delta V)_a}{V \times \Delta\theta}$$

(ii) **Co-efficient of real expansion (γ) :** It is due to the actual increase in volume of liquid due to heating.

$$\gamma_r = \frac{\text{Real increase in volume}}{\text{Initial vdume} \times \Delta\theta} = \frac{(\Delta V)_r}{V \times \Delta\theta}$$

(iii) Also coefficient of expansion of flask $\gamma_{Vessel} = \frac{(\Delta V)_{Vessel}}{V \times \Delta\theta}$

(iv) $\gamma_{Real} = \gamma_{Apparent} + \gamma_{Vessel}$

(v) Change (apparent change) in volume in liquid relative to vessel is

$$\Delta V_{app} = V\gamma_{app}\,\Delta\theta = V(\gamma_{Real} - \gamma_{Vessel})\Delta\theta = V(\gamma_r - 3\alpha)\Delta\theta$$

α = Coefficient of linear expansion of the vessel.

Table 12.4 : Different level of liquid in vessel

γ	ΔV	Level
$\gamma_{Real} > \gamma_{Vessel}\,(=3\alpha) \Rightarrow \gamma_{app} > 0$	ΔV_{app} is positive	Level of liquid in vessel will rise on heating.
$\gamma_{Real} < \gamma_{Vessel}\,(=3\alpha) \Rightarrow \gamma_{app} < 0$	ΔV_{app} is negative	Level of liquid in vessel will fall on heating.
$\gamma_{Real} = \gamma_{Vessel}\,(=3\alpha) \Rightarrow \gamma_{app} = 0$	$\Delta V_{app} = 0$	level of liquid in vessel will remain same.

(5) **Anomalous expansion of water :** Generally matter expands on heating and contracts on cooling. In case of water, it expands on heating if its temperature is greater than $4°C$. In the range $0°C$ to $4°C$, water contracts on heating and expands on cooling, *i.e.* γ is negative. This behaviour of water in the range from $0°C$ to $4°C$ is called anomalous expansion.

This anomalous behaviour of water causes ice to form first at the surface of a lake in cold weather. As winter approaches, the water temperature increases initially at the surface. The water there sinks because of its increased density. Consequently, the surface reaches $0°C$ first and the lake becomes covered with ice. Aquatic life is able to survive the cold winter as the lake bottom remains unfrozen at a temperature of about $4°C$.

At $4°C$, density of water is maximum while its specific volume is minimum.

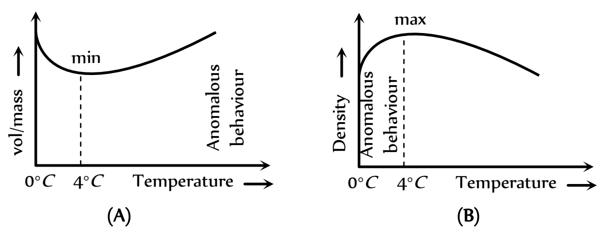


Fig. 12.12

(6) **Effect of temperature on upthrust :** The thrust on V volume of a body in a liquid of density σ is given by $Th = V\sigma g$

Now with rise in temperature by $\Delta\theta\,°C$, due to expansion, volume of the body will increase while density of liquid will decrease according to the relations $V' = V(1+\gamma_S\Delta\theta)$ and $\sigma' = \sigma/(1+\gamma_L\Delta\theta)$

So the thrust $Th' = V'\sigma' g \Rightarrow \quad \dfrac{Th'}{Th} = \dfrac{V'\sigma' g}{V\sigma g} = \dfrac{(1+\gamma_S\Delta\theta)}{(1+\gamma_L\Delta\theta)}$

and apparent weight of the body W_{app} = Actual weight – Thrust

As $\gamma_S < \gamma_L \;\; \therefore \;\; Th' < Th$ with rise in temperature thrust also decreases and apparent weight of body increases.

Variation of Density with Temperature

Most substances (solid and liquid) expand when they are heated, *i.e.*, volume of a given mass of a substance increases on heating, so the density should decrease $\left(\text{as } \rho \propto \dfrac{1}{V}\right)$. For a given mass $\rho \propto \dfrac{1}{V} \Rightarrow$

$$\frac{\rho'}{\rho} = \frac{V}{V'} = \frac{V}{V+\Delta V} = \frac{V}{V+\gamma V\Delta\theta} = \frac{1}{1+\gamma\,\Delta\theta}$$

$$\Rightarrow \rho' = \frac{\rho}{1+\gamma\,\Delta\theta} = \rho(1+\gamma\,\Delta\theta)^{-1} = \rho(1-\gamma\,\Delta\theta)$$

Expansion of Gases

Gases have no definite shape, therefore gases have only volume expansion. Since the expansion of container is negligible in comparison to the gases, therefore gases have only real expansion.

(1) **Coefficient of volume expansion :** At constant pressure, the unit volume of a given mass of a gas, increases with 1°C rise of temperature, is called coefficient of volume expansion.

$$\alpha = \frac{\Delta V}{V_0} \times \frac{1}{\Delta\theta} \quad \Rightarrow \text{Final volume } V' = V(1+\alpha\Delta\theta)$$

(2) **Coefficient of pressure expansion :** $\beta = \dfrac{\Delta P}{P} \times \dfrac{1}{\Delta\theta}$

$\therefore$ Final pressure $P' = P(1+\beta\Delta\theta)$

For an ideal gas, coefficient of volume expansion is equal to the coefficient of pressure expansion *i.e.* $\alpha = \beta = \dfrac{1}{273}\,°C^{-1}$

Heat

(1) The form of energy which is exchanged among various bodies or system on account of temperature difference is defined as heat.

(2) We can change the temperature of a body by giving heat (temperature rises) or by removing heat (temperature falls) from body.

(3) The amount of heat (Q) is given to a body depends upon it's mass (m), change in it's temperature ($\Delta\theta\,° = \Delta\theta$) and nature of material *i.e.* $Q = m.c.\Delta\theta$; where c = specific heat of material.

(4) Heat is a scalar quantity. It's units are *joule, erg, cal, kcal etc.*

(5) The calorie (*cal*) is defined as the amount of heat required to raise the temperature of 1 *gm* of water from 14.5°C to 15.5°C.

Also 1 *kcal* = 1000 *cal* = 4186 J and 1 *cal* = 4.18 J

(6) **British Thermal Unit (BTU) :** One BTU is the quantity of heat required to raise the temperature of one pound ($1\,lb$) of water from 63°F to 64°F

1 *BTU* = 778 *ft. lb* = 252 *cal* = 1055 J

(7) In solids thermal energy is present in the form of kinetic energy, in liquids, in the form of translatory energy of molecules. In gas it is due to the random motion of molecules.

(8) Heat always flows from a body of higher temperature to lower temperature till their temperature becomes equal (Thermal equilibrium).

(9) The heat required for a given temperature increase depends only on how many atoms the sample contains, not on the mass of an individual atom.

Specific Heat

When a body is heated it's temperature rises (except during a change in phase).

(1) **Gram specific heat :** The amount of heat energy required to raise the temperature of unit mass of a body through 1°C (or K) is called specific heat of the material of the body.

If Q heat changes the temperature of mass m by $\Delta\theta$ then specific heat $c = \dfrac{Q}{m\Delta\theta}$

(i) Units : *Calorie/gm* × °C (practical), *J/kg* × K (S.I.) Dimension : $[L^2T^{-2}\theta^{-1}]$

(ii) For an infinitesimal temperature change $d\theta$ and corresponding quantity of heat dQ.

Specific heat $c = \dfrac{1}{m}.\dfrac{dQ}{d\theta}$

(2) **Molar specific heat :** Molar specific heat of a substance is defined as the amount of heat required to raise the temperature of one gram mole of the substance through a unit degree it is represented by (capital) C.

Molar specific heat (C) $= M \times \text{Gram specific heat } (c)$

(M = Molecular mass of substance)

$$C = M\frac{Q}{m\Delta\theta} = \frac{1}{\mu}\frac{Q}{\Delta\theta} \quad \left(\text{where, Number of moles } \mu = \frac{m}{M}\right)$$

Units : *calorie/mole* × °C (practical); *J/mole* × *kelvin* (S.I.) Dimension : $[ML^2T^{-2}\theta^{-1}]$

Specific Heat of Solids

When a solid is heated through a small range of temperature, its volume remains more or less constant. Therefore specific heat of a solid may be called its specific heat at constant volume C_v.

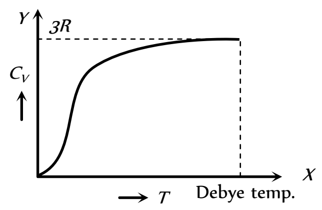


Fig. 12.13

(1) From the graph it is clear that at T = 0, C_v tends to zero

(2) With rise in temperature, C_v increases and at a particular temperature (called Debey's temperature) it becomes constant = $3R$ = 6 *cal/mole* × *kelvin* = 25 *J/mole* × *kelvin*

(3) For most of the solids, Debye temperature is close to room temperature.

(4) **Dulong and Petit law :** Average molar specific heat of all metals at room temperature is constant, being nearly equal to $3R$ = 6 *cal. mole*⋅ *K* = 25 *J mole*⋅ *K*, where *R* is gas constant for one mole of the gas. This statement is known as Dulong and Petit law.

(5) **Debey's law :** It was observed that at very low temperature molar specific heat $\propto T^3$ (exception are *Sn*, *Pb* and *Pt*)

(6) **Specific heat of ice :** In C.G.S. $c_{\text{ice}} = 0.5 \dfrac{cal}{gm \times °C}$

In S.I. $c_i = 500 \dfrac{cal}{kg \times °C} = 2100 \dfrac{Joule}{kg \times °C}$.

Table 12.5 : Specific heat of some solids at room temperature and atmospheric pressure

Substance	Specific heat ($J\text{-}kg^{-1}\ K^{-1}$)	Molar specific heat ($J\text{-}g\ mole^{-1}\ K^{-1}$)
Aluminium	900.0	24.4
Copper	386.4	24.5
Silver	236.1	25.5
Lead	127.7	26.5
Tungsten	134.4	24.9

Specific Heat of Liquid (Water)

(1) Among all known solids and liquids specific heat of water is maximum *i.e.* water takes more time to heat and more time to cool *w.r.t.* other solids and liquids.

(2) It is observed that by increasing temperature, initially specific heat of water goes on decreasing, becomes minimum at 37°*C* and then it start increasing. Specific heat of water is –

$$\frac{1\,cal}{gm \times °C} = 1000 \frac{cal}{kg \times °C} = 4200 \frac{J}{kg \times °C}$$

(This value is obtained between the temperature 14.5°*C* to 15.5°*C*)

(3) The variation of specific heat with temperature for water is shown in the figure. Usually this temperature dependence of specific heat is neglected.

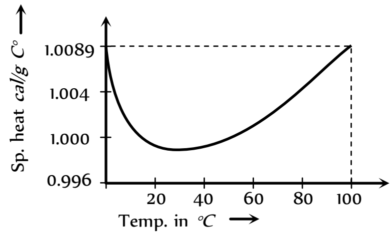


Fig. 12.14

(4) As specific heat of water is very large; by absorbing or releasing large amount of heat its temperature changes by small amount. This is why, it is used in hot water bottles or as coolant in radiators.

Specific Heat of Gases

(1) In case of gases, heat energy supplied to a gas is spent not only in raising the temperature of the gas but also in expansion of gas against atmospheric pressure.

(2) Hence specific heat of a gas, which is the amount of heat energy required to raise the temperature of one gram of gas through a unit degree shall not have a single or unique value.

(3) If the gas is compressed suddenly and no heat is supplied from outside *i.e.* ΔQ = 0, but the temperature of the gas raises on the account of compression.

$$\Rightarrow \quad c = \frac{Q}{m(\Delta\theta)} = \frac{0}{m\Delta\theta} = 0$$

(4) If the gas is heated and allowed to expand at such a rate that rise in temperature due to heat supplied is exactly equal to fall in temperature due to expansion of the gas. *i.e.* $\Delta\theta$ = 0

$$\Rightarrow \quad c = \frac{Q}{m(\Delta\theta)} = \frac{Q}{0} = \infty$$

(5) If rate of expansion of the gas were slow, the fall in temperature of the gas due to expansion would be smaller than the rise in temperature of the gas due to heat supplied. Therefore, there will be some net rise in temperature of the gas *i.e.* ΔT will be positive.

$$\Rightarrow \quad c = \frac{Q}{m(\Delta\theta)} = \text{ Positive}$$

(6) If the gas were to expand very fast, fall of temperature of gas due to expansion would be greater than rise in temperature due to heat supplied. Therefore, there will be some net fall in temperature of the gas *i.e.* $\Delta\theta$ will be negative.

$$\Rightarrow \quad c = \frac{Q}{m(-\Delta\theta)} = \text{Negative}$$

Hence the specific heat of gas can have any positive value ranging from zero to infinity. Further it can even be negative. The exact value depends upon the mode of heating the gas. Out of many values of specific heat of a gas, two are of special significance, namely C_v and C_p, in the chapter "Kinetic theory of gases" we will discussed this topic in detail.

Specific heat of steam : $c_{\text{steam}} = 0.47\, cal/gm \times °C$

Phase Change and Latent Heat

(1) **Phase :** We use the term phase to describe a specific state of matter, such as solid, liquid or gas. A transition from one phase to another is called a phase change.

(i) For any given pressure a phase change takes place at a definite temperature, usually accompanied by absorption or emission of heat and a change of volume and density.

(ii) In phase change ice at $0°C$ melts into water at $0°C$. Water at $100°C$ boils to form steam at $100°C$.

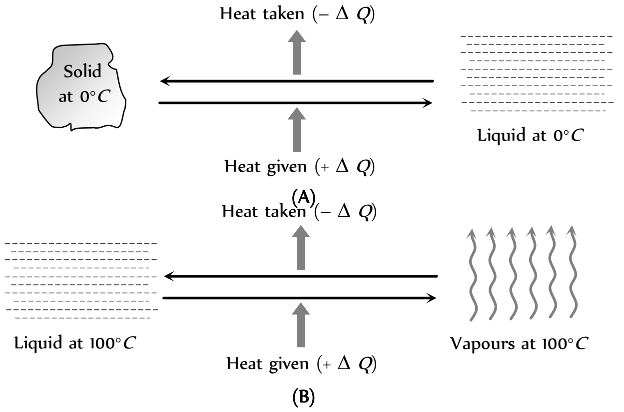


Fig. 12.15

(iii) In solids, the forces between the molecules are large and the molecules are almost fixed in their positions inside the solid. In a liquid, the forces between the molecules are weaker and the molecules may move freely inside the volume of the liquid. However, they are not able to come out of the surface. In vapours or gases, the intermolecular forces are almost negligible and the molecules may move freely anywhere in the container. When a solid melts, its molecules move apart against the strong molecular attraction. This needs energy which must be supplied from outside. Thus, the internal energy of a given body is larger in liquid phase than in solid phase. Similarly, the internal energy of a given body in vapour phase is larger than that in liquid phase.

(iv) In case of change of state if the molecules come closer, energy is released and if the molecules move apart, energy is absorbed.

(2) **Latent heat :** The amount of heat required to change the state of the mass m of the substance is written as : $Q = mL$, where L is the latent heat. Latent heat is also called as Heat of Transformation. It's unit is *cal/gm* or *J/kg* and Dimension: $[L^2T^{-2}]$

(i) **Latent heat of fusion :** The latent heat of fusion is the heat energy required to change 1 *kg* of the material in its solid state at its melting point to 1 *kg* of the material in its liquid state. It is also the amount of heat energy released when at melting point 1 *kg* of liquid changes to 1 *kg* of solid. For water at its normal freezing temperature or melting point ($0°C$), the latent heat of fusion (or latent heat of ice) is

$$L_F = L_{\text{ice}} \approx 80\, cal/gm \approx 60\, kJ/mol \approx 336\, kilo\, joule/kg$$

(ii) **Latent heat of vaporisation :** The latent heat of vaporisation is the heat energy required to change 1 *kg* of the material in its liquid state at its boiling point to 1 *kg* of the material in its gaseous state. It is also the amount of heat energy released when 1 *kg* of vapour changes into 1 *kg* of liquid. For water at its normal boiling point or condensation temperature ($100°C$), the latent heat of vaporisation (latent heat of steam) is

$$L_V = L_{\text{steam}} \approx 540\, cal/gm \approx 40.8\, kJ/mol \approx 2260\, kilo\, joule/kg$$

(iii) Latent heat of vaporisation is more than the latent heat of fusion. This is because when a substance gets converted from liquid to vapour, there is a large increase in volume. Hence more amount of heat is required. But when a solid gets converted to a liquid, then the increase in volume is negligible. Hence very less amount of heat is required. So, latent heat of vaporisation is more than the latent heat of fusion.

Thermal Capacity and Water Equivalent

(1) **Thermal capacity :** It is defined as the amount of heat required to raise the temperature of the whole body (mass m) through $0°C$ or $1K$.

Thermal capacity $= mc = \mu C = \dfrac{Q}{\Delta\theta}$

The value of thermal capacity of a body depends upon the nature of the body and its mass.

Dimension : $[ML^2T^{-2}\theta^{-1}]$, Unit : *cal/°C* (practical) *Joule/k* (S.I.)

(2) **Water Equivalent :** Water equivalent of a body is defined as the mass of water which would absorb or evolve the same amount of heat as is done by the body in rising or falling through the same range of temperature. It is represented by W.

If m = Mass of the body, c = Specific heat of body, $\Delta\theta$ = Rise in temperature.

Then heat given to body $\Delta Q = mc\Delta\theta$ ..... (i)

If same amount of heat is given to W *gm* of water and its temperature also rises by $\Delta\theta$. Then

heat given to water $Q = W \times 1 \times \Delta\theta$ ... (ii) [As $c_{\text{water}} = 1$]

From equation (i) and (ii) $\Delta Q = mc\Delta\theta = W \times 1 \times \Delta\theta$

$\Rightarrow$ Water equivalent (W) = *mc gm*

(i) Unit : *Kg* (S.I.) Dimension : $[ML^0T^0]$

(ii) Unit of thermal capacity is *J/kg* while unit of water equivalent is *kg*.

(iii) Thermal capacity of the body and its water equivalent are numerically equal.

(iv) If thermal capacity of a body is expressed in terms of mass of water it is called water-equivalent of the body.

Some Important Terms

(1) **Evaporation :** Vaporisation occurring from the free surface of a liquid is called evaporation. Evaporation is the escape of molecules from the surface of a liquid. This process takes place at all temperatures and increases with the increase of temperature. Evaporation leads to cooling because the faster molecules escape and, therefore, the average kinetic energy of the molecules of the liquid (and hence the temperature) decreases.

Evaporation cools hot water produced by power plants

Fig. 12.16

(2) **Melting (or fusion)/freezing (or solidification) :** The phase change of solid to liquid is called melting or fusion. The reverse phenomenon is called freezing or solidification.

Fig. 12.17

When pressure is applied on ice, it melts. As soon as the pressure is removed, it freezes again. This phenomenon is called **regelation**.

(3) **Vaporisation/liquefication (condensation) :** The phase change from liquid to vapour is called vaporisation. The reverse transition is called liquefication or condensation.

(4) **Sublimation :** Sublimation is the conversion of a solid directly into vapours. Sublimation takes place when boiling point is less than the melting point. A block of ice sublimates into vapours on the surface of moon because of very very low pressure on its surface. Heat required to change unit mass of solid directly into vapours at a given temperature is called heat of sublimation at that temperature.

(5) **Hoar frost :** Direct conversion of vapours into solid is called hoar frost. This process is just reverse of the process of sublimation, *e.g.*, formation of snow by freezing of clouds.


Fig. 12.18

(6) **Vapour pressure :** When the space above a liquid is closed, it soon becomes saturated with vapour and a dynamic equilibrium is established. The pressure exerted by this vapour is called Saturated Vapour Pressure (S.V.P.) whose value depends only on the temperature – it is independent of any external pressure. If the volume of the space is reduced, some vapour liquefies, but the pressure is unchanged.

A saturated vapour does not obey the gas law whereas the unsaturated vapour obeys them fairly well. However, a vapour differs from a gas in that the former can be liquefied by pressure alone, whereas the latter cannot be liquefied unless it is first cooled.

(7) **Boiling :** As the temperature of a liquid is increased, the rate of evaporation also increases. A stage is reached when bubbles of vapour start forming in the body of the liquid which rise to the surface and escape. A liquid boils at a temperature at which the S.V.P. is equal to the external pressure.


Fig. 12.19

It is a fast process. The boiling point changes on mixing impurities.

(8) **Dew point :** It is that temperature at which the mass of water vapour present in a given volume of air is just sufficient to saturate it, *i.e.* the temperature at which the actual vapour pressure becomes equal to the saturated vapuor pressure.

(9) **Humidity :** Atmospheric air always contains some water vapour. The mass of water vapour per unit volume is called absolute humidity.

The ratio of the mass of water vapour (m) actually present in a given volume of air to the mass of water vapour (M) required to saturate the same volume at the same temperature is called the relative humidity (R.H.).

Generally, it is expressed as a percentage, *i.e.*, $\text{R.H.}(\%) = \frac{m}{M} \times 100(\%)$

R.H. May also be defined as the ratio of the actual vapour pressure (p) of water at the same temperature, *i.e.* $\text{R.H.}(\%) = \frac{p}{P} \times 100(\%)$

Thus R.H. may also be defined as

$$\text{R.H.}(\%) = \frac{\text{S.V.P. at dew point}}{\text{S.V. P. at given temperature}} \times 100$$

(10) **Variation of melting point with pressure :** For those substances with contract on melting (*e.g.* water and rubber), the melting point decreases with pressure. The reason is the pressure helps shrinking and hence melting. Most substances expand on melting. (*e.g.* max, sulpher *etc.*)

An increase of pressure opposes the melting of such substances and their melting point is raised.

(11) **Variation of latent heat with temperature and pressure :** The latent heat of vapourization of a substance varies with temperature and hence pressure because the boiling point depends on pressure. It increases as the temperature is decreased. For example, water at 1 *atm* boils at 100°*C* and has latent heat 2259 Jg^{-1} but at 0.5 *atm* it boils at 82°*C* and has latent heat 2310 Jg^{-1}

The latent heat of fusion shows similar but less pronounced pressure dependence.

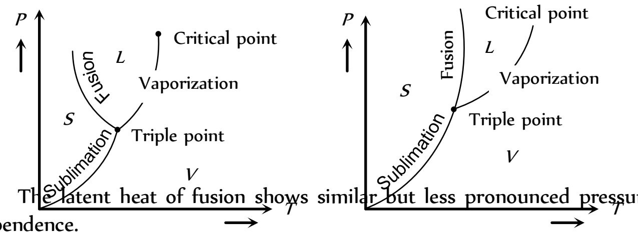


Fig. 12.20

The figures show the *P-T* graphs for (*a*) a substance (*e.g.*, water) which contracts on melting an (*b*) a substance (*e.g.* wax) which expands on melting. The *P-T* graph consists of three curves.

(i) Sublimation curve which connects points at which vapour (*V*) and solid (*S*) exist in equilibrium.

(ii) Vapourization curve which shows vapour and liquid (*L*) existing in equilibrium.

(iii) Fusion curve which shows liquid and solid existing in equilibrium.

The three curves meet at a single point which is called the triple point. It is that unique temperature-pressure point for a substance at which all the three phases exist in equilibrium.

(12) **Freezing mixture :** If salt is added to ice, then the temperature of mixture drops down to less than 0°*C*. This is so because, some ice melts down to cool the salt to 0°*C*. As a result, salt gets dissolved in the water formed and saturated solution of salt is obtained; but the ice point (freeing point) of the solution formed is always less than that of pure water. So, ice cannot be in the solid state with the salt solution at 0°*C*. The ice which is in contact with the solution, starts melting and it absorbs the required latent heat from the mixture, so the temperature of mixture falls down.

Joule's Law (Heat and Mechanical Work)

Whenever heat is converted into mechanical work or mechanical work is converted into heat, then the ratio of work done to heat produced always remains constant. *i.e.* $W \propto Q$ or $\frac{W}{Q} = J$

This is Joule's law and J is called mechanical equivalent of heat.

(1) From $W = JQ$ if $Q = 1$ then $J = W$. Hence the amount of work done necessary to produce unit amount of heat is defined as the mechanical equivalent of heat.

(2) J is neither a constant, nor a physical quantity rather it is a conversion factor which used to convert *Joule* or *erg* into *calorie* or *kilo calories* vice-versa.

(3) Value of $J = 4.2\dfrac{Joule}{cal} = 4.2\times10^7\dfrac{erg}{cal}$

$$= 4.2\times10^3\frac{Joule}{kcal}.$$

(4) When water in a stream falls from height h, then its potential energy is converted into heat and temperature of water rises slightly.

From $W = JQ \Rightarrow mgh = J(mc\,\Delta\theta)$

[where m = Mass of water, c = Specific heat of water, $\Delta\theta$ = temperature rise]

$\Rightarrow$ Rise in temperature $\Delta\theta = \dfrac{gh}{Jc}\,°C$

(5) The kinetic energy of a bullet fired from a gun gets converted into heat on striking the target. By this heat the temperature of bullet increases by $\Delta\theta$.

From $W = JQ \Rightarrow \dfrac{1}{2}mv^2 = J(m\,s\,\Delta\theta)$

[where m = Mass of the bullet, v = Velocity of the bullet, c = Specific heat of the bullet]

$\Rightarrow$ Rise in temperature $\Delta t = \dfrac{v^2}{2Jc}\,°C$

If the temperature of bullet rises upto the melting point of the bullet and bullet melts then.

From $W = J(Q_{\text{Temperature change}} + Q_{\text{Phase change}})$

$\Rightarrow \dfrac{1}{2}mv^2 = J(mc\,\Delta\theta + mL)$; L = Latent heat of bullet

$\Rightarrow$ Rise in temperature $\Delta\theta = \left[\dfrac{\left(\dfrac{v^2}{2J} - L\right)}{c}\right]°C$

(6) If m *kg* ice-block falls down through some height (h) and melts partially (m' *kg*) then its potential energy gets converted into heat of melting.

From $W = JQ \Rightarrow mgh = J\,m'L \Rightarrow h = \dfrac{m'}{m}\left(\dfrac{JL}{g}\right)$

If ice-block melts completely then $m' = m \Rightarrow h = \dfrac{JL}{g}\,meter$

Principle of Calorimetry

Calorimetry means 'measuring heat'.

When two bodies (one being solid and other liquid or both being liquid) at different temperatures are mixed, heat will be transferred from body at higher temperature to a body at lower temperature till both acquire same temperature. The body at higher temperature releases heat while body at lower temperature absorbs it, so that

Heat lost = Heat gained

i.e. principle of calorimetry represents the law of conservation of heat energy.

(1) Temperature of mixture (θ_{mix}) is always ≥ lower temperature (θ_L) and ≤ higher temperature (θ_H), *i.e.*, $\theta_L \le \theta_{mix} \le \theta_H$.

It means the temperature of mixture can never be lesser than lower temperatures (as a body cannot be cooled below the temperature of cooling body) and greater than higher temperature (as a body cannot be heated above the temperature of heating body). Furthermore usually rise in temperature of one body is not equal to the fall in temperature of the other body though heat gained by one body is equal to the heat lost by the other.

(2) **Mixing of two substances when temperature changes only** : It means no phase change. Suppose two substances having masses m_1 and m_2, gram specific heat c_1 and c_2, temperatures θ_1 and θ_2 $(\theta_1 > \theta_2)$ are mixed together such that temperature of mixture at equilibrium is θ_{mix}

Hence, Heat lost = Heat gained

$$\Rightarrow m_1c_1(\theta_1 - \theta_{mix}) = m_2c_2(\theta_{mix} - \theta_2) \Rightarrow$$

$$\theta_{mix} = \frac{m_1c_1\theta_1 + m_2c_2\theta_2}{m_1c_1 + m_2c_2}$$

Table 12.6 : Temperature of mixture in different cases

Condition	Temperature of mixture
If bodies are of same material *i.e.* $c_1 = c_2$	$\theta_{mix} = \dfrac{m_1\theta_1 + m_2\theta_2}{m_1 + m_2}$
If bodies are of same mass $m_1 = m_2$	$\theta_{mix} = \dfrac{\theta_1c_1 + \theta_2c_2}{c_1 + c_2}$
If $m_1 = m_2$ and $c_1 = c_2$	$\theta_{mix} = \dfrac{\theta_1 + \theta_2}{2}$

(3) **Mixing of two substances when temperature and phase both changes or only phase changes:** A very common example for this category is ice-water mixing.

Suppose water at temperature $\theta_W\,°C$ is mixed with ice at $0°C$, first ice will melt and then it's temperature rises to attain thermal equilibrium. Hence; Heat given = Heat taken

$$\Rightarrow m_WC_W(\theta_W - \theta_{mix}) = m_iL_i + m_iC_W(\theta_{mix} - 0°)$$

$$\Rightarrow \theta_{mix} = \frac{m_W\theta_W - \dfrac{m_iL_i}{C_W}}{m_W + m_i}$$

(i) If $m_W = m_i$ then $\theta_{mix} = \dfrac{\theta_W - \dfrac{L_i}{C_W}}{2}$

(ii) By using this formulae if $\theta_{mix} < \theta_i$ then take $\theta_{mix} = 0°C$

Heating Curve

If to a given mass (m) of a solid, heat is supplied at constant rate P and a graph is plotted between temperature and time, the graph is as shown in figure and is called heating curve. From this curve it is clear that

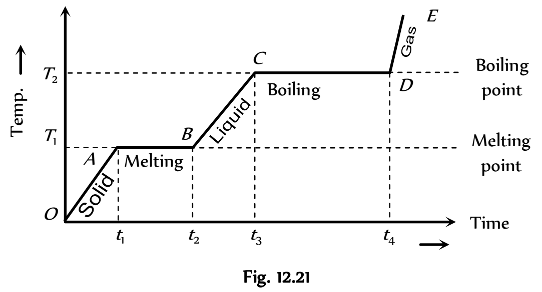


Fig. 12.21

(1) In the region *OA* temperature of solid is changing with time so, $Q = mc_S \Delta T \Rightarrow P\,\Delta t = mc_S \Delta T$ [as $Q = P\Delta t$]

But as $(\Delta T/\Delta t)$ is the slope of temperature-time curve

$$c_s \propto \frac{1}{\text{Slope of line } OA}$$

i.e. specific heat (or thermal capacity) is inversely proportional to the slope of temperature-time curve.

(2) In the region *AB* temperature is constant, so it represents change of state, *i.e.*, melting of solid with melting point T_1. At *A* melting starts and at *B* all solid is converted into liquid. So between *A* and *B* substance is partly solid and partly liquid. If L_F is the latent heat of fusion. $Q = mL_F$ or

$$L_F = \frac{P(t_2 - t_1)}{m} \quad [\text{as } Q = P(t_2 - t_1)]$$

or $L_F \propto$ length of line *AB*

i.e. Latent heat of fusion is proportional to the length of line of zero slope. [In this region specific heat $\propto \frac{1}{\tan 0} = \infty$]

(3) In the region *BC* temperature of liquid increases so specific heat (or thermal capacity) of liquid will be inversely proportional to the slope of line *BC*

i.e., $c_L \propto \frac{1}{\text{Slope of line } BC}$

(4) In the region *CD* temperature is constant, so it represents the change of state, *i.e.*, boiling with boiling point T_2. At *C* all substance is in liquid state while at *D* in vapour state and between *C* and *D* partly liquid and partly gas. The length of line *CD* is proportional to latent heat of vaporisation

i.e., $L_V \propto$ Length of line *CD* [In this region specific heat $\propto \frac{1}{\tan 0} = \infty$]

(5) The line *DE* represents gaseous state of substance with its temperature increasing linearly with time. The reciprocal of slope of line will be proportional to specific heat or thermal capacity of substance in vapour state.

Tips & Tricks

✍ After snow falls, the temperature of the atmosphere becomes very low. This is because the snow absorbs the heat from the atmosphere to melt down. So, in the mountains, when snow falls, one does not feel too cold, but when ice melts, he feels too cold.

✍ There is more shivering effect of **ice-cream** on teeth as compared to that of water (obtained from ice). This is because, when ice-cream melts down, it absorbs large amount of heat from teeth.

✍ Branch of physics dealing with production and measurement of temperatures close to $0K$ is known as cryogenics while that dealing with the measurement of very high temperature is called as pyrometry.

✍ It is more painful to get burnt by steam rather than by boiling water at same temperature. This is so because when steam at 100°*C* gets converted to water at 100°*C*, then it gives out 536 *calories* of heat. So, it is clear that steam at 100°*C* has more heat than water at 100°*C* (*i.e.*, boiling of water).

✍ A solid and hollow sphere of same radius and material, heated to the same temperature then expansion of both will be equal because thermal expansion of isotropic solids is similar to true photographic enlargement. It means the expansion of cavity is same as if it has been a solid body of the same material. But if same heat is given to the two spheres, due to lesser mass, rise in temperature of hollow sphere will be more $\left\{\text{As}\left(\Delta\theta = \frac{\Delta Q}{mc}\right)\right\}$.

Hence its expansion will be more.

✍ Specific heat of a substance can also be negative. Negative specific heat means that in order to raise the temperature, a certain quantity of heat is to be withdrawn from the body.

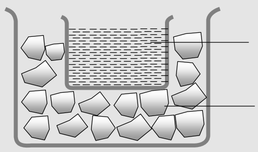
e.g. Specific heat of saturated vapours.

✍ Specific heat for hydrogen is maximum $(3.5\ cal/gm \times {}^oC)$ and it is minimum for radon and actinium $(\simeq 0.022\,cal/gm \times {}^\circ C)$.

✍ The minimum possible temperature is 0 *K*.

✍ Amount of steam at 100°*C* required to just melt *m gm* of ice at 0°*C* is *m*/8 *gm*.

✍ If we put the beaker containing water in melting ice, the water in the beaker will cool to 0°*C* but will never freeze.



✍ A pressure in excess of 25 *atm* is required to make helium solidfy. At 1 *atm* pressure, helium remains a liquid down to absolute zero.

✍ Boiling temperature of water, if pressure is different from normal pressure is $t_{boiling}$ = [100°*C* – (760 – *P* in *mm*) × 0.037]°*C*

✍ **Confusing S.I. and C.G.S. units**

It is advised to do questions on calorimetry in C.G.S. as calculations becomes simple. If the final answer is in joules, then convert cal into joules.

✍ Invar and quartz have very small values of co-efficient of linear expansion.

✍ In S.I. nomenclature " degree" is not used with the kelvin scale; *e.g.* $273°K$ is wrong while $273\ K$ is correct to write.

✍ Magnetic thermometer is recommended for measuring very low temperature ($2K$).

✍ The most sensitive thermometer is gas thermometer.

✍ Dew formation is more probable on a cloudiness calm night.

✍ In winters, generally fog disappear before noon. Because, the atmosphere warms up and tends to be unsaturated. The condensed vapours reevaporates and the fog disappears.

✍ Standardisation of thermometer is obtained with gas thermometer. Because coefficient of expansion of gas is very large.

✍ Dogs hang their tongues in order to expose a surface to the air for evaporation and hence, cooling. They do not sweat.

OT Ordinary Thinking

Objective Questions

Thermometry

1. On the Celsius scale the absolute zero of temperature is at [CBSE PMT 1994]

(a) $0°C$ (b) $-32°C$

(c) $100°C$ (d) $-273.15°C$

2. Oxygen boils at $-183°C$. This temperature is approximately [CPMT 1992]

(a) $215°F$ (b) $-297°F$

(c) $329°F$ (d) $361°F$

3. Recently, the phenomenon of superconductivity has been observed at $95\ K$. This temperature is nearly equal to [CPMT 1990]

(a) $-288°F$ (b) $-146°F$

(c) $-368°F$ (d) $+178°F$

4. The temperature of a substance increases by $27°C$. On the Kelvin scale this increase is equal to [CPMT 1993]

(a) $300\ K$ (b) $2.46\ K$

(c) $27\ K$ (d) $7\ K$

5. The resistance of a resistance thermometer has values 2.71 and 3.70 *ohm* at $10°C$ and $100°C$. The temperature at which the resistance is 3.26 *ohm* is [CPMT 1994]

(a) $40°C$ (b) $50°C$

(c) $60°C$ (d) $70°C$

6. No other thermometer is as suitable as a platinum resistance thermometer to measure temperature in the entire range of [MNR 1993]

(a) $0°C$ to $100°C$ (b) $100°C$ to $1500°C$

(c) $-50°C$ to $+350°C$ (d) $-200°C$ to $600°C$

7. The temperature of the sun is measured with [Pb. PMT 1998; CPMT 1998; Pb. PET 1997, 2001]
(a) Platinum thermometer
(b) Gas thermometer
(c) Pyrometer
(d) Vapour pressure thermometer

8. Absolute temperature can be calculated by [AFMC 1994]
(a) Mean square velocity (b) Motion of the molecule
(c) Both (a) and (b) (d) None of the above

9. Thermoelectric thermometer is based on [CPMT 1993, 95; AFMC 1998]
(a) Photoelectric effect (b) Seeback effect
(c) Compton effect (d) Joule effect

10. Maximum density of H_2O is at the temperature [CPMT 1996; Pb. PMT 1996]
(a) $32°F$ (b) $39.2°F$
(c) $42°F$ (d) $4°F$

11. The study of physical phenomenon at low temperatures (below liquid nitrogen temperature) is called [CPMT 1992]
(a) Refrigeration (b) Radiation
(c) Cryogenics (d) Pyrometry

12. 'Stem Correction' in platinum resistance thermometers are eliminated by the use of [AIIMS 1998]
(a) Cells (b) Electrodes
(c) Compensating leads (d) None of the above

13. The absolute zero is the temperature at which [AIIMS 1998]
(a) Water freezes
(b) All substances exist in solid state
(c) Molecular motion ceases
(d) None of the above

14. Absolute scale of temperature is reproduced in the laboratory by making use of a [SCRA 1998]
(a) Radiation pyrometer
(b) Platinum resistance thermometer
(c) Constant volume helium gas thermometer
(d) Constant pressure ideal gas thermometer

15. Absolute zero (0 K) is that temperature at which [AFMC 1993]
(a) Matter ceases to exist
(b) Ice melts and water freezes
(c) Volume and pressure of a gas becomes zero
(d) None of these

16. On which of the following scales of temperature, the temperature is never negative [EAMCET 1997]
(a) Celsius (b) Fahrenheit
(c) Reaumur (d) Kelvin

17. The temperature on Celsius scale is $25°C$. What is the corresponding temperature on the Fahrenheit scale [AFMC 2001]
(a) $40°F$ (b) $77°F$
(c) $50°F$ (d) $45°F$

18. One quality of a thermometer is that its heat capacity should be small. If P is a mercury thermometer, Q is a resistance thermometer and R thermocouple type then [CPMT 1997]
(a) P is best, R worst (b) R is best, P worst
(c) R is best, Q worst (d) P is best, Q worst

19. Two thermometers are used to record the temperature of a room. If the bulb of one is wrapped in wet hanky [AFMC 1997]
(a) The temperature recorded by both will be same
(b) The temperature recorded by wet-bulb thermometer will be greater than that recorded by the other
(c) The temperature recorded by dry-bulb thermometer will be greater than that recorded by the other
(d) None of the above

20. The temperature of a body on Kelvin scale is found to be x K. When it is measured by Fahrenheit thermometer, it is found to be $x°F$, then the value of x is [UPSEAT 2000; Pb. CET 2004]
(a) 40 (b) 313
(c) 574.25 (d) 301.25

21. A centigrade and a Fahrenheit thermometer are dipped in boiling water. The water temperature is lowered until the Fahrenheit thermometer registers 140°. What is the fall in temperature as registered by the Centigrade thermometer [CBSE PMT 1992; AIIMS 1998]
(a) 30° (b) 40°
(c) 60° (d) 80°

22. At what temperature the centigrade (Celsius) and Fahrenheit, readings are the same [RPMT 1997, 99, 2003; BHU 1997; MNR 1992; DPMT 1998; CPMT 1995; UPSEAT 1999; KCET 2000]
(a) – 40° (b) + 40°
(c) 36.6° (d) – 37°

23. Standardisation of thermometers is obtained with [CPMT 1996]
(a) Jolly's thermometer
(b) Platinum resistance thermometer
(c) Thermocouple thermometer
(d) Gas thermometer

24. The gas thermometers are more sensitive than liquid thermometers because [CPMT 1993]
(a) Gases expand more than liquids
(b) Gases are easily obtained
(c) Gases are much lighter
(d) Gases do not easily change their states

25. Mercury thermometers can be used to measure temperatures upto [CBSE PMT 1992, 96; BHU 1998; UPSEAT 1998]
(a) $100°C$ (b) $212°C$
(c) $360°C$ (d) $500°C$

26. A constant volume gas thermometer shows pressure reading of $50 cm$ and $90 cm$ of mercury at $0°C$ and $100°C$ respectively. When the pressure reading is 60 cm of mercury, the temperature is [MNR 1991; UPSEAT 2000; Pb. CET 2004]
(a) $25°C$ (b) $40°C$

(c) $15^{\circ}C$ (d) $12.5^{\circ}C$

27. Mercury boils at $367^{\circ}C$. However, mercury thermometers are made such that they can measure temperature up to $500^{\circ}C$. This is done by [CPMT 2004]

(a) Maintaining vacuum above mercury column in the stem of the thermometer

(b) Filling nitrogen gas at high pressure above the mercury column

(c) Filling nitrogen gas at low pressure above the mercury level

(d) Filling oxygen gas at high pressure above the mercury column

28. A device used to measure very high temperature is [KCET 1998]

(a) Pyrometer (b) Thermometer

(c) Bolometer (d) Calorimeter

29. The absolute zero temperature in Fahrenheit scale is [DCE 1996]

(a) $-273^{\circ}F$ (b) $-32^{\circ}F$

(c) $-460^{\circ}F$ (d) $-132^{\circ}F$

30. A constant pressure air thermometer gave a reading of 47.5 units of volume when immersed in ice cold water, and 67 units in a boiling liquids. The boiling point of the liquid will be

(a) $135^{\circ}C$ (b) $125^{\circ}C$

(c) $112^{\circ}C$ (d) $100^{\circ}C$

31. If a thermometer reads freezing point of water as $20^{\circ}C$ and boiling point as $150^{\circ}C$, how much thermometer read when the actual temperature is $60^{\circ}C$ [AFMC 2004]

(a) $98^{\circ}C$ (b) $110^{\circ}C$

(c) $40^{\circ}C$ (d) $60^{\circ}C$

32. If temperature of an object is $140^{\circ}F$, then its temperature in centigrade is [RPMT 1999]

(a) $105^{\circ}C$ (b) $32^{\circ}C$

(c) $140^{\circ}C$ (d) $60^{\circ}C$

33. Of the following thermometers, the one which can be used for measuring a rapidly changing temperature is a [CPMT 1992]

(a) Thermocouple thermometer

(b) Gas thermometer

(c) Maximum resistance thermometer

(d) Vapour pressure thermometer

34. On centigrade scale the temperature of a body increases by 30 degrees. The increase in temperature on Fahrenheit scale is

(a) 50° (b) 40°

(c) 30° (d) 54°

35. The correct value of $0^{\circ}C$ on Kelvin scale will be [RPMT 1999]

(a) 273.15 K (b) 273.00 K

(c) 273.05 K (d) 273.63 K

Thermal Expansion

1. When a copper ball is heated, the largest percentage increase will occur in its [EAMCET 1992]

(a) Diameter (b) Area

(c) Volume (d) Density

2. A vertical column 50 cm long at $50^{\circ}C$ balances another column of same liquid 60 cm long at $100^{\circ}C$. The coefficient of absolute expansion of the liquid is [EAMCET 1990]

(a) $0.005/^{\circ}C$ (b) $0.0005/^{\circ}C$

(c) $0.002/^{\circ}C$ (d) $0.0002/^{\circ}C$

3. The apparent coefficient of expansion of a liquid when heated in a copper vessel is C and when heated in a silver vessel is S. If A is the linear coefficient of expansion of copper, then the linear coefficient of expansion of silver is [EAMCET 1991]

(a) $\frac{C+S-3A}{3}$ (b) $\frac{C+3A-S}{3}$

(c) $\frac{S+3A-C}{3}$ (d) $\frac{C+S+3A}{3}$

4. A uniform metal rod is used as a bar pendulum. If the room temperature rises by $10^{\circ}C$, and the coefficient of linear expansion of the metal of the rod is 2×10^{-6} per $^{\circ}C$, the period of the pendulum will have percentage increase of [NSEP 1992]

(a) -2×10^{-3} [AIIMS 1994] (b) -1×10^{-3}

(c) 2×10^{-3} (d) 1×10^{-3}

5. A bar of iron is 10 cm at $20^{\circ}C$. At $19^{\circ}C$ it will be (α of iron $= 11 \times 10^{-6}/^{\circ}C$) [EAMCET 1997]

(a) 11×10^{-6} cm longer (b) 11×10^{-6} cm shorter

(c) 11×10^{-5} cm shorter (d) 11×10^{-5} cm longer

6. When a rod is heated but prevented from expanding, the stress developed is independent of [EAMCET 1997]

(a) Material of the rod (b) Rise in temperature

(c) Length of rod (d) None of above

7. Expansion during heating [CBSE PMT 1994]

(a) Occurs only in solids

(b) Increases the weight of a material

(c) Decreases the density of a material

(d) Occurs at the same rate for all liquids and solids

8. On heating a liquid of coefficient of cubical expansion γ in a container having coefficient of linear expansion $\gamma / 3$, the level of liquid in the container will [EAMCET 1993] [UPSEAT 2005]

(a) Rise

(b) Fall

(c) Will remain almost stationary

(d) It is difficult to say

9. A pendulum clock keeps correct time at $0^{\circ}C$. Its mean coefficient of linear expansions is $\alpha / {}^{\circ}C$, then the loss in seconds per day by the clock if the temperature rises by $t^{\circ}C$ is

(a) $\frac{\frac{1}{2}\alpha t \times 864000}{1-\frac{\alpha t}{2}}$ (b) $\frac{1}{2}\alpha t \times 86400$

(c) $\dfrac{\frac{1}{2}\alpha t \times 86400}{\left(1-\frac{\alpha t}{2}\right)^2}$ (d) $\dfrac{\frac{1}{2}\alpha t \times 86400}{1+\frac{\alpha t}{2}}$

10. When a bimetallic strip is heated, it [CBSE PMT 1990]

(a) Does not bend at all

(b) Gets twisted in the form of an helix

(c) Bend in the form of an arc with the more expandable metal outside

(d) Bends in the form of an arc with the more expandable metal inside

11. A solid ball of metal has a concentric spherical cavity within it. If the ball is heated, the volume of the cavity will

[AFMC 1997; Orissa PMT 2004]

(a) Increase (b) Decrease

(c) Remain unaffected (d) None of these

12. A *litre* of alcohol weighs [AFMC 1994]

(a) Less in winter than in summer

(b) Less in summer than in winter

(c) Some both in summer and winter

(d) None of the above

13. 5 *litre* of benzene weighs [MNR 1996]

(a) More in summer than in winter

(b) More in winter than in summer

(c) Equal in winter and summer

(d) None of the above

14. Water has maximum density at [Pb. PMT 1997]

(a) $0°C$ (b) $32°F$

(c) $-4°C$ (d) $4°C$

15. At some temperature T, a bronze pin is a little large to fit into a hole drilled in a steel block. The change in temperature required for an exact fit is minimum when

[SCRA 1998]

(a) Only the block is heated

(b) Both block and pin are heated together

(c) Both block and pin are cooled together

(d) Only the pin is cooled

16. If the length of a cylinder on heating increases by 2%, the area of its base will increase by [CPMT 1993; BHU 1997]

(a) 0.5% (b) 2%

(c) 1% (d) 4%

17. The volume of a gas at $20°C$ is 100 cm^3 at normal pressure. If it is heated to $100°C$, its volume becomes 125 cm^3 at the same pressure, then volume coefficient of the gas at normal pressure is

(a) $0.0015/°C$ (b) $0.0045/°C$

(c) $0.0025/°C$ (d) $0.0033/°C$

18. The coefficient of superficial expansion of a solid is 2×10^{-5} /°C. It's coefficient of linear expansion is [KCET 1999]

(a) $4 \times 10^{-5}/°C$ (b) $3 \times 10^{-5}/°C$

(c) $2 \times 10^{-5}/°C$ (d) $1 \times 10^{-5}/°C$

19. Density of substance at $0°C$ is 10 *gm/cc* and at $100°C$, its density is 9.7 *gm/cc*. The coefficient of linear expansion of the substance will be

[BHU 1996; Pb. PMT 1999; DPMT 1998, 2003]

(a) 10^{2} (b) 10^{-2}

(c) 10^{-3} (d) 10^{-4}

20. Coefficient of real expansion of mercury is $0.18 \times 10^{-3}/°C$. If the density of mercury at $0°C$ is 13.6 *gm/cc*. its density at $473K$ is

(a) 13.11 *gm/cc* (b) 26.22 *gm/cc*

(c) 52.11 *gm/cc* (d) None of these

21. The real coefficient of volume expansion of glycerine is 0.000597 per$°C$ and linear coefficient of expansion of glass is 0.000009 per$°C$. Then the apparent volume coefficient of expansion of glycerine is

(a) 0.000558 per$°C$ (b) 0.00057 per$°C$

(c) 0.00027 per$°C$ (d) 0.00066 per$°C$

22. A beaker is completely filled with water at $4°C$. It will overflow if [EAMCET 199

(a) Heated above $4°C$

(b) Cooled below $4°C$

(c) Both heated and cooled above and below $4°C$ respectively

(d) None of the above

23. The volume of a metal sphere increases by 0.24% when its temperature is raised by $40°C$. The coefficient of linear expansion of the metal is .... $°C$ [Kerala PMT 2005]

(a) 2×10^{-5} (b) 6×10^{-5}

(c) 2.1×10^{-5} (d) 1.2×10^{-5}

24. Ratio among linear expansion coefficient (α), areal expansion coefficient (β) and volume expansion coefficient (γ) is

(a) 1 : 2 : 3 (b) 3 : 2 : 1

(c) 4 : 3 : 2 (d) None of these

25. If on heating liquid through $80°C$, the mass expelled is $(1/100)^{th}$ of mass still remaining, the coefficient of apparent expansion of liquid is [RPMT 2004]

(a) $1.25 \times 10^{-4}/°C$ (b) $12.5 \times 10^{-4}/°C$

(c) $1.25 \times 10^{-5}/°C$ (b) None of these

26. In cold countries, water pipes sometimes burst, because

(a) Pipe contracts

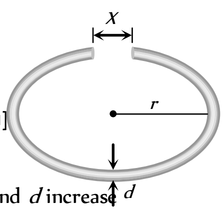
(b) Water expands on freezing

(c) When water freezes, pressure increases

(d) When water freezes, it takes heat from pipes

27. A cylindrical metal rod of length L_0 is shaped into a ring with a small gap as shown. On heating the system

[Pb. PET 2002; DPMT 2001]



(a) x decreases, r and d increase

(b) x and r increase, d decreases

(c) x, r and d all increase

(d) Data insufficient to arrive at a conclusion

28. The length of a metallic rod is $5m$ at $0°C$ and becomes 5.01 m, on heating upto $100°C$. The linear expansion of the metal will be

(a) 2.33×10^{-5} /°C (b) 6.0×10^{-5} /°C

(c) 4.0×10^{-5} /°C (d) 2.0×10^{-5} /°C

29. A metal rod of silver at 0°C is heated to 100°C. It's length is increased by 0.19 *cm*. Coefficient of cubical expansion of the silver rod is [UPSEAT 2001]

(a) 5.7×10^{-5}/°C (b) 0.63×10^{-5}/°C

(c) 1.9×10^{-5}/°C (d) 16.1×10^{-5}/°C

30. A brass disc fits simply in a hole of a steel plate. The disc from the hole can be loosened if the system [UPSEAT 2001]

(a) First heated then cooled (b) First cooled then heated

(c) Is heated (d) Is cooled

31. An iron bar of length 10 *m* is heated from 0°C to 100°C. If the coefficient of linear thermal expansion of iron is 10×10^{-6} /°C, the increase in the length of bar is [UPSEAT 2005]

(a) 0.5 *cm* (b) 1.0 *cm*

(c) 1.5 *cm* (d) 2.0 *cm*

32. If a cylinder of diameter 1.0 *cm* at 30°C is to be solid into a hole of diameter 0.9997 *cm* in a steel plate at the same temperature, then minimum required rise in the temperature of the plate is : (Coefficient of linear expansion of steel $= 12 \times 10^{-6} / °C$)

(a) 25°C (b) 35°C

(c) 45°C (d) 55°C

33. Surface of the lake is at 2°C. Find the temperature of the bottom of the lake [Orissa JEE 2002]

(a) 2°C (b) 3°C

(c) 4°C (d) 1°C

34. Two rods, one of aluminum and the other made of steel, having initial length l_1 and l_2 are connected together to form a single rod of length $l_1 + l_2$. The coefficients of linear expansion for aluminum and steel are α_a and α_s respectively. If the length of each rod increases by the same amount when their temperature are raised by t^oC, then find the ratio $\frac{l_1}{(l_1 + l_2)}$

(a) $\frac{\alpha_s}{\alpha_a}$ (b) $\frac{\alpha_a}{\alpha_s}$

(c) $\frac{\alpha_s}{(\alpha_a + \alpha_s)}$ (d) $\frac{\alpha_a}{(\alpha_a + \alpha_s)}$

Calorimetry

1. When vapour condenses into liquid [CPMT 1990]

(a) It absorbs heat (b) It liberates heat

(c) Its temperature increases (d) Its temperature decreases

2. At NTP water boils at 100°C. Deep down the mine, water will boil at a temperature [CPMT 1996]

(a) 100°C (b) > 100°C

(c) < 100°C (d) Will not boil at all

3. If specific heat of a substance is infinite, it means [AIIMS 1997]

(a) Heat is given out

(b) Heat is taken in

(c) No change in temperature takes place whether heat is taken in or given out

(d) All of the above

4. A gas in an airtight container is heated from 25°C to 90°C. The density of the gas will [BCECE 1997]

(a) Increase slightly (b) Increase considerably

(c) Remain the same (d) Decrease slightly

5. A quantity of heat required to change the unit mass of a solid substance, from solid state to liquid state, while the temperature remains constant, is known as [AIIMS 1998]

(a) Latent heat (b) Sublimation

(c) Hoar frost (d) Latent heat of fusion

6. The latent heat of vaporization of a substance is always [SCRA 1998]

(a) Greater than its latent heat of fusion

(b) Greater than its latent heat of sublimation

(c) Equal to its latent heat of sublimation

(d) Less than its latent heat of fusion

7. The factor not needed to calculate heat lost or gained when there is no change of state is [AFMC 1997; BHU 1997]

(a) Weight (b) Specific heat

(c) Relative density [EAMCET 2001] (d) Temperature change

8. 540 *g* of ice at 0°C is mixed with 540 *g* of water at 80°C. The final temperature of the mixture is [AFMC 1994]

(a) 0°C (b) 40°C

(c) 80°C (d) Less than 0°C

9. Water is used to cool radiators of engines, because [AFMC 2001]

(a) Of its lower density (b) It is easily available

(c) It is cheap (d) It has high specific heat

10. How much heat energy is gained when 5 *kg* of water at 20°C is brought to its boiling point

(Specific heat of water = 4.2 $kJ\ kg^{-1} c^{-1}$) [BHU 2001]

(a) 1680 *kJ* (b) 1700 *kJ*

[IIT-JEE (Screening) 2003] (c) 1720 *kJ* (d) 1740 *kJ*

11. Melting point of ice [CBSE PMT 1993]

(a) Increases with increasing pressure

(b) Decreases with increasing pressure

(c) Is independent of pressure

(d) Is proportional to pressure

12. Heat required to convert one gram of ice at 0°C into steam at 100°C is (given L_{steam} = 536 *cal/gm*) [Pb. PMT 1990]

(a) 100 *calorie* (b) 0.01 *kilocalorie*

(c) 716 *calorie* (d) 1 *kilocalorie*

13. 80 *gm* of water at 30°C are poured on a large block of ice at 0°C. The mass of ice that melts is [CBSE PMT 1989]

(a) 30 *gm* (b) 80 *gm*

(c) 1600 *gm* (d) 150 *gm*

14. The saturation vapour pressure of water at 100°C is [EAMCET 1997]

(a) 739 *mm* of mercury (b) 750 *mm* of mercury

(c) 760 *mm* of mercury (d) 712 *mm* of mercury

15. Two spheres made of same substance have diameters in the ratio 1 : 2. Their thermal capacities are in the ratio of [JIPMER 1999]

(a) 1 : 2 (b) 1 : 8
(c) 1 : 4 (d) 2 : 1

16. Work done in converting one gram of ice at –10°*C* into steam at 100°*C* is [MP PET/PMT 1988; EAMCET (Med.) 1995; MP PMT 2003]

(a) 3045 *J* (b) 6056 *J*
(c) 721 *J* (d) 616 *J*

17. If mass energy equivalence is taken into account, when water is cooled to form ice, the mass of water should [AIEEE 2002]

(a) Increase (b) Remain unchanged
(c) Decrease (d) First increase then decrease

18. Compared to a burn due to water at 100°*C*, a burn due to steam at 100°*C* is [KCET 1999; UPSEAT 1999]

(a) More dangerous (b) Less dangerous
(c) Equally dangerous (d) None of these

19. 50 *gm* of copper is heated to increase its temperature by 10°*C*. If the same quantity of heat is given to 10 *gm* of water, the rise in its temperature is (Specific heat of copper = 420 *Joule-kg*°*C*)

(a) 5°*C* (b) 6°*C*
(c) 7°*C* (d) 8°*C*

20. Two liquids *A* and *B* are at 32°*C* and 24°*C*. When mixed in equal masses the temperature of the mixture is found to be 28°*C*. Their specific heats are in the ratio of [DPMT 1996]

(a) 3 : 2 (b) 2 : 3
(c) 1 : 1 (d) 4 : 3

21. A beaker contains 200 *gm* of water. The heat capacity of the beaker is equal to that of 20 *gm* of water. The initial temperature of water in the beaker is 20°*C*. If 440 *gm* of hot water at 92°*C* is poured in it, the final temperature (neglecting radiation loss) will be nearest to[NSEP 1994]

(a) 58°*C* (b) 68°*C*
(c) 73°*C* (d) 78°*C*

22. Amount of heat required to raise the temperature of a body through 1*K* is called its [KCET 1996; MH CET 2001; AIEEE 2002]

(a) Water equivalent (b) Thermal capacity
(c) Entropy (d) Specific heat

23. A metallic ball and highly stretched spring are made of the same material and have the same mass. They are heated so that they melt, the latent heat required [AIIMS 2002]

(a) Are the same for both
(b) Is greater for the ball
(c) Is greater for the spring
(d) For the two may or may not be the same depending upon the metal

24. A liquid of mass *m* and specific heat *c* is heated to a temperature 2*T*. Another liquid of mass *m*/2 and specific heat 2*c* is heated to a temperature *T*. If these two liquids are mixed, the resulting temperature of the mixture is [EAMCET 1992]

(a) (2/3)*T* (b) (8/5)*T*
(c) (3/5)*T* (d) (3/2)*T*

25. Calorie is defined as the amount of heat required to raise temperature of 1*g* of water by 1°*C* and it is defined under which of the following conditions [IIT-JEE (Screening) 2005]

(a) From 14.5°*C* to 15.5°*C* at 760 *mm* of *Hg*
(b) From 98.5°*C* to 99.5°*C* at 760 *mm* of *Hg*
(c) From 13.5°*C* to 14.5°*C* at 76 *mm* of *Hg*
(d) From 3.5°*C* to 4.5°*C* at 76 *mm* of *Hg*

26. 100 *gm* of ice at 0°*C* is mixed with 100 *g* of water at 100°*C*. What will be the final temperature of the mixture [SCRA 1996; AMU 1999]

(a) 10°*C* (b) 20°*C*
(c) 30°*C* (d) 40°*C*

27. At atmospheric pressure, the water boils at 100°*C*. If pressure is reduced, it will boil at [MP PMT 1984]

(a) Higher temperature (b) Lower temperature
(c) At the same temperature (d) At critical temperature

28. A closed bottle containing water at 30°*C* is carried to the moon in a space-ship. If it is placed on the surface of the moon, what will happen to the water as soon as the lid is opened [EAMCET (Med.) 2000]

(a) Water will boil
(b) Water will freeze
(c) Nothing will happen on it
(d) It will decompose into H_2 and O_2

29. The thermal capacity of 40 *gm* of aluminium (specific heat = 0.2 *cal*/*gm*/°*C*) is [CBSE PMT 1990]

(a) 40 *cal*/°*C* (b) 160 *cal*/°*C*
(c) 200 *cal*/°*C* (d) 8 *cal*/°*C*

30. If temperature scale is changed from °*C* to °*F*, the numerical value of specific heat will [CPMT 1984]

(a) Increases (b) Decreased
(c) Remains unchanged (d) None of the above

31. By exerting a certain amount of pressure on an ice block, you

(a) Lower its melting point
(b) Make it melt at 0°*C* only
(c) Make it melt at a faster rate
(d) Raise its melting point

32. When we rub our palms they gets heated but to a maximum temperature because

(a) Heat is absorbed by our palm
(b) Heat is lost in the environment
(c) Produced of heat is stopped
(d) None of the above

33. A bullet moving with a uniform velocity *v*, stops suddenly after hitting the target and the whole mass melts be *m*, specific heat *S*, initial temperature 25°*C*, melting point 475°*C* and the latent heat *L*. Then *v* is given by [NCERT 1972]

(a) $mL = mS\,(475-25)+\frac{1}{2}\cdot\frac{mv^2}{J}$

(b) $mS(475-25)+mL=\frac{mv^2}{2J}$

(c) $mS\,(475-25)+mL=\frac{mv^2}{J}$

(d) $mS\,(475-25)-mL=\dfrac{mv^2}{2J}$

34. A water fall is 84 *metres* high. If half of the potential energy of the falling water gets converted to heat, the rise in temperature of water will be [JIPMER 2002]

(a) $0.098°C$ (b) $0.98°C$

(c) $9.8°C$ (d) $0.0098°C$

35. A body of mass 5 *kg* falls from a height of 30 *metre*. If its all mechanical energy is changed into heat, then heat produced will be

(a) 350 *cal* (b) 150 *cal*

(c) 60 *cal* (d) 6 *cal*

36. In supplying 400 calories of heat to a system, the work done will be

(a) 400 *joules* (b) 1672 *joules*

(c) 1672 *watts* (d) 1672 *ergs*

37. 0.93 *watt-hour* of energy is supplied to a block of ice weighing 10 *gm*. It is found that [NCERT 1973; DPMT 1999]

(a) Half of the block melts

(b) The entire block melts and the water attains a temperature of $4°C$

(c) The entire block just melts

(d) The block remains unchanged

38. The weight of a person is 60 *kg*. If he gets 10^5 calories heat through food and the efficiency of his body is 28%, then upto how much height he can climb (approximately) [AFMC 1997]

(a) 100 *m* (b) 200 *m*

(c) 400 *m* (d) 1000 *m*

39. The temperature of *Bhakhra dam* water at the ground level with respect to the temperature at high level should be

(a) Greater (b) Less

(c) Equal (d) $0°C$

40. The height of a waterfall is 84 *metre*. Assuming that the entire kinetic energy of falling water is converted into heat, the rise in temperature of the water will be

$(g=9.8\,m/s^2,\ J=4.2\ joule/cal)$ [MP PET 1994]

(a) $0.196°C$ (b) $1.960°C$

(c) $0.96°C$ (d) $0.0196°C$

41. Hailstone at $0°C$ falls from a height of 1 *km* on an insulating surface converting whole of its kinetic energy into heat. What part of it will melt $(g=10\,m/s^2)$ [MP PMT 1994]

(a) $\dfrac{1}{33}$ (b) $\dfrac{1}{8}$

(c) $\dfrac{1}{33}\times 10^{-4}$ (d) All of it will melt

42. The *SI* unit of mechanical equivalent of heat is [MP PMT/PET 1998]

(a) *Joule* × *Calorie* (b) *Joule*/*Calorie*

(c) *Calorie* × *Erg* (d) *Erg*/*Calorie*

43. Of two masses of 5 *kg* each falling from height of 10 *m*, by which 2*kg* water is stirred. The rise in temperature of water will be

(a) $2.6°C$ (b) $1.2°C$

(c) $0.32°C$ (d) $0.12°C$

44. A lead ball moving with a velocity *V* strikes a wall and stops. If 50% of its energy is converted into heat, then what will be the increase in temperature (Specific heat of lead is *S*) [CPMT 1975] [RPMT 1996]

(a) $\dfrac{2V^2}{JS}$ (b) $\dfrac{V^2}{4JS}$

[MP PMT 1989]

(c) $\dfrac{V^2}{J}$ (d) $\dfrac{V^2S}{2J}$

45. The mechanical equivalent of heat *J* is [MP PET 2000]

(a) A constant (b) A physical quantity

(c) A conversion factor (d) None of the above

46. Water falls from a height of 210*m*. Assuming whole of energy due to fall is converted into heat the rise in temperature of water would be (*J* = 4.3 *Joule*/*cal*) [Pb. PMT 2002]

(a) $42°C$ (b) $49°C$

(c) $0.49°C$ (d) $4.9°C$

47. A block of mass 100 *gm* slides on a rough horizontal surface. If the speed of the block decreases from 10 *m*/*s* to 5 *m*/*s*, the thermal energy developed in the process is [UPSEAT 2002]

(a) 3.75 *J* (b) 37.5 *J*

(c) 0.375 *J* (d) 0.75 *J*

48. 4200 *J* of work is required for [MP PMT 1986]

(a) Increasing the temperature of 10 *gm* of water through $10°C$

(b) Increasing the temperature of 100 *gm* of water through $10°C$

(c) Increasing the temperature of 1 *kg* of water through $10°C$

(d) Increasing the temperature of 10 *kg* of water through $10°C$

49. At $100°C$, the substance that causes the most severe burn, is [KCET 1999; UPSEAT 1999]

(a) Oil (b) Steam

(c) Water (d) Hot air

50. In a water-fall the water falls from a height of 100 *m*. If the entire K.E. of water is converted into heat, the rise in temperature of water will be [MP PMT 2001]

(a) $0.23°C$ (b) $0.46°C$

(c) $2.3°C$ (d) $0.023°C$

51. A lead bullet of 10 *g* travelling at 300 *m*/*s* strikes against a block of wood and comes to rest. Assuming 50% of heat is absorbed by the bullet, the increase in its temperature is

(Specific heat of lead = 150*J*/*kg*, *K*) [EAMCET 2001]

(a) $100°C$ (b) $125°C$

(c) $150°C$ (d) $200°C$

52. The temperature at which the vapour pressure of a liquid becomes equals to the external (atmospheric) pressure is its [Kerala (Engg.) 2001]

(a) Melting point (b) Sublimation point

(c) Critical temperature (d) Boiling point

53. When the pressure on water is increased the boiling temperature of water as compared to 100°*C* will be [RPET 1999]

(a) Lower

(b) The same

(c) Higher

(d) On the critical temperature

54. Calorimeters are made of which of the following [AFMC 2000]

(a) Glass (b) Metal

(c) Wood (d) Either (a) or (c)

55. Triple point of water is [CPMT 2002]

(a) 273.16°*F* (b) 273.16 *K*

(c) 273.16°*C* (d) 273.16 *R*

56. A liquid boils when its vapour pressure equals [MP PET 2002]

(a) The atmospheric pressure

(b) Pressure of 76.0 *cm* column of mercury

(c) The critical pressure

(d) The dew point of the surroundings

57. The amount of work, which can be obtained by supplying 200 *cal* of heat, is [Pb. PET 2001, 03; BHU 2004]

(a) 840 *dyne* (b) 840 *W*

(c) 840 *erg* (d) 840 *J*

58. How many grams of a liquid of specific heat 0.2 at a temperature 40°*C* must be mixed with 100 *gm* of a liquid of specific heat of 0.5 at a temperature 20°*C*, so that the final temperature of the mixture becomes 32°*C* [Pb. PET 1999]

(a) 175 *gm* (b) 300 *g*

(c) 295 *gm* (d) 375 *g*

59. 1 *g* of a steam at 100°*C* melt how much ice at 0°*C*? (Latent heat of ice = 80 *cal/gm* and latent heat of steam = 540 *cal/gm*)

(a) 1 *gm* (b) 2 *gm*

(c) 4 *gm* (d) 8 *gm*

60. 5 *g* of ice at 0°*C* is dropped in a beaker containing 20 *g* of water at 40°*C*. The final temperature will be [Pb. PET 2003]

(a) 32°*C* (b) 16°*C*

(c) 8°*C* (d) 24°*C*

61. One kilogram of ice at 0°*C* is mixed with one kilogram of water at 80°*C*. The final temperature of the mixture is

(Take : specific heat of water $= 4200\ J\ kg^{-1}\ K^{-1}$, latent heat of ice $= 336\ kJ\ kg^{-1}$) [KCET 2002]

(a) 40°*C* (b) 60°*C*

(c) 0°*C* (d) 50°*C*

62. During constant temperature, we feel colder on a day when the relative humidity will be [Pb. PMT 1996]

(a) 25% (b) 12.5%

(c) 50% (d) 75%

63. Which of the following is the unit of specific heat [MH CET 2004]

(a) $J\ kg\ °C^{-1}$ (b) $J / kg\ °C$

(c) $kg\ °C / J$ (d) $J / kg\ °C^{-2}$

64. 50 *gm* of ice at 0°*C* is mixed with 50 *gm* of water at 80°*C*, final temperature of mixture will be [DCE 2002]

(a) 0°*C* (b) 40°*C*

(c) 40°*C* (d) 4°*C*

65. The freezing point of the liquid decreases when pressure is increased, if the liquid [DCE 1995]

(a) Expands while freezing

(b) Contracts while freezing

(c) Does not change in volume while freezing

(d) None of these

66. The relative humidity on a day, when partial pressure of water vapour is $0.012 \times 10^5\ Pa$ at 12°*C* is (take vapour pressure of water at this temperature as $0.016 \times 10^5\ Pa$) [AIIMS 1998]

(a) 70% (b) 40%

(c) 75% (d) 25%

67. A hammer of mass 1*kg* having speed of 50 *m/s*, hit a iron nail of mass 200 *gm*. If specific heat of iron is 0.105 *cal/gm°C* and half the energy is converted into heat, the raise in temperature of nail is

(a) 7.1°*C* (b) 9.2°*C*

(c) 10.5°*C* (d) 12.1°*C*

68. Latent heat of 1*gm* of steam is 536 *cal/gm*, then its value in *joule/kg* is [RPMT 1999]

(a) 2.25×10^6 (b) 2.25×10^3

(c) 2.25 (d) None

69. Which of the following has maximum specific heat [RPMT 1999]

(a) Water (b) Alcohol

(c) Glycerine [Pb. PET 2000] (d) Oil

70. 50 *gm* ice at 0°*C* in insulator vessel, 50*g* water of 100°*C* is mixed in it, then final temperature of the mixture is (neglect the heat loss)

(a) 10°*C* (b) $0° << T_m < 20°C$

(c) 20°*C* (d) Above 20°*C*

71. A stationary object at 4°*C* and weighing 3.5 *kg* falls from a height of 2000 *m* on a snow mountain at 0°*C*. If the temperature of the object just before hitting the snow is 0°*C* and the object comes to rest immediately $(g = 10\ m/s^2)$ and (latent heat of ice $= 3.5 \times 10^5$ *joule/sec*), then the object will melt

(a) 2 *kg* of ice (b) 200 *gm* of ice

(c) 20 *gm* ice (d) 2 *gm* of ice

72. 300 *gm* of water at 25°*C* is added to 100 *gm* of ice at 0°*C*. The final temperature of the mixture is **[MP PET 2004]**

(a) $-\frac{5}{3}°C$ (b) $-\frac{5}{2}°C$

(c) – 5°*C* (d) 0°*C*

73. Calculate the amount of heat (in calories) required to convert 5 *gm* of ice at 0°*C* to steam at 100°*C* **[DPMT 2005]**

(a) 3100 (b) 3200

(c) 3600 (d) 4200

74. 2*gm* of steam condenses when passed through 40*gm* of water initially at 25°*C*. The condensation of steam raises the temperature of water to 54.3°*C*. What is the latent heat of steam

(a) 540 *cal/g* (b) 536 *cal/g*

(c) 270 *cal/g* (d) 480 *cal/g*

75. 10 *gm* of ice at 0°*C* is mixed with 100 *gm* of water at 50°*C*. What is the resultant temperature of mixture **[AFMC 2005]**

(a) 31.2°*C* (b) 32.8°*C*

(c) 36.7°*C* (d) 38.2°*C*

76. Three liquids with masses m_1, m_2, m_3 are thoroughly mixed. If their specific heats are c_1, c_2, c_3 and their temperatures T_1, T_2, T_3 respectively, then the temperature of the mixture is

(a) $\frac{c_1T_1 + c_2T_2 + c_3T_3}{m_1c_1 + m_2c_2 + m_3c_3}$

(b) $\frac{m_1c_1T_1 + m_2c_2T_2 + m_3c_3T_3}{m_1c_1 + m_2c_2 + m_3c_3}$

(c) $\frac{m_1c_1T_1 + m_2c_2T_2 + m_3c_3T_3}{m_1T_1 + m_2T_2 + m_3T_3}$

(d) $\frac{m_1T_1 + m_2T_2 + m_3T_3}{c_1T_1 + c_2T_2 + c_3T_3}$

77. The point on the pressure temperature phase diagram where all the phases co-exist is called **[MH CET 2005]**

(a) Sublimation (b) Fusion point

(c) Triple point (d) Vaporisation point

78. Boiling water is changing into steam. At this stage the specific heat of water is **[UPSEAT 1998]**

(a) < 1 (b) ∞

(c) 1 (d) 0

79. A vessel contains 110 *g* of water. The heat capacity of the vessel is equal to 10 *g* of water. The initial temperature of water in vessel is 10°*C*. If 220 *g* of hot water at 70°*C* is poured in the vessel, the final temperature neglecting radiation loss, will be

(a) 70°*C* (b) 80°*C*

(c) 60°*C* (d) 50°*C*

80. The thermal capacity of a body is 80 *cal*, then its water equivalent is **[UPSEAT 2001]**

(a) 80 *cal / gm* (b) 8 *gm*

(c) 80 *gm* (d) 80 *kg*

81. A liquid of mass *M* and specific heat *S* is at a temperature 2*t*. If another liquid of thermal capacity 1.5 times, at a temperature of $\frac{t}{3}$ is added to it, the resultant temperature will be

(a) $\frac{4}{3}t$ (b) t

(c) $\frac{t}{2}$ (d) $\frac{2}{3}t$

82. Dry ice is **[CPMT 2000]**

(a) Ice cube (b) Sodium chloride

(c) Liquid nitrogen **[J & K CET 2005]** (d) Solid carbon dioxide

Critical Thinking

Objective Questions

1. A glass flask is filled up to a mark with 50 *cc* of mercury at 18°*C*. If the flask and contents are heated to 38°*C*, how much mercury will be above the mark ? (α for glass is 9 × 10^{-6}/°*C* and coefficient of real expansion of mercury is 180 × 10^{-6}/°*C*)

(a) 0.85 *cc* (b) 0.46 *cc*

(c) 0.153 *cc* (d) 0.05 *cc*

2. The coefficient of apparent expansion of mercury in a glass vessel is 153 × 10^{-6}/°*C* and in a steel vessel is 144 × 10^{-6}/°*C*. If α for steel is 12 × 10^{-6}/°*C*, then that of glass is **[EAMCET 1997]**

(a) 9 × 10^{-6}/°*C* (b) 6 × 10^{-6}/°*C*

(c) 36 × 10^{-6}/°*C* (d) 27 × 10^{-6}/°*C*

3. Solids expand on heating because **[CPMT 1990]**

(a) Kinetic energy of the atoms increases

(b) Potential energy of the atoms increases

(c) Total energy of the atoms increases

(d) The potential energy curve is asymmetric about the equilibrium distance between neighbouring atoms

4. An iron tyre is to be fitted on to a wooden wheel 1*m* in diameter. The diameter of tyre is 6 *mm* smaller than that of wheel. The tyre should be heated so that its temperature increases by a minimum of (the coefficient of cubical expansion of iron is 3.6 × 10^{-5}/°*C*) **[CPMT 1989]**

(a) 167°*C* (b) 334°*C*

(c) 500°*C* (d) 1000°*C*

5. A glass flask of volume one *litre* at 0°*C* is filled, level full of mercury at this temperature. The flask and mercury are now heated to 100°*C*. How much mercury will spill out, if coefficient of volume expansion of mercury is $1.82 \times 10^{-4}/°C$ and linear expansion of glass **[UPSEAT 2000]** is $0.1 \times 10^{-4}/°C$ respectively **[MNR 1994]**

(a) 21.2 *cc* (b) 15.2 *cc*

(c) 1.52 *cc* (d) 2.12 *cc*

6. A steel scale measures the length of a copper wire as $80.0\,cm$, when both are at $20°C$ (the calibration temperature for scale). What would be the scale read for the length of the wire when both are at $40°C$? (Given $\alpha_{steel} = 11 \times 10^{-6}$ per°C and $\alpha_{copper} = 17 \times 10^{-6}\ per\,°C$)

[CPMT 2004]

(a) $80.0096\,cm$ (b) $80.0272\,cm$

(c) $1\,cm$ (d) $25.2\,cm$

7. A bimetallic strip is formed out of two identical strips, one of copper and other of brass. The coefficients of linear expansion of the two metals are α_C and α_B. On heating, the temperature of the strip goes up by ΔT and the strip bends to form an arc of radius of curvature *R*. Then *R* is

[IIT-JEE (Screening) 1999]

(a) Proportional to ΔT

(b) Inversely proportional to ΔT

(c) Proportional to $|\alpha_B - \alpha_C|$

(d) Inversely proportional to $|\alpha_B - \alpha_C|$

8. Two metal strips that constitute a thermostat must necessarily differ in their **[IIT-JEE 1992]**

(a) Mass

(b) Length

(c) Resistivity

(d) Coefficient of linear expansion

9. A metal ball immersed in alcohol weighs W_1 at 0°*C* and W_2 at 59°*C*. The coefficient of cubical expansion of the metal is less than that of alcohol. Assuming that the density of metal is large compared to that of alcohol, it can be shown that

[CPMT 1998]

(a) $W_1 > W_2$ (b) $W_1 = W_2$

(c) $W_1 < W_2$ (d) $W_2 = (W_1 / 2)$

10. The coefficient of volumetric expansion of mercury is $18 \times 10^{-5}/°C$. A thermometer bulb has a volume $10^{-6}\ m^3$ and cross section of stem is 0.004 *cm*. Assuming that bulb is filled with mercury at 0°*C* then the length of the mercury column at 100°*C* is **[Pb. PMT 1998, DPMT 1997, 2001]**

(a) 18.8 *mm* (b) 9.2 *mm*

(c) 7.4 *cm* (d) 4.5 *cm*

11. A piece of metal weight 46 *gm* in air, when it is immersed in the liquid of specific gravity 1.24 at 27°*C* it weighs 30 *gm*. When the temperature of liquid is raised to 42°*C* the metal piece weight 30.5 *gm*, specific gravity of the liquid at 42°*C* is 1.20, then the linear expansion of the metal will be

[BHU 1995]

(a) $3.316 \times 10^{-5}/°C$ (b) $2.316 \times 10^{-5}/°C$

(c) $4.316 \times 10^{-5}/°C$ (d) None of these

12. It is known that wax contracts on solidification. If molten wax is taken in a large vessel and it is allowed to cool slowly, then

(a) It will start solidifying from the top downward

(b) It will start solidifying from the bottom upward

(c) It will start solidifying from the middle, upward and downward at equal rates

(d) The whole mass will solidify simultaneously

13. A substance of mass *m* kg requires a power input of *P* watts to remain in the molten state at its melting point. When the power is turned off, the sample completely solidifies in time *t* sec. What is the latent heat of fusion of the substance

[IIT JEE 1992]

(a) $\frac{Pm}{t}$ (b) $\frac{Pt}{m}$

(c) $\frac{m}{Pt}$ (d) $\frac{t}{Pm}$

14. Steam at 100°*C* is passed into 1.1 *kg* of water contained in a calorimeter of water equivalent 0.02 *kg* at 15°*C* till the temperature of the calorimeter and its contents rises to 80°*C*. The mass of the steam condensed in *kg* is

[IIT 1995]

(a) 0.130 (b) 0.065

(c) 0.260 (d) 0.135

15. 2 *kg* of ice at – 20°*C* is mixed with 5 *kg* of water at 20°*C* in an insulating vessel having a negligible heat capacity. Calculate the final mass of water remaining in the container. It is given that the specific heats of water and ice are 1 *kcal/kg per* °*C* and 0.5 *kcal/kg/*°*C* while the latent heat of fusion of ice is 80 *kcal/kg* **[IIT-JEE (Screening) 2003]**

(a) 7 *kg* (b) 6 *kg*

(c) 4 *kg* (d) 2 *kg*

16. Water of volume 2 *litre* in a container is heated with a coil of $1\,kW$ at $27\,°C$. The lid of the container is open and energy dissipates at rate of $160\,J/s$. In how much time temperature will rise from $27\,°C$ to $77\,°C$ [Given specific heat of water is $4.2\,kJ/kg$]

(a) 8 *min* 20 *s* (b) 6 *min* 2 *s*

(c) 7 *min* (d) 14 *min*

17. A lead bullet at 27°*C* just melts when stopped by an obstacle. Assuming that 25% of heat is absorbed by the obstacle, then the velocity of the bullet at the time of striking (M.P. of lead = 327°*C*, specific heat of lead = 0.03 *cal/gm*°*C*, latent heat of fusion of lead = 6 *cal/gm* and *J* = 4.2 *joule/cal*) **[IIT 1981]**

(a) 410 *m/sec* (b) 1230 *m/sec*

(c) 307.5 *m/sec* (d) None of the above

18. If two balls of same metal weighing 5 *gm* and 10 *gm* strike with a target with the same velocity. The heat energy so developed is used for raising their temperature alone, then the temperature will be higher

(a) For bigger ball

(b) For smaller ball

(c) Equal for both the balls

(d) None is correct from the above three

19. The temperature of equal masses of three different liquids *A*, *B* and *C* are 12°*C*, 19°*C* and 28°*C* respectively. The temperature when *A* and *B* are mixed is 16°*C* and when *B* and *C* are mixed is 23°*C*. The temperature when *A* and *C* are mixed is **[CBSE PMT 1994]**

(a) 18.2°*C* (b) 22°*C*

(c) 20.2°*C* (d) 25.2°*C*

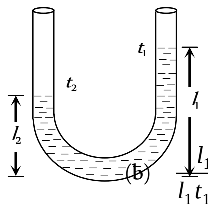
20. In an industrial process 10 *kg* of water per hour is to be heated from 20°*C* to 80°*C*. To do this steam at 150°*C* is passed from a boiler into a copper coil immersed in water. The steam condenses in the coil and is returned to the boiler as water at 90°*C*. how many *kg* of steam is required per hour.

(Specific heat of steam = 1 *calorie* per *gm*°*C*, Latent heat of vaporisation = 540 *cal/gm*)

(a) 1 *gm* (b) 1 *kg*

(c) 10 *gm* (d) 10 *kg*

21. In a vertical U-tube containing a liquid, the two arms are maintained at different temperatures t_1 and t_2. The liquid columns in the two arms have heights l_1 and l_2 respectively. The coefficient of volume expansion of the liquid is equal to



(a) $\frac{l_1 - l_2}{l_2 t_1 - l_1 t_2}$ (b) $\frac{l_1 - l_2}{l_1 t_1 - l_2 t_2}$

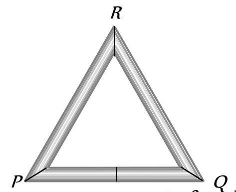
(c) $\frac{l_1 + l_2}{l_2 t_1 + l_1 t_2}$ (d) $\frac{l_1 + l_2}{l_1 t_1 + l_2 t_2}$

22. The coefficient of linear expansion of crystal in one direction is α_1 and that in every direction perpendicular to it is α_2. The coefficient of cubical expansion is

(a) $\alpha_1 + \alpha_2$ (b) $2\alpha_1 + \alpha_2$

(c) $\alpha_1 + 2\alpha_2$ (d) None of these

23. Three rods of equal length *l* are joined to form an equilateral triangle *PQR*. *O* is the mid point of *PQ*. Distance *OR* remains same for small change in temperature. Coefficient of linear expansion for *PR* and *RQ* is same *i.e.* α_2 but that for *PQ* is α_1. Then



(a) $\alpha_2 = 3\alpha_1$

(b) $\alpha_2 = 4\alpha_1$

(c) $\alpha_1 = 3\alpha_2$

(d) $\alpha_1 = 4\alpha_2$

24. A one *litre* glass flask contains some mercury. It is found that at different temperatures the volume of air inside the flak remains the same. What is the volume of mercury in this flask if coefficient of linear expansion of glass is $9 \times 10^{-6}/°C$ while of volume expansion of mercury is $1.8 \times 10^{-4}/°C$

(a) 50 *cc* (b) 100 *cc*

(c) 150 *cc* (d) 200 *cc*

25. 10 *gm* of ice at – 20°*C* is dropped into a calorimeter containing 10 *gm* of water at 10°*C*; the specific heat of water is twice that of ice. When equilibrium is reached, the calorimeter will contain

(a) 20 *gm* of water

(b) 20 *gm* of ice

(c) 10 *gm* ice and 10 *gm* water

(d) 5 *gm* ice and 15 *gm* water

26. A rod of length 20 *cm* is made of metal. It expands by 0.075 *cm* when its temperature is raised from 0°*C* to 100°*C*. Another rod of a different metal *B* having the same length expands by 0.045 *cm* for the same change in temperature. A third rod of the same length is composed of two parts, one of metal *A* and the other of metal *B*. This rod expands by 0.060 *cm* for the same change in temperature. The portion made of metal *A* has the length

(a) 20 *cm* (b) 10 *cm*

(c) 15 *cm* (d) 18 *cm*

27. Steam is passed into 22 *gm* of water at 20°*C*. The mass of water that will be present when the water acquires a temperature of 90°*C* (Latent heat of steam is 540 *cal/gm*) is

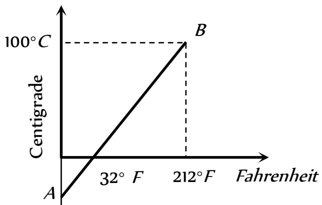
[SCRA 1994]

(a) 24.8 *gm* (b) 24 *gm*

(c) 36.6 *gm* (d) 30 *gm*

GQ Graphical Questions

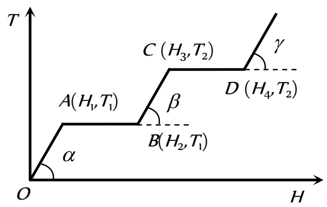
1. The graph *AB* shown in figure is a plot of temperature of a body in degree celsius and degree Fahrenheit. Then



(a) Slope of line *AB* is 9/5 (b) Slope of line *AB* is 5/9

(c) Slope of line *AB* is 1/9 (d) Slope of line *AB* is 3/9

2. The graph shows the variation of temperature (*T*) of one *kilogram* of a material with the heat (*H*) supplied to it. At *O*, the substance is in the solid state. From the graph, we can conclude that



(a) T_2 is the melting point of the solid

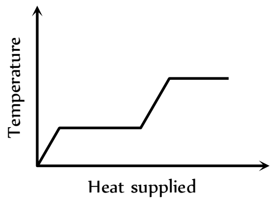
(b) *BC* represents the change of state from solid to liquid

(c) $(H_2 - H_1)$ represents the latent heat of fusion of the substance

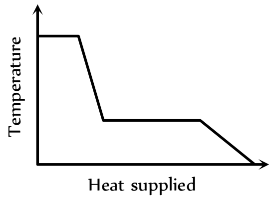
(d) $(H_3 - H_1)$ represents the latent heat of vaporization of the liquid

3. A block of ice at $-10°C$ is slowly heated and converted to steam at $100°C$. Which of the following curves represents the phenomenon qualitatively [IIT-JEE (Screening) 2000]

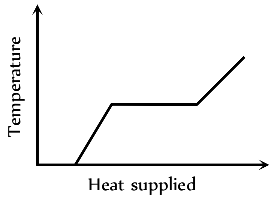
(a)



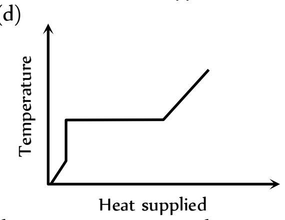
(b)



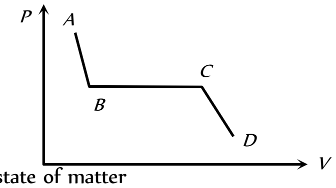
(c)



(d)



4. The portion AB of the indicator diagram representing the state of matter denotes



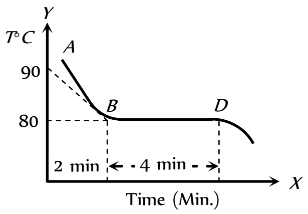
(a) The liquid state of matter

(b) Gaseous state of matter

(c) Change from liquid to gaseous state

(d) Change from gaseous state to liquid state

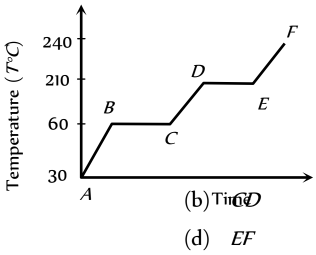
5. The figure given below shows the cooling curve of pure wax material after heating. It cools from A to B and solidifies along BD. If L and C are respective values of latent heat and the specific heat of the liquid wax, the ratio L/C is



(a) 40 (b) 80

(c) 100 (d) 20

6. A solid substance is at $30°C$. To this substance heat energy is supplied at a constant rate. Then temperature versus time graph is as shown in the figure. The substance is in liquid state for the portion (of the graph) [RPET 1990, 94]

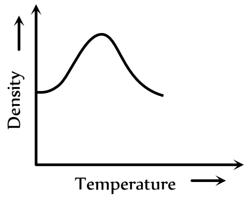


(a) BC (b) CD

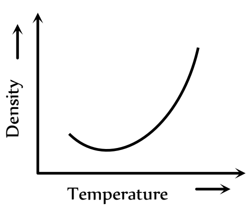
(c) ED (d) EF

7. The variation of density of water with temperature is represented by the

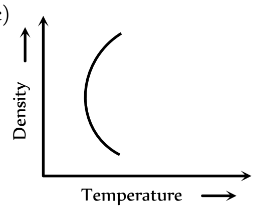
(a)



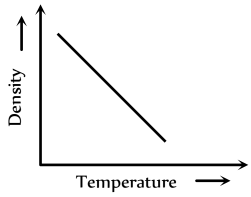
(b)



(c)



(d)



8. If a graph is plotted taking the temperature in Fahrenheit along Y-axis and the corresponding temperature in Celsius along the X-axis, it will be a straight line [AIIMS 1997]

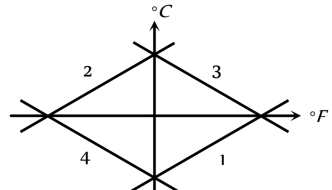
(a) Having a $+ve$ intercept on Y-axis

(b) Having a $+ve$ intercept on X-axis

(c) Passing through the origin

(d) Having $a-ve$ intercepts on both the axis

9. Which of the curves in figure represents the relation between Celsius and Fahrenheit temperatures



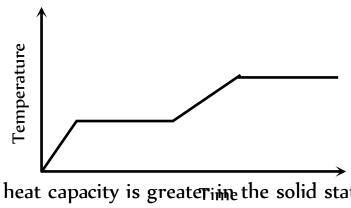
(a) 1

(b) 2

(c) 3

(d) 4

10. Heat is supplied to a certain homogenous sample of matter, at a uniform rate. Its temperature is plotted against time, as shown. Which of the following conclusions can be drawn



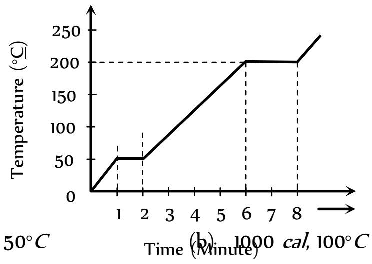
(a) Its specific heat capacity is greater in the solid state than in the liquid state

(b) Its specific heat capacity is greater in the liquid state than in the solid state

(c) Its latent heat of vaporization is greater than its latent heat of fusion

(d) Its latent heat of vaporization is smaller than its latent heat of fusion

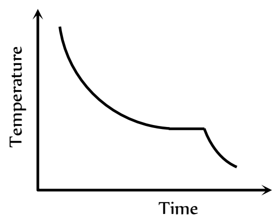
11. A student takes 50gm wax (specific heat = 0.6 $kcal/kg°C$) and heats it till it boils. The graph between temperature and time is as follows. Heat supplied to the wax per minute and boiling point are respectively [BHU 1994]



(a) 500 cal, $50°C$ (b) 1000 cal, $100°C$

(c) 1500 cal, $200°C$ (d) $200°C$

12. The graph signifies [JIPMER 1999]



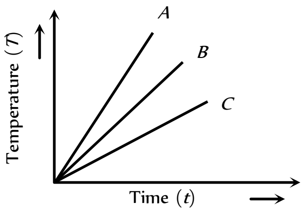
(a) Adiabatic expansion of a gas

(b) Isothermal expansion of a gas

(c) Change of state from liquid to solid

(d) Cooling of a heated solid

13. Which of the substances A, B or C has the highest specific heat ? The temperature *vs* time graph is shown



(a) A (b) B

(c) C (d) All have equal specific heat

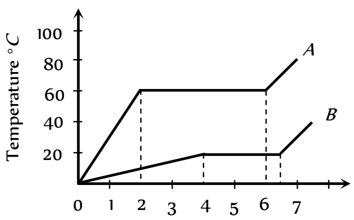
14. Two substances A and B of equal mass m are heated at uniform rate of 6 cal/s under similar conditions. A graph between temperature and time is shown in figure. Ratio of heat absorbed H_A / H_B by them for complete fusion is

(a) $\frac{9}{4}$

(b) $\frac{4}{9}$

(c) $\frac{8}{5}$

(d) $\frac{5}{8}$



AR Assertion & Reason

For AIIMS Aspirants

Read the assertion and reason carefully to mark the correct option out of the options given below :

(a) If both assertion and reason are true and the reason is the correct explanation of the assertion.
(b) If both assertion and reason are true but reason is not the correct explanation of the assertion.
(c) If assertion is true but reason is false.
(d) If the assertion and reason both are false.
(e) If assertion is false but reason is true.

1. Assertion : The melting point of ice decreases with increase of pressure.

   Reason : Ice contracts on melting. [AIIMS 2004]

2. Assertion : Fahrenheit is the smallest unit measuring temperature.

   Reason : Fahrenheit was the first temperature scale used for measuring temperature. [AIIMS 1999]

3. Assertion : Melting of solid causes no change in internal energy.

   Reason : Latent heat is the heat required to melt a unit mass of solid. [AIIMS 1998]

4. Assertion : Specific heat capacity is the cause of formation of land and sea breeze.

   Reason : The specific heat of water is more than land. [AIIMS 1995]

5. Assertion : A brass disc is just fitted in a hole in a steel plate. The system must be cooled to loosen the disc from the hole.

   Reason : The coefficient of linear expansion for brass is greater than the coefficient of linear expansion for steel.

6. Assertion : The coefficient of volume expansion has dimension K.

   Reason : The coefficient of volume expansion is defined as the change in volume per unit volume per unit change in temperature.

7. Assertion : The temperature at which Centigrade and Fahrenheit thermometers read the same is – 40°.

   Reason : There is no relation between Fahrenheit and Centigrade temperature.

8. Assertion : When a solid iron ball is heated, percentage increase is its volume is largest.

   Reason : Coefficient of superficial expansion is twice that of linear expansion where as coefficient of volume expansion is three time of linear expansion.

9. Assertion : A beaker is completely filled with water at 4°C. It will overflow, both when heated or cooled.

   Reason : There is expansion of water below and above 4°C.

10. Assertion : Latent heat of fusion of ice is 336000 J/kg.

    Reason : Latent heat refers to change of state without any change in temperature

11. Assertion : Two bodies at different temperatures, if brought in thermal contact do not necessary settle to the mean temperature.

    Reason : The two bodies may have different thermal capacities.

12. Assertion : Specific heat of a body is always greater than its thermal capacity.

    Reason : Thermal capacity is the required for raising temperature of unit mass of the body through unit degree.

13. Assertion : Water kept in an open vessel will quickly evaporate on the surface of the moon.

    Reason : The temperature at the surface of the moon is much higher than boiling point of the water.

14. Assertion : The molecules at 0°C ice and 0°C water will have same potential energy.

    Reason : Potential energy depends only on temperature of the system.

Answers

Thermometry

1	d	2	b	3	a	4	c	5	b
6	d	7	c	8	a	9	b	10	b
11	c	12	c	13	c	14	c	15	c
16	d	17	b	18	c	19	c	20	c
21	c	22	a	23	d	24	a	25	c
26	a	27	b	28	a	29	c	30	c
31	a	32	d	33	a	34	d	35	a

Thermal Expansion

1	c	2	a	3	b	4	d	5	c
6	c	7	c	8	c	9	b	10	c
11	a	12	b	13	b	14	d	15	a
16	d	17	d	18	d	19	d	20	a
21	b	22	c	23	a	24	a	25	a
26	b	27	c	28	d	29	a	30	d
31	b	32	a	33	a	34	c		

Calorimetry

1	b	2	b	3	c	4	c	5	d
6	a	7	c	8	a	9	d	10	a
11	b	12	c	13	a	14	c	15	b
16	a	17	b	18	a	19	a	20	c
21	b	22	b	23	a	24	d	25	a
26	a	27	b	28	a	29	d	30	b
31	a	32	b	33	b	34	a	35	a
36	b	37	c	38	b	39	a	40	a
41	a	42	b	43	d	44	b	45	c
46	c	47	a	48	b	49	b	50	a
51	c	52	d	53	c	54	b	55	b
56	a	57	d	58	d	59	d	60	b
61	c	62	a	63	a	64	a	65	a
66	c	67	a	68	a	69	a	70	a
71	b	72	d	73	c	74	a	75	d
76	b	77	c	78	b	79	d	80	c
81	b	82	d						

Critical Thinking Questions

1	c	2	a	3	d	4	c	5	b
6	a	7	bd	8	d	9	c	10	d
11	b	12	b	13	b	14	a	15	b
16	a	17	a	18	c	19	c	20	b
21	a	22	c	23	d	24	c	25	c
26	b	27	a						

Graphical Questions

1	b	2	c	3	a	4	a	5	d
6	b	7	a	8	a	9	a	10	bc
11	c	12	c	13	c	14	c		

Assertion and Reason

1	a	2	c	3	e	4	a	5	a
6	a	7	c	8	a	9	a	10	b
11	a	12	d	13	a	14	d		

Answers and Solutions

Thermometry

1. (d) $T = 273.15 + t°C \Rightarrow 0 = 273.15 + t°C$
$\Rightarrow t = -273.15°C$

2. (b) $\frac{C}{5} = \frac{F-32}{9} \Rightarrow \frac{-183}{5} = \frac{F-32}{9} \Rightarrow F = -297°F$

3. (a) $\frac{F-32}{9} = \frac{K-273}{5} \Rightarrow \frac{F-32}{9} = \frac{95-273}{5} \Rightarrow F = -288°F$

4. (c) Temperature change in Celsius scale = Temperature change in Kelvin scale = 27 K

5. (b) Change in resistance $3.70 - 2.71 = 0.99\,\Omega$ corresponds to interval of temperature $90°C$.

So change in resistance $3.26 - 2.71 = 0.55\,\Omega$ Corresponds to change in temperature

$$= \frac{90}{0.99} \times 0.55 = 50°C$$

6. (d) $-200°C$ to $600°C$ can be measured by platinum resistance thermometer.

7. (c) Pyrometer can measure temperature from $800°C$ to $6000°C$. Hence temperature of sun is measured with pyrometer.

8. (a) $\overline{v^2} \propto T$

9. (b) Thermoelectric thermometer is based on Seeback Effect.

10. (b) Maximum density of water is at $4°C$

Also $\frac{C}{5} = \frac{F-32}{9} \Rightarrow \frac{4}{5} = \frac{F-32}{9} \Rightarrow F = 39.2°F$

11. (c) Production and measurement of temperature close to $0\,K$ is done in cryogenics

12. (c)

13. (c) At absolute zero (*i.e.* 0 K) v_{rms} becomes zero.

14. (c)

15. (c) We know that $P = P_0(1+\gamma t)$ and $V = V_0(1+\gamma t)$

and $\gamma = (1/273)/°C$ for $t = -273°C$, we have $P = 0$ and $V = 0$

Hence, at absolute zero, the volume and pressure of the gas become zero.

16. (d) Zero kelvin $= -273°C$ (absolute temperature). As no matter can attain this temperature, hence temperature can never be negative on Kelvin scale.

17. (b) $\frac{C}{5} = \frac{F-32}{9} \Rightarrow \frac{25}{5} = \frac{F-32}{9} \Rightarrow F = 77°F$.

18. (c) Thermoelectric thermometer is used for finding rapidly varying temperature.

19. (c) Due to evaporation cooling is caused which lowers the temperature of bulb wrapped in wet hanky.

20. (c) $\frac{F-32}{9} = \frac{K-273}{5} \Rightarrow \frac{x-32}{9} = \frac{x-273}{5} \Rightarrow x = 574.25$

21. (c) $\frac{C}{5} = \frac{F-32}{9} \Rightarrow \frac{C}{5} = \frac{(140-32)}{9} \Rightarrow C = 60°$

22. (a) $\frac{C}{5} = \frac{F-32}{9} \Rightarrow \frac{t}{5} = \frac{t-32}{9} \Rightarrow t = -40°$

23. (d) Standardisation of thermometers is done with gas thermometer.

24. (a) For gases γ is more.

25. (c) The boiling point of mercury is 400°*C*. Therefore, the mercury thermometer can be used to measure the temperature upto 360°*C*.

26. (a) $t = \frac{(P_t - P_0)}{(P_{100} - P_0)} \times 100°C = \frac{(60-50)}{(90-50)} \times 100 = 25°C$

27. (b) By filling nitrogen gas at high pressure, the boiling point of mercury is increased which extend the range upto 500°*C*.

28. (a) Pyrometer is used to measure very high temperature.

29. (c) $\frac{F-32}{9} = \frac{K-273}{5} \Rightarrow \frac{F-32}{9} = \frac{0-273}{5}$

$\Rightarrow F = -459.4°F \approx -460°F$

30. (c) Initial volume $V_1 = 47.5$ units

Temperature of ice cold water $T_1 = 0°C = 273\,K$

Final volume of $V_2 = 67$ units

Applying Charle's law, we have $\frac{V_1}{T_1} = \frac{V_2}{T_2}$

(where temperature T_2 is the boiling point)

or $T_2 = \frac{V_2}{V_1} \times T_1 = \frac{67 \times 273}{47.5} = 385\,K = 112°C$

31. (a) Temperature on any scale can be converted into other scale by $\frac{x - LFP}{UFP - LFP}$ = Constant for all scales $\frac{x-20}{150-20} = \frac{60}{100} \Rightarrow$ $x = 98°C$

32. (d) $\frac{C}{5} = \frac{F-32}{9} \Rightarrow \frac{C}{5} = \frac{140-32}{9} \Rightarrow C = 60°C$

33. (a) Rapidly changing temperature is measured by thermocouple thermometers.

34. (d) Difference of 100° C = difference of 180° F

$\therefore$ Difference of 30° $= \frac{180}{100} \times 30 = 54°$

35. (a)

Thermal Expansion

1. (c) When a copper ball is heated, it's size increases. As Volume $\propto (\text{radius})^3$ and Area $\propto ((\text{radius})^2$, so percentage increase will be largest in it's volume. Density will decrease with rise in temperature.

2. (a) $\frac{h_1}{h_2} = \frac{\rho_1}{\rho_2} = \frac{(1+\gamma\theta_1)}{(1+\gamma\theta_2)}$ $\left[\because \rho = \frac{\rho_0}{(1+\gamma\theta)}\right]$

$\Rightarrow \frac{50}{60} = \frac{1+\gamma \times 50}{1+\gamma \times 100} \Rightarrow \gamma = 0.005/°C$

3. (b) $\gamma_r = \gamma_a + \gamma_v$; where γ_r = coefficient of real expansion, γ_a = coefficient of apparent expansion and γ_v = coefficient of expansion of vessel.

For copper $\gamma_r = C + 3\alpha_{Cu} = C + 3A$

For silver $\gamma_r = S + 3\alpha_{Ag}$

$\Rightarrow C + 3A = S + 3\alpha_{Ag} \Rightarrow \alpha_{Ag} = \frac{C - S + 3A}{3}$

4. (d) Fractional change in period

$\frac{\Delta T}{T} = \frac{1}{2}\alpha\Delta\theta = \frac{1}{2} \times 2 \times 10^{-6} \times 10 = 10^{-5}$

% change $= \frac{\Delta T}{T} \times 100 = 10^{-5} \times 100 = 10^{-3}$ %

5. (c) $L = L_0(1 + \alpha\Delta\theta) \Rightarrow \frac{L_1}{L_2} = \frac{1+\alpha(\Delta\theta)_1}{1+\alpha(\Delta\theta)_2}$

$\Rightarrow \frac{10}{L_2} = \frac{1 + 11 \times 10^{-6} \times 20}{1 + 11 \times 10^{-6} \times 19} \Rightarrow L_2 = 9.99989$

$\Rightarrow$ Length is shorten by

$10 - 9.99989 = 0.00011 = 11 \times 10^{-5}\ cm$

6. (c) Stress $= Y\alpha\Delta\theta$; hence it is independent of length.

7. (c) Solids, liquids and gases all expand on being heated as result density (= mass/volume) decreases.

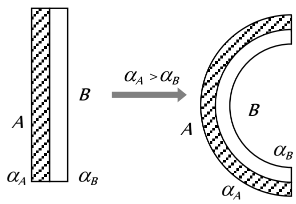
8. (c) As coefficient of cubical expansion of liquid equals coefficient of cubical expansion of vessel, the level of liquid will not change on heating.

9. (b) Loss in time per second $\frac{\Delta T}{T} = \frac{1}{2}\alpha\Delta\theta = \frac{1}{2}\alpha(t-0)$

$\Rightarrow$ loss in time per day

$\Delta t = \left(\frac{1}{2}\alpha t\right)t = \frac{1}{2}\alpha t \times (24 \times 60 \times 60) = \frac{1}{2}\alpha t \times 86400$

10. (c) A bimetallic strip on being heated bends in the form of an arc with more expandable metal (A) outside (as shown) correct.



11. (a) When the ball is heated, expansion of ball and cavity both occurs, hence volume of cavity increases.

12. (b) In summer alcohol expands, density decreases, so 1 litre of alcohol will weigh less in summer than in winter.

13. (b) Similar to previous question, benzene contracts in winter. So 5 litre of benzene will weigh more in winter than in summer.

14. (d) Water has maximum density at 4° C.

15. (a) Since coefficient of expansion of steel is greater than that of bronze. Hence with small increase in it's temperature the hole expand sufficiently.

16. (d) $A \propto L^2 \Rightarrow \frac{\Delta A}{A} = 2 \cdot \frac{\Delta L}{L} \Rightarrow \frac{\Delta A}{A} = 2 \times 2 = 4\%$.

17. (d) $\frac{V_1}{V_2} = \frac{1+\gamma t_1}{1+\gamma t_2} \Rightarrow \frac{100}{125} = \frac{1+\gamma \times 20}{1+\gamma \times 100} \Rightarrow \gamma = 0.0033/°C$

18. (d) $\alpha = \frac{\beta}{2} = \frac{2 \times 10^{-5}}{2} = 10^{-5}/°C$

19. (d) Coefficient of volume expansion

$\gamma = \frac{\Delta\rho}{\rho.\Delta T} = \frac{(\rho_1 - \rho_2)}{\rho.(\Delta\theta)} = \frac{(10-9.7)}{10 \times (100-0)} = 3 \times 10^{-4}$

Hence, coefficient of linear expansion

$\alpha = \frac{\gamma}{3} = 10^{-4}/°C$

20. (a) $\rho = \rho_0(1 - \gamma.\Delta\theta) = 13.6[1 - 0.18 \times 10^{-3}(473 - 273)]$

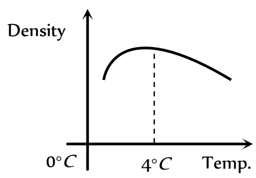
$= 13.6[1 - 0.036] = 13.11\ gm/cc$.

21. (b) As we know $\gamma_{\text{real}} = \gamma_{\text{app.}} + \gamma_{\text{vessel}}$

$\Rightarrow \gamma_{\text{app.}} = \gamma_{\text{glycerine}} - \gamma_{\text{glass}}$

$= 0.000597 - 0.000027 = 0.00057/°C$

22. (c) Water has maximum density at 4° C, so if the water is heated above 4° C or cooled below 4° C density decreases *i.e.* volume increases. In other words, it expands so it overflows in both the cases.



23. (a)

$$\gamma = \frac{\Delta V}{V.\Delta T} = \frac{0.24}{100 \times 40} = 6 \times 10^{-5} / °C$$

$$\Rightarrow \alpha = \frac{\gamma}{3} = 2 \times 10^{-5} / °C$$

24. (a) As $\alpha = \frac{\beta}{2} = \frac{\gamma}{3} \Rightarrow \alpha : \beta : \gamma = 1 : 2 : 3$

25. (a) $\gamma_{app.} = \frac{\text{Mass expelled}}{\text{Mass remained } \times \Delta T}$

$$= \frac{x/100}{x \times 80} = \frac{1}{8000} = 1.25 \times 10^{-4} / °C$$

26. (b) In anomalous expansion, water contracts on heating and expands on cooling in the range 0°*C* to 4°*C*. Therefore water pipes sometimes burst, in cold countries.

27. (c) On heating the system; *x, r, d* all increases, since the expansion of isotropic solids is similar to true photographic enlargement.

28. (d) $\alpha = \frac{\Delta L}{L_0 \times \Delta\theta} = \frac{0.01}{5 \times 100} = 2 \times 10^{-5} / °C$

29. (a) $\alpha = \frac{\Delta L}{L_0(\Delta\theta)} = \frac{0.19}{100(100-0)} = 1.9 \times 10^{-5} / °C$

Now $\gamma = 3\alpha = 3 \times 1.9 \times 10^{-5}/°C = 5.7 \times 10^{-5}/°C$

30. (d) Since, the coefficient of linear expansion of brass is greater than that of steel. On cooling, the brass contracts more, so, it get loosened.

31. (b) Increase in length $\Delta L = L_0 \alpha \Delta\theta$

$= 10 \times 10 \times 10^{-6} \times (100 - 0) = 10^{-2}\ m = 1\ cm$

32. (a) $\alpha = \frac{\Delta L}{L_0 \Delta\theta} = \frac{(1-0.9997)}{0.9997 \times 12 \times 10^{-6}} = 25°C$

33. (c) The densest layer of water will be at bottom. The density of water is maximum at 4°*C*. So the temperature of bottom of lake will be 4°*C*.

34. (c) Given $\Delta l_1 = \Delta l_2$ or $l_1 \alpha_a t = l_2 \alpha_s t$

$\therefore \frac{l_1}{l_2} = \frac{\alpha_s}{\alpha_a}$ or $\frac{l_1}{l_1 + l_2} = \frac{\alpha_s}{\alpha_a + \alpha_s}$.

Calorimetry

1. (b) In vapor to liquid phase transition, heat liberates.

2. (b) Pressure inside the mines is greater than that of normal. Pressure. Also we know that boiling point increases with increase in pressure.

3. (c) $Q = m.c.\Delta\theta \Rightarrow c = \frac{Q}{m.\Delta\theta}$; when $\Delta\theta = 0 \Rightarrow c = \infty$

4. (c) Mass and volume of the gas will remain same, so density will also remain same.

5. (d)

6. (a) The latent heat of vaporization is always greater than latent heat of fusion because in liquid to vapour phase change there is a large increase in volume. Hence more heat is required as compared to solid to liquid phase change.

7. (c) When state is not changing $\Delta Q = mc\Delta\theta$.

8. (a) Heat taken by ice to melt at 0°*C* is

$Q_1 = mL = 540 \times 80 = 43200\, cal$

Heat given by water to cool upto 0°*C* is

$Q_2 = ms\Delta\theta = 540 \times 1 \times (80 - 0) = 43200\, cal$

Hence heat given by water is just sufficient to melt the whole ice and final temperature of mixture is 0°*C*.

Short trick : For these type of frequently asked questions you can remember the following formula

$$\theta_{mix} = \frac{m_W \theta_W - \frac{m_i L_i}{c_W}}{m_i + m_W}$$ (See theory for more details)

If $m_W = m_i$ then $\theta_{mix} = \frac{\theta_W - \frac{L_i}{c_W}}{2} = \frac{80 - \frac{80}{1}}{2} = 0°C$

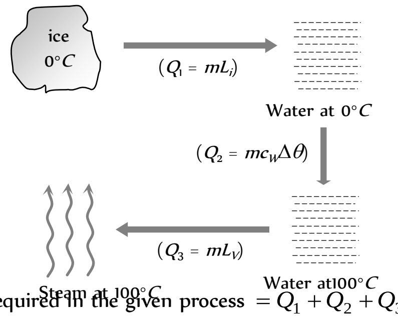
9. (d) Due to large specific heat of water, it releases large heat with very small temperature change.

10. (a) $Q = m.c.\Delta\theta = 5 \times (1000 \times 4.2) \times (100 - 20)$

$= 1680 \times 10^3\, J = 1680\, kJ$

11. (b) Melting point of ice decreases with increase in pressure (as ice expands on solidification).

12. (c) Conversion of ice (0°*C*) into steam (100°*C*) is as follows



Heat required in the given process $= Q_1 + Q_2 + Q_3$

$= 1 \times 80 + 1 \times 1 \times (100 - 0) + 1 \times 536 = 716\, cal$

13. (a) If *m gm* ice melts then

Heat lost = Heat gain

$80 \times 1 \times (30 - 0) = m \times 80 \Rightarrow m = 30\, gm$

14. (c) At boiling point saturation vapour pressure becomes equal to atmospheric pressure. Therefore, at 100°*C* for water. S.V.P. = 760 *mm* of *Hg* (atm pressure).

15. (b) Thermal capacity = Mass × Specific heat

Due to same material both spheres will have same specific heat.

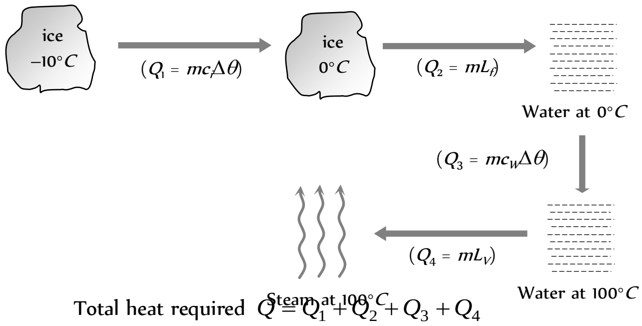
Also mass = Volume (*V*) × Density (ρ)

$\therefore$ Ratio of thermal capacity

$$= \frac{m_1}{m_2} = \frac{V_1 \rho}{V_2 \rho} = \frac{\frac{4}{3}\pi r_1^3}{\frac{4}{3}\pi r_2^3} = \left(\frac{r_1}{r_2}\right)^3 = \left(\frac{1}{2}\right)^3 = 1 : 8$$

16. (a) Ice (–10°*C*) converts into steam as follows

(c_i = Specific heat of ice, c_w = Specific heat of water)



Total heat required $Q = Q_1 + Q_2 + Q_3 + Q_4$

$\Rightarrow Q = 1 \times 0.5(10) + 1 \times 80 + 1 \times 1 \times (100 - 0) + 1 \times 540$

$= 725\, cal$

Hence work done $W = JQ = 4.2 \times 725 = 3045\,J$

17. (b) When water is cooled at 0°C to form ice then 80 *calorie/gm* (latent heat) energy is released. Because potential energy of the molecules decreases. Mass will remain constant in the process of freezing of water.

18. (a) Steam at 100°C contains extra 540 *calorie/gm* energy as compare to water at 100°C. So it's more dangerous to burn with steam then water.

19. (a) Same amount of heat is supplied to copper and water so $m_c c_c \Delta\theta_c = m_W c_W \Delta\theta_W$

$$\Rightarrow \Delta\theta_W = \frac{m_c c_c (\Delta\theta)_c}{m_W c_W} = \frac{50\times10^{-3}\times420\times10}{10\times10^{-3}\times4200} = 5°C$$

20. (c) Temperature of mixture $\theta_{mix} = \dfrac{\theta_A c_A + \theta_B c_B}{c_A + c_B}$

$$\Rightarrow 28 = \frac{32\times c_A + 24\times c_B}{c_A + c_B}$$

$$\Rightarrow 28\,c_A + 28 c_B = 32\,c_A + 24\,c_B \Rightarrow \frac{c_A}{c_B} = \frac{1}{1}$$

21. (b) Heat lost by hot water = Heat gained by cold water in beaker + Heat absorbed by beaker

$$\Rightarrow 440(92-\theta) = 200\times(\theta-20) + 20\times(\theta-20)$$

$$\Rightarrow \theta = 68°C$$

22. (b) $Q = m.c.\Delta\theta$; if $\Delta\theta = 1\,K$ then $Q = mc$ = Thermal capacity.

23. (a) Latent heat is independent of configuration. Ordered energy spent in stretching the spring will not contribute to heat which is disordered kinetic energy of molecules of substance.

24. (d) Temperature of mixture

$$\theta_{mix} = \frac{m_1c_1\theta_1 + m_2c_2\theta_2}{m_1c_1 + m_2c_2} = \frac{m\times c\times 2T + \frac{m}{2}(2c)T}{m.c + \frac{m}{2}(2c)} = \frac{3}{2}T$$

25. (a)

26. (a) $\theta_{mix} = \dfrac{\theta_W - \frac{L_i}{c_W}}{2} = \dfrac{100 - \frac{80}{1}}{2} = 10°C$

27. (b) When pressure decreases, boiling point also decreases.

28. (a) Boiling occurs when the vapour pressure of liquid becomes equal to the atmospheric pressure. At the surface of moon, atmospheric pressure is zero, hence boiling point decreases and water begins to boil at 30°C.

29. (d) Thermal capacity $= mc = 40\times0.2 = 8\,cal/°C$.

30. (b) $Q = m.c.\Delta\theta \Rightarrow c = \dfrac{Q}{m.\Delta\theta}$

In temperature measurement scale $\Delta\theta°F > \Delta\theta°C$ so $(c)_{°F} < (c)_{°C}$.

31. (a) Increasing pressure lowers melting point of ice.

32. (b) Work done changes into heat energy, when the temperature of palm becomes above the atmosphere so it starts losing heat to the surroundings.

33. (b) Firstly the temperature of bullet rises up to melting point, then it melts. Hence according to $W = JQ$.

$$\Rightarrow \frac{1}{2}mv^2 = J.[m.c.\Delta\theta + mL] = J[m\,S\,(475-25) + mL]$$

$$\Rightarrow mS(475-25) + mL = \frac{mv^2}{2J}$$

34. (a) As $W = JQ \Rightarrow \dfrac{1}{2}(mgh) = J\times mc\Delta\theta \Rightarrow \Delta\theta = \dfrac{gh}{2Jc}$

$$\Delta\theta = \frac{9.8\times84}{2\times4.2\times1000} = 0.098°C$$

$$(\because c_{\text{water}} = 1000\,\frac{cal}{kg\times°C})$$

Short trick : Remember the value of $\dfrac{g}{Jc_W} = 0.0023$, here

$$\Delta\theta = \frac{1}{2}\times(0.0023)h = \frac{1}{2}\times0.0023\times84 = 0.098°C$$

35. (a) $W = JQ \Rightarrow mgh = J\times Q$

$$\Rightarrow Q = \frac{mgh}{J} = \frac{5\times9.8\times30}{4.2} = 350\,cal$$

36. (b) $W = JQ = 4.18\times400 = 1672\,joule$

37. (c) Energy supplied $= 0.93\times3600\,joules$ = 3348 *joules*

Heat required to melt 10 *gms* of ice

$= 10\times80\times4.18 = 3344\,joules$

Hence block of ice just melts.

38. (b) Suppose person climbs upto height *h*, then by using

$W = JQ \Rightarrow mgh = JQ$

$$\Rightarrow 60\times9.8\times h = 4.2\times\left(10^5\times\frac{28}{100}\right) \Rightarrow h = 200\,m$$

39. (a) When water falls from a height, loss of potential energy causes rise in temperature.

40. (a) $W = JQ \Rightarrow mg\,h = J(m.c.\Delta\theta)$

$$\Rightarrow \Delta\theta = \frac{g\,h}{Jc} = 0.0023\,h = 0.0023\times84 = 0.196°C$$

41. (a) Suppose *m′ kg* ice melts out of *m kg* then by using

$W = JQ \Rightarrow mgh = J(m'L)$. Hence fraction of ice melts

$$= \frac{m'}{m} = \frac{gh}{JL} = \frac{9.8\times1000}{4.18\times80} = \frac{1}{33}$$

42. (b) $J = \dfrac{W}{Q} = \dfrac{Joule}{cal}$

43. (d) $W = JQ \Rightarrow (2m)gh = J\times m'c\Delta\theta$

$$\Rightarrow 2\times5\times10\times10 = 4.2(2\times1000\times\Delta\theta)$$

$$\Rightarrow \Delta\theta = 0.1190°C = 0.12°C$$

44. (b) $W = JQ \Rightarrow \dfrac{1}{2}\left(\dfrac{1}{2}mV^2\right) = J\times mS\Delta\theta \Rightarrow \Delta\theta = \dfrac{V^2}{4\,JS}$

45. (c) '*J*' is a conversion

46. (c) $\Delta\theta = 0.0023\,h = 0.0023\times 210 = 0.483°C \approx 0.49°C$.

47. (a) According to energy conservation, change in kinetic energy appears in the form of heat (thermal energy).

$\Rightarrow$ *i.e.* Thermal energy $= \frac{1}{2}m(v_1^2 - v_2^2)$ $\left[\because \underset{(Joule)}{W} = \underset{(Joule)}{Q}\right]$

$= \frac{1}{2}(100\times 10^{-3})(10^2 - 5^2) = 3.75\ J$

48. (b) Work done to raise the temperature of 100 gm water through $10°C$ is

$W = JQ = 4.2\times(100\times 10^{-3}\times 1000\times 10) = 4200\ J$

49. (b) Among all the option, latent heat of steam is highest.

50. (a) $\Delta\theta = 0.0023\,h = 0.0023\times 100 = 0.23°C$

51. (c) Since specific heat of lead is given in *Joules*, hence use $W = Q$ instead of $W = JQ$.

$\Rightarrow \frac{1}{2}\times\left(\frac{1}{2}mv^2\right) = m.c.\Delta\theta \Rightarrow \Delta\theta = \frac{v^2}{4c} = \frac{(300)^2}{4\times 150} = 150°C$.

52. (d) At boiling point, vapour pressure becomes equal to the external pressure.

53. (c) When pressure increases boiling point also increases.

54. (b) Calorimeters are made by conducting materials.

55. (b) Triple point of water is 273.16 K.

56. (a)

57. (d) $W = JQ \Rightarrow W = 4.2\times 200 = 840\ J$.

58. (d) Temperature of mixture $\theta = \frac{m_1c_1\theta_1 + m_2c_2\theta_2}{m_1c_1 + m_2\theta_2}$

$\Rightarrow 32 = \frac{m_1\times 0.2\times 40 + 100\times 0.5\times 20}{m_1\times 0.2 + 100\times 0.5} \Rightarrow m_1 = 375\ gm$

59. (d) Suppose $m\ gm$ ice melted, then heat required for its melting $= mL = m\times 80\ cal$

Heat available with steam for being condensed and then brought to $0°C$

$= 1\times 540 + 1\times 1\times(100-0) = 640\ cal$

$\Rightarrow$ Heat lost = Heat taken

$\Rightarrow 640 = m\times 80 \Rightarrow m = 8\ gm$

Short trick: You can remember that amount of steam (m') at $100°C$ required to melt $m\ gm$ ice at $0°C$ is $m' = \frac{m}{8}$.

Here, $m = 8\times m' = 8\times 1 = 8\ gm$

60. (b) For water and ice mixing $\theta_{mix} = \frac{m_W\theta_W - \frac{m_iL_i}{c_W}}{m_i + m_W}$

$= \frac{20\times 40 - \frac{5\times 80}{1}}{5+20} = 16°C$

61. (c) $\theta_{mix} = \frac{m_W\theta_W - \frac{m_iL_i}{c_W}}{m_i + m_W}$

$\because m_i = m_W \Rightarrow \theta_{mix} = \frac{\theta_W - \frac{L_i}{c_W}}{2} = \frac{80 - \frac{336}{4.2}}{2} = 0°C$

62. (a) When the relative humidity is low (approx. 25%), the evaporation from our body is faster. Thus we feel colder.

63. (a) $c = \frac{Q}{m.\Delta\theta} \to \frac{J}{kg\times °C}$

64. (a) $\theta_{mix} = \frac{\theta_W - \frac{L_i}{c_W}}{2} = \frac{80 - \frac{80}{1}}{2} = 0$

65. (a) Freezing point of water decreases when pressure increases, because water expands on solidification while "except water" for other liquid freezing point increases with increase in pressure.

Since the liquid in question is water. Hence, it expands on freezing.

66. (c) Partial pressure of water vapour $P_W = 0.012\times 10^5\ Pa$,

Vapour pressure of water $P_V = 0.016\times 10^5\ Pa$.

The relative humidity at a given temperature is given by

$= \frac{\text{Partial pressure of water vapour}}{\text{Vapour pressure of water}}$

$= \frac{0.012\times 10^5}{0.016\times 10^5} = 0.75 = 75\%$

67. (a) $W = JQ \Rightarrow \frac{1}{2}\left(\frac{1}{2}Mv^2\right) = J(m.c.\Delta\theta)$

$\Rightarrow \frac{1}{4}\times 1\times(50)^2 = 4.2[200\times 0.105\times\Delta\theta] \Rightarrow \Delta\theta = 7.1°C$

68. (a) $536\frac{cal}{gm} = \frac{536\times 4.2\ J}{10^{-3}\ kg} = 2.25\times 10^6\ J/kg$

69. (a) Water has maximum specific heat.

70. (a) $\theta_{mix} = \frac{\theta_W - \frac{L_i}{C_W}}{2} = \frac{100 - \frac{80}{1}}{2} = 10°C$

71. (b) Suppose $m\ kg$ of ice melts then by using $\underset{(Joules)}{W} = \underset{(Joules)}{H}$

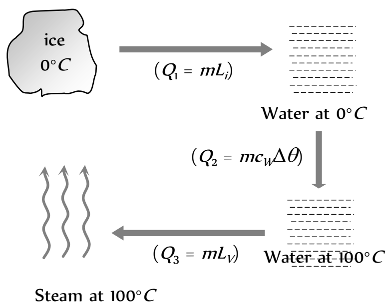
$\Rightarrow Mgh = mL \Rightarrow 3.5\times 10\times 2000 = m\times 3.5\times 10^5$

$\Rightarrow m = 0.2\ kg = 200\ gm$

72. (d) $\theta_{mix} = \frac{m_W\theta_W - \frac{m_iL_i}{S_W}}{m_i + m_W} = \frac{300\times 25 - \frac{100\times 80}{1}}{100+300} = -1.25^oC$

Which is not possible. Hence $\theta_{mix} = 0^oC$

73. (c) Ice ($0°C$) converts into steam ($100°C$) in following three steps.



Total heat required $Q = Q_1 + Q_2 + Q_3$

$= 5 \times 80 + 5 \times 1 \times (100 - 0) + 5 \times 540 = 3600\, cal$

74. (a) Let L be the latent heat and using principle of calorimetry.

$2L + 2\ (100 - 54.3) = 40 \times (54.3 - 25.3)$

$\Rightarrow L = 540.3$ *cal/gm.*

75. (d) $\theta_{\text{mix}} = \dfrac{m_W\theta_W - \dfrac{m_i L_i}{c_W}}{m_i + m_W} = \dfrac{100 \times 50 - 10 \times \dfrac{80}{1}}{10 + 100} \approx 38.2°C$

76. (b) Let the final temperature be $T°C$.

Total heat supplied by the three liquids in coming down to $0°C$ =

$m_1c_1T_1 + m_2c_2T_2 + m_3c_3T_3$ ..... (i)

Total heat used by three liquids in raising temperature from $0°C$ to $T°C$

= $m_1c_1T + m_2c_2T + m_3c_3T$ .....(ii)

By equating (i) and (ii) we get

$(m_1c_1 + m_2c_2 + m_3c_3)T$

= $m_1c_1T_1 + m_2c_2T_2 + m_3c_3T_3$

$\Rightarrow T = \dfrac{m_1c_1T_1 + m_2c_2T_2 + m_3c_3T_3}{m_1c_1 + m_2c_2 + m_3c_3}$.

77. (c) At triple point all the phases co-exist

78. (b) $c = \dfrac{Q}{m.\Delta\theta}$; as $\Delta\theta = 0$, hence c becomes ∞.

79. (d) Let final temperature of water be θ

Heat taken = Heat given

$110 \times 1\ (\theta - 10) + 10\ (\theta - 10) = 220 \times 1\ (70 - \theta)$

$\Rightarrow \theta = 48.8°C \approx 50°C$.

80. (c) We know that thermal capacity of a body expressed in calories is equal to water equivalent of the body expressed in grams.

81. (b) $\theta_{mix} = \dfrac{m_1c_1\theta_1 + m_2c_2\theta_2}{m_1c_1 + m_2c_2} = \dfrac{m\,s\,(2t) + 1.5\,(m\,s) \times \dfrac{t}{3}}{m\,s + 1.5\,(m\,s)} = t$

82. (d) We know that when solid carbondioxide is heated, it becomes vapour directly without passing through its liquid phase. Therefore it is called dry ice.

Critical Thinking Questions

1. (c) Due to volume expansion of both mercury and flask, the change in volume of mercury relative to flask is given by $\Delta V = V_0[\gamma_L - \gamma_g]\Delta\theta = V[\gamma_m - 3\alpha_g]\Delta\theta$

$= 50\,[180 \times 10^{-6} - 3 \times 9 \times 10^{-6}](38 - 18) = 0.153\, cc$

2. (a) $\gamma_{real} = \gamma_{app.} + \gamma_{vessel}$

So $(\gamma_{app.} + \gamma_{vessel})_{glass} = (\gamma_{app.} + \gamma_{vessel})_{steel}$

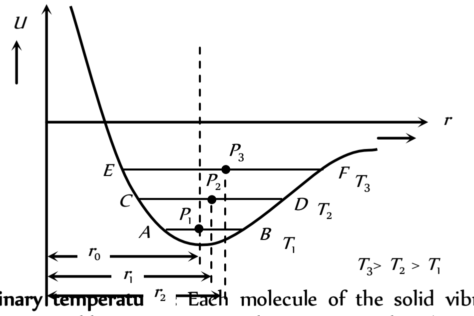
$\Rightarrow 153 \times 10^{-6} + (\gamma_{vessel})_{glass} = (144 \times 10^{-6} + \gamma_{vessel})_{steel}$

Further, $(\gamma_{vessel})_{steel} = 3\alpha = 3 \times (12 \times 10^{-6}) = 36 \times 10^{-6}/°C$

$\Rightarrow 153 \times 10^{-6} + (\gamma_{vessel})_{glass} = 144 \times 10^{-6} + 36 \times 10^{-6}$

$\Rightarrow (\gamma_{vessel})_{glass} = 3\alpha = 27 \times 10^{-6}/°C \Rightarrow \alpha = 9 \times 10^{-6}/°C$

3. (d) The expansion of solids can be well understood by potential energy curve for two adjacent atoms in a crystalline solid as a function of their internuclear separation (r).



At ordinary temperature : Each molecule of the solid vibrate about it's equilibrium position P_1 between A and B (r_1 is the equilibrium distance of it from some other molecule)

At high temperature : Amplitude of vibration increase ($C \leftrightarrow D$ and $E \leftrightarrow F$). Due to asymmetry of the curve, the equilibrium positions (P_2 and P_3) of molecule displaced. Hence it's distance from other molecule increases ($r_3 > r_2 > r_1$).

Thus, on raising the temperature, the average equilibrium distance between the molecules increases and the solid as a whole expands.

4. (c) Initial diameter of tyre = (1000 – 6) *mm* = 994 *mm*, so initial radius of tyre $R = \dfrac{994}{2} = 497\, mm$

and change in diameter ΔD = 6 *mm* so $\Delta R = \dfrac{6}{2} = 3\, mm$

After increasing temperature by $\Delta\theta$ tyre will fit onto wheel

Increment in the length (circumference) of the iron tyre

$\Delta L = L \times \alpha \times \Delta\theta = L \times \dfrac{\gamma}{3} \times \Delta\theta$ [As $\alpha = \dfrac{\gamma}{3}$]

$2\pi\Delta R = 2\pi R\left(\dfrac{\gamma}{3}\right)\Delta\theta \Rightarrow \Delta\theta = \dfrac{3}{\gamma}\dfrac{\Delta R}{R} = \dfrac{3 \times 3}{3.6 \times 10^{-5} \times 497}$

$\Rightarrow \Delta\theta = 500^o C$

5. (b) Due to volume expansion of both liquid and vessel, the change in volume of liquid relative to container is given by $\Delta V = V_0[\gamma_L - \gamma_g]\Delta\theta$

Given $V_0 = 1000\ cc$, $\alpha_g = 0.1 \times 10^{-4}/°C$

$\therefore \gamma_g = 3\alpha_g = 3 \times 0.1 \times 10^{-4}\ /°C = 0.3 \times 10^{-4}\ /°C$

$\therefore \Delta V = 1000\ [1.82 \times 10^{-4} - 0.3 \times 10^{-4}] \times 100 = 15.2\ cc$

6. (a) With temperature rise (same $25°C$ for both), steel scale and copper wire both expand. Hence length of copper wire *w.r.t.* steel scale or apparent length of copper wire after rise in temperature

$L_{app} = L'_{cu} - L'_{steel} = [L_0(1 + \alpha_{Cu}\Delta\theta) - L_0(1 + \alpha_s\Delta\theta)$

$\Rightarrow L_{app} = L_0(\alpha_{Cu} - \alpha_s)\Delta\theta$

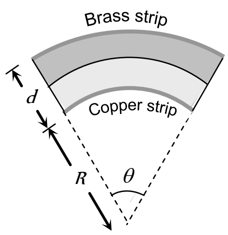
$= 80(17 \times 10^{-6} - 11 \times 10^{-6}) \times 20 = 80.0096\ cm$

7. (b, d) Let L_0 be the initial length of each strip before heating.

Length after heating will be

$L_B = L_0(1 + \alpha_B\Delta T) = (R + d)\theta$

$L_C = L_0(1 + \alpha_C\Delta T) = R\theta$



$$\Rightarrow \frac{R+d}{R} = \frac{1+\alpha_B \Delta T}{1+\alpha_C \Delta T}$$

$$\Rightarrow 1+\frac{d}{R} = 1+(\alpha_B-\alpha_C)\Delta T$$

$$\Rightarrow R = \frac{d}{(\alpha_B-\alpha_C)\Delta T} \Rightarrow R \propto \frac{1}{\Delta T} \text{ and } R \propto \frac{1}{(\alpha_B-\alpha_C)}$$

8. (d) Thermostat is used in electric apparatus like refrigerator, Iron *etc* for automatic cut off. Therefore for metallic strips to bend on heating their coefficient of linear expansion should be different.

9. (c) As the coefficient of cubical expansion of metal is less as compared to the coefficient of cubical expansion of liquid, we may neglect the expansion of metal ball. So when the ball is immersed in alcohol at 0°C, it displaces some volume V of alcohol at $0°C$ and has weight W_1.

$\therefore \quad W_1 = W_0 - V\rho_0 g$

where W_0 = weight of ball in air

Similarly, $W_2 = W_0 - V\rho_{50} g$

where ρ_0 = density of alcohol at 0°C

and ρ_{50} = density of alcohol at 50°C

As $\rho_{50} < \rho_0$, $\Rightarrow W_2 > W_1$ or $W_1 < W_2$

10. (d) $V = V_0(1 + \gamma\Delta\theta)$ $\Rightarrow$ Change in volume

$$V - V_0 = \Delta V = A.\Delta l = V_0\gamma\Delta\theta$$

$$\Rightarrow \Delta l = \frac{V_0.\Delta\theta}{A} = \frac{10^{-6}\times 18\times 10^{-5}\times(100-0)}{0.004\times 10^{-4}}$$

$= 45 \times 10^{-3}\ m = 4.5\ cm$

11. (b) Loss of weight at $27°C$ is

$= 46 - 30 = 16 = V_1 \times 1.24\ \rho_l \times g$ ...(i)

Loss of weight at $42°C$ is

$= 46 - 30.5 = 15.5 = V_2 \times 1.2\ \rho_l \times g$ ...(ii)

Now dividing (i) by (ii), we get $\frac{16}{15.5} = \frac{V_1}{V_2}\times\frac{1.24}{1.2}$

But $\frac{V_2}{V_1} = 1 + 3\alpha\,(t_2 - t_1) = \frac{15.5\times 1.24}{16\times 1.2} = 1.001042$

$\Rightarrow 3\alpha\,(42° - 27°) = 0.001042 \Rightarrow \alpha = 2.316 \times 10^{-5}/°C$.

12. (b) Substances are classified into two categories

(i) water like substances which expand on solidification.

(ii) CO_2 like (Wax, Ghee *etc.*) which contract on solidification.

Their behaviour regarding solidification is opposite.

Melting point of ice decreases with rise of temp but that of wax etc increases with increase in temperature. Similarly ice starts forming from top downwards whereas wax starts its formation from bottom.

13. (b) Heat lost in t sec = mL or heat lost per sec = $\frac{mL}{t}$. This must be the heat supplied for keeping the substance in molten state per sec.

$$\therefore \quad \frac{mL}{t} = P \quad \text{or} \quad L = \frac{Pt}{m}$$

14. (a) Heat is lost by steam in two stages (i) for change of state from steam at $100°C$ to water at $100°C$ is $m \times 540$ (ii) to change water at $100°C$ to water at $80°C$ is $m \times 1 \times (100 - 80)$, where m is the mass of steam condensed.

Total heat lost by steam is m × 540 + m × 20 = 560 m (*cals*)

Heat gained by calorimeter and its contents is

= (1.1 + 0.02) × (80 – 15) = 1.12 × 65 *cals.*

using Principle of calorimetery, Heat gained = heat lost

$\therefore$ 560 m = 1.12 × 65, m = 0.130 *gm*

15. (b) Initially ice will absorb heat to raise it's temperature to $0°C$ then it's melting takes place

If m_i = Initial mass of ice, m_i' = Mass of ice that melts and m_w = Initial mass of water

By Law of mixture Heat gained by ice = Heat lost by water $\Rightarrow$

$m_i \times c \times (20) + m_i' \times L = m_W c_W [20]$

$\Rightarrow 2\times 0.5(20) + m_i'\times 80 = 5\times 1\times 20 \Rightarrow m_i' = 1 kg$

So final mass of water = Initial mass of water + Mass of ice that melts = 5 + 1= 6 *kg.*

16. (a) Heat gained by the water = (Heat supplied by the coil) – (Heat dissipated to environment)

$$\Rightarrow mc\,\Delta\theta = P_{Coil}\,t - P_{Loss}\,t$$

$$\Rightarrow 2\times 4.2\times 10^3\times(77-27) = 1000\,t - 160\,t$$

$$\Rightarrow t = \frac{4.2\times 10^5}{840} = 500\ sec = 8\ min\ 20\ sec$$

17. (a) If mass of the bullet is m *gm*,

then total heat required for bullet to just melt down

$Q_1 = m\,c\,\Delta\theta + m\,L = m \times 0.03\ (327 - 27) + m \times 6$

$= 15\ m\ cal = (15m\times 4.2)J$

Now when bullet is stopped by the obstacle, the loss in its mechanical energy $= \frac{1}{2}(m\times 10^{-3})v^2 J$

(As $m\ gm = m\times 10^{-3} kg$)

As 25% of this energy is absorbed by the obstacle,

The energy absorbed by the bullet

$$Q_2 = \frac{75}{100}\times\frac{1}{2}mv^2\times 10^{-3} = \frac{3}{8}mv^2\times 10^{-3}\ J$$

Now the bullet will melt if $Q_2 \geq Q_1$

i.e. $\frac{3}{8}mv^2\times 10^{-3} \geq 15m\times 4.2 \Rightarrow v_{min} = 410\ m/s$

18. (c) Energy = $\frac{1}{2}mv^2 = mc\,\Delta\theta$; $\Rightarrow \Delta\theta \propto v^2$

Temperature does not depend upon the mass of the balls.

19. (c) Heat gain = heat lost

$C_A(16 - 12) = C_B\,(19 - 16) \Rightarrow \frac{C_A}{C_B} = \frac{3}{4}$

and $C_B(23 - 19) = C_C\,(28 - 23) \Rightarrow \frac{C_B}{C_C} = \frac{5}{4}$

$\Rightarrow \frac{C_A}{C_C} = \frac{15}{16}$ ...(i)

If θ is the temperature when A and C are mixed then,

$C_A(\theta - 12) = C_C(28 - \theta) \Rightarrow \frac{C_A}{C_C} = \frac{28-\theta}{\theta - 12}$ ...(ii)

On solving equation (i) and (ii) $\theta = 20.2°C$.

20. (b) Suppose m *kg* steam required per hour

Heat released by steam in following three steps

(i) When $150°C$ steam $\xrightarrow[Q_1]{} 100°C$ steam

$Q = mc_{\_}\,\Delta\theta = m \times 1\,(150 - 100) = 50\ m\ cal$

(ii) When $150°C$ steam $\xrightarrow[Q_2]{} 100°C$ water

$Q = mL = m \times 540 = 540\ m\ cal$

(iii) When $100°C$ water $\xrightarrow[Q_2]{} 90°C$ water

$Q = mc\,\Delta\theta = m \times 1 \times (100 - 90) = 10\ m\ cal$

Hence total heat given by the steam $Q = Q + Q + Q = 600\ m cal$ ... (i)

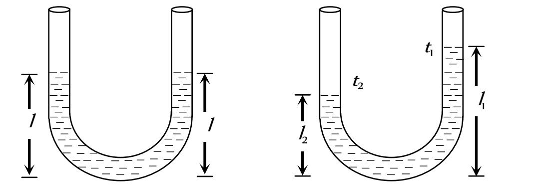
Heat taken by 10 *kg* water

$Q' = mc_W\Delta\theta = 10 \times 10^3 \times 1 \times (80 - 20) = 600 \times 10^3\, cal$

Hence $Q = Q' \Rightarrow 600\ m = 600 \times 10$

$\Rightarrow m = 10\ gm = 1 kg$.

21. (a) Suppose, height of liquid in each arm before rising the temperature is l.



With temperature rise height of liquid in each arm increases *i.e.* $l_1 > l$ and $l_2 > l$

Also $l = \frac{l_1}{1+\gamma\, t_1} = \frac{l_2}{1+\gamma\, t_2}$

$\Rightarrow l_1 + \gamma\, l_1 t_2 = l_2 + \gamma\, l_2 t_1 \Rightarrow \gamma = \frac{l_1 - l_2}{l_2 t_1 - l_1 t_2}$

22. (c) $V = V_0(1 + \gamma\Delta\theta)$

$L^3 = L_0(1+\alpha_1\Delta\theta)L_0^2(1+\alpha_2\Delta\theta)^2 = L_0^3(1+\alpha_1\Delta\theta)(1+\alpha_2\Delta\theta)^2$

Since $L_0^3 = V_0$ and $L^3 = V$

Hence $1 + \gamma\Delta\theta = (1+\alpha_1\Delta\theta)(1+\alpha_2\Delta\theta)^2$

$\cong (1+\alpha_1\Delta\theta)(1+2\alpha_2\Delta\theta) \cong (1+\alpha_1\Delta\theta + 2\alpha_2\Delta\theta)$

$\Rightarrow \gamma = \alpha_1 + 2\alpha_2$

23. (d) $(OR)^2 = (PR)^2 - (PO)^2 = l^2 - \left(\frac{l}{2}\right)^2$

$= [l(1+\alpha_2 t)]^2 - \left[\frac{l}{2}(1+\alpha_1 t)\right]^2$

$l^2 - \frac{l^2}{4} = l^2(1+\alpha_2^2t^2 + 2\alpha_2 t) - \frac{l^2}{4}(1+\alpha_1^2t^2 + 2\alpha_1 t)$

Neglecting $\alpha_2^2t^2$ and $\alpha_1^2t^2$

$0 = l^2(2\alpha_2 t) - \frac{l^2}{4}(2\alpha_1 t) \Rightarrow 2\alpha_2 = \frac{2\alpha_1}{4} \Rightarrow; \alpha_1 = 4\alpha_2$

24. (c) It is given that the volume of air in the flask remains the same. This means that the expansion in volume of the vessel is exactly equal to the volume expansion of mercury.

i.e., $\Delta V_g = \Delta V_L$ or $V_g\gamma_g\Delta\theta = V_L\gamma_L\Delta\theta$

$\therefore V_L = \frac{V_g\gamma_g}{\gamma_L} = \frac{1000 \times (3 \times 9 \times 10^{-6})}{1.8 \times 10^{-4}} = 150\, cc$

25. (c) Heat given by water $Q_1 = 10 \times 10 = 100\, cal.$

Heat taken by ice to melt

$Q = 10 \times 0.5 \times [0 - (-20)] + 10 \times 80 = 900\ cal$

As $Q_1 < Q_2$, so ice will not completely melt and final temperature $= 0°C$.

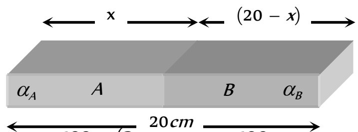
As heat given by water in cooling up to $0°C$ is only just sufficient to increase the temperature of ice from $-20°C$ to $0°C$, hence mixture in equilibrium will consist of 10 *gm* ice and 10 *gm* water at $0°C$.

26. (b) $\Delta L = L_0\, \alpha\Delta\theta$

Rod A : $0.075 = 20 \times \alpha \times 100 \Rightarrow \alpha_A = \frac{75}{2} \times 10^{-6}\ /°C$

rod B : $0.045 = 20 \times \alpha \times 100 \Rightarrow \alpha_B = \frac{45}{2} \times 10^{-6}\ /°C$

For composite rod : x *cm* of A and $(20 - x)$ *cm* of B we have

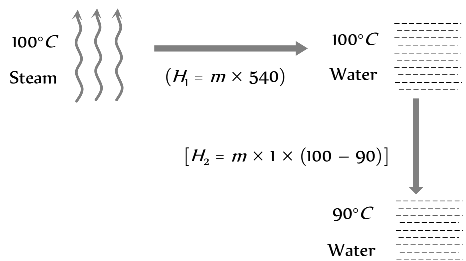


$0.060 = x\,\alpha \times 100 + (20 - x)\,\alpha \times 100$

$= x\left[\frac{75}{2} \times 10^{-6} \times 100 + (20 - x) \times \frac{45}{2} \times 10^{-6} \times 100\right]$

On solving we get $x = 10\ cm$.

27. (a) Let m *gm* of steam get condensed into water (By heat loss). This happens in following two steps.



Heat gained by water ($20°C$) to raise it's temperature upto 90° $= 22 \times 1 \times (90 - 20)$

Hence, in equilibrium heat lost = Heat gain

$\Rightarrow m \times 540 + m \times 1 \times (100 - 90) = 22 \times 1 \times (90 - 20)$

$\Rightarrow m = 2.8\ gm$

The net mass of the water present in the mixture $= 22 + 2.8 = 24.8\ gm.$

Graphical Questions

1. (b) Relation between Celsius and Fahrenheit scale of temperature is $\frac{C}{5} = \frac{F-32}{9} \Rightarrow C = \frac{5}{9}F - \frac{160}{9}$

Equating above equation with standard equation of line

$y = mx + c$ we get slope of the line AB is $m = \frac{5}{9}$

2. (c) Since in the region AB temperature is constant therefore at this temperature phase of the material changes from solid to liquid and $(H_2 - H_1)$ heat will be absorb by the material. This heat is known as the heat of melting of the solid.

Similarly in the region CD temperature is constant therefore at this temperature phase of the material changes from liquid to gas and $(H_4 - H_3)$ heat will be absorb by the material. This heat as known as the heat of vaporisation of the liquid.

3. (a) Initially, on heating temperature rises from $-10°C$ to $0°C$. Then ice melts and temperature does not rise. After the whole ice has melted, temperature begins to rise until it reaches $100°C$. Then it becomes constant, as at the boiling point will not rise.

4. (a) The volume of matter in portion AB of the curve is almost constant and pressure is decreasing. These are the characteristics of liquid state.

5. (d) Let the quantity of heat supplied per minute be Q. Then quantity of heat supplied in 2 *min* $= mC(90-80)$

In 4 *min*, heat supplied $= 2m\,C(90-80)$

$\therefore\ 2m\,C(90-80) = m\,L \Rightarrow \frac{L}{C} = 20$

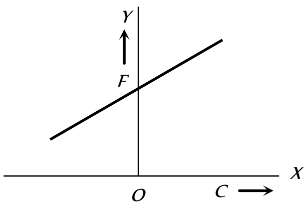
6. (b) In the given graph CD represents liquid state.

7. (a) Density of water is maximum at $4°C$ and is less on either side of this temperature.

8. (a) We know that, $\frac{C}{100} = \frac{F-32}{180}$ or $F = \frac{9}{5}C + 32$

Equation of straight line is, $y = mx + c$

Hence, $m = (9/5)$, positive and $c = 32$ positive. The graph is shown in figure.



9. (a)

$$\frac{C}{5} = \frac{F-32}{9} \Rightarrow$$

$C = \left(\frac{5}{9.}\right)F - \frac{20}{3}$. Hence graph between $°C$ and $°F$ will be a straight line with positive slope and negative intercept.

10. (bc) The horizontal parts of the curve, where the system absorbs heat at constant temperature must depict changes of state. Here the latent heats are proportional to lengths of the horizontal parts. In the sloping parts, specific heat capacity is inversely proportional to the slopes.

11. (c) Since specific heat = 0.6 *kcal/gm* × °*C* = 0.6 *cal/gm* ×°*C*

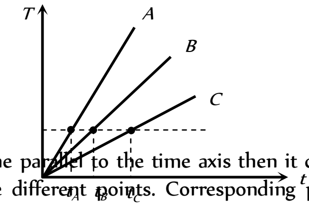
From graph it is clear that in a minute, the temperature is raised from $0°C$ to $50°C$.

$\Rightarrow$ Heat required for a minute = 50 × 0.6 × 50 = 1500 *cal.*

Also from graph, Boiling point of wax is $200°C$.

12. (c)

13. (c) Substances having more specific heat take longer time to get heated to a higher temperature and longer time to get cooled.



If we draw a line parallel to the time axis then it cuts the given graphs at three different points. Corresponding points on the times axis shows that

$t_C > t_B > t_A \Rightarrow C_C > C_B > C_A$

14. (c) From given curve,

Melting point for $A = 60°C$

and melting point for $B = 20°C$

Time taken by A for fusion $= (6-2) = 4$ minute

Time taken by B for fusion $= (6.5-4) = 2.5$ minute

Then $\frac{H_A}{H_B} = \frac{6 \times 4 \times 60}{6 \times 2.5 \times 60} = \frac{8}{5}$.

Assertion and Reason

1. (a) With rise in pressure melting point of ice decreases. Also ice contracts on melting

2. (c) Celsius scale was the first temperature scale and Fahrenheit is the smallest unit measuring temperature.

3. (e) Melting is associated with increasing of internal energy without change in temperature. In view of the reason being correct the amount of heat absorbed or given out during change of state is expressed $Q = mL$, where m is the mass of the substance and L is the latent heat of the substance.

4. (a) The temperature of land rises rapidly as compared to sea because of specific heat of land is five times less than that of sea water. Thus, the air above the land become hot and light so rises up so because of pressure drops over land. To compensate the drop of pressure, the cooler air starts from sea starts blowing towards lands, so, setting up sea breeze. During night land as well sea radiate heat energy. The temperature of land falls more rapidly as compared to sea water, as sea water consists of higher specific heat capacity. The air above sea water being warm and light rises up and to take its place the cold air from land starts blowing towards sea and so et up breeze.

5. (a) Linear expansion for brass $(19 \times 10^{-4}) >$ linear expansion for steel (11×10^{-4}). On cooling the disk shrinks to a greater extent than the hole and hence it will get loose.

6. (a) As, $\gamma = \frac{\Delta V}{V \Delta T}$ *i.e.*, units of coefficient of volume expansion is K.

7. (c) The relation between F and C scale is, $\frac{C}{5} = \frac{F-32}{9}$. If $F = C$

$\Rightarrow C = -40°C$ *i.e.*, at $-40°$ the Centigrade and Fahrenheit thermometers reads the same.

8. (a) As $\beta = 2\alpha$ and $\gamma = 3\alpha$, *i.e.*, coefficient of volume expansion of solid is three time coefficient of linear expansion and 1.5 times the coefficient of superficial expansion, on heating a solid iron ball, percentage increase in its volume is largest.

9. (a) Water has maximum density at $4°C$. On heating above $4°C$ or cooling below $4°C$, density of water decreases and its volume increases. Therefore, water overflows in both the cases.

10. (b) The Latent heat of fusion of ice is amount of heat required to convert unit mass of ice at $0°C$ into water at $0°C$. For fusion of ice

$$L = 80\, cal/gm = 80000\, cal/gm = 8000 \times 4.2\ J/kg$$

$$= 336000\ J/kg\ .$$

11. (a) When two bodies at temperature T_1 and T_2 are brought in thermal contact, they do settle to the mean temperature $(T_1 + T_2)/2$. They will do so, in case the two bodies were of same mass and material *i.e.*, same thermal capacities. In other words, the two bodies may be having different thermal capacities, that's why they do not settle to the mean temperature, when brought together.

12. (d) Specific heat of a body is the amount of heat required to raise the temperature of unit mass of the body through unit degree. When mass of a body is less than unity, then its thermal capacity is less than its specific heat and vice-versa.

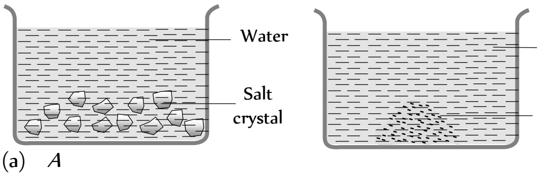
13. (a) Water would evaporate quickly because there is no atmosphere on moon, due to which surface temperature of moon is much higher than earth (Maximum surface temperature of moon is $123°C$).

14. (d) The potential energy of water molecules is more. The heat given to melt the ice at $0°C$ is used up in increasing the potential energy of water molecules formed at $0°C$.

Thermometry, Thermal Expansion and Calorimetry

SET Self Evaluation Test -12

1. Out of the following, in which vessel will the temperature of the solution be higher after the salt is completely dissolved.



(a) *A*
(b) *B*
(c) Equal in both
(d) Information is not sufficient

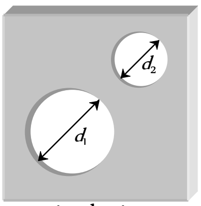
2. Fire is extinguished more effectively by
(a) Hot water (b) Cold water
(c) Equally by both (d) Ice

3. An ideal thermometer should have
(a) Large heat capacity (b) Medium heat capacity
(c) Small heat capacity (d) Variable heat capacity

4. A steel meter scale is to be ruled so that millimeter intervals are accurate within about 5×10^{-5} *mm* at a certain temperature. The maximum temperature variation allowable during the ruling is (Coefficient of linear expansion of steel $=10\times10^{-6}K^{-1}$)
(a) $2°C$ (b) $5°C$
(c) $7°C$ (d) $10°C$

5. During illness an 80 *kg* man ran a fever of $102.2°F$ instead of normal body temperature of $98.6°F$. Assuming that human body is mostly water, how much heat is required to raise his temperature by that amount
(a) 100 *kcal* (b) 160 *kcal*
(c) 50 *kcal* (d) 92 *kcal*

6. Two holes of unequal diameters d_1 and d_2 $(d_1 > d_2)$ are cut in a metal sheet. If the sheet is heated



(a) Both d_1 and d_2 will decrease
(b) Both d_1 and d_2 will increase
(c) d_1 will increase, d_2 will decrease
(d) d_1 will decrease, d_2 will increase

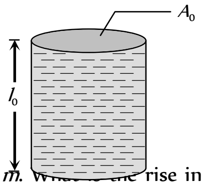
7. If earth suddenly stops rotating about its own axis, the increase in it's temperature will be
(a) $\frac{R^2\omega^2}{5Js}$ (b) $\frac{R^2\omega^2}{Js}$
(c) $\frac{Rm\omega^2}{5Js}$ (d) None of these

8. Latent heat of ice is 80 *cal/gm*. A man melts 60 *g* of ice by chewing in 1 *minute*. His power is
(a) 4800 *W* (b) 336 *W*
(c) 1.33 *W* (d) 0.75 *W*

9. A faulty thermometer has its lower fixed point marked as $-10°C$ and upper fixed point marked as 110° and upper fixed point marked as 110°. If the temperature of the body shown in this scale is 62°, the temperature shown on the Celsius scale is
(a) $72° C$ (b) $82° C$
(c) $60° C$ (d) $42° C$

10. If there are no heat losses, the heat released by the condensation of *x gm* of steam at $100°C$ into water at $100°C$ can be used to convert *y gm* of ice at $0°C$ into water at $100°C$. Then the ratio *y* : *x* is nearly
(a) 1 : 1 (b) 2.5 : 1
(c) 2 : 1 (d) 3 : 1

11. The figure shows a glass tube (linear co-efficient of expansion is α) completely filled with a liquid of volume expansion co-efficient γ. On heating length of the liquid column does not change. Choose the correct relation between γ and α



(a) $\gamma = \alpha$
(b) $\gamma = 2\alpha$
(c) $\gamma = 3\alpha$ [EAMCET 2001]
(d) $\gamma = \frac{\alpha}{3}$

12. Water falls from a height 500*m*. What is the rise in temperature of water at bottom if whole energy remains in the water
(a) $0.96°C$ (b) $1.02°C$
(c) $1.16°C$ (d) $0.23°C$

13. A steel ball of mass 0.1 *kg* falls freely from a height of 10 *m* and bounces to a height of 5.4*m* from the ground. If the dissipated energy in this process is absorbed by the ball, the rise in its temperature is

(Specific heat of steel $= 460\ Joule-kg^{-1}°C^{-1}$, $g = 10\ ms^{-2}$)

[EAMCET (Med.) 2000]

(a) $0.01°C$ (b) $0.1°C$
(c) $1°C$ (d) $1.1°C$

14. 1*gm* of ice at $0°C$ is mixed with 1*gm* of water at $100°C$ the resulting temperature will be [AIIMS 1994]
(a) $5°C$ (b) $0°C$
(c) $10°C$ (d) ∞

15. The amount of heat required to change 1 *gm* ($0°C$) of ice into water of $100°C$, is [RPMT 1999]
(a) 716 *cal* (b) 500 *cal*
(c) 180 *cal* (d) 100 *cal*

AS Answers and Solutions

(SET -12)

1. (b) When salt crystals dissolves, crystal lattice is destroyed. The process requires a certain amount of energy (latent heat) which is taken from the water.

In vessel (B), a part of intermolecular bonds has already been destroyed in crushing the crystal. Hence less energy is require to dissolve the powder and the water will be at higher temperature.

2. (a) Fire is extinguished by the vaporisation do water which lowers the temperature of the burning body. Further, the water vapour envelops the body, keeping oxygen away. Hot water evaporates more than cold water as

3. (c) The thermometer has to attain the temperature of the body. To do this, it should draw as little heat from the body as possible, so that the existing temperature of the body is not disturbed.

4. (b) As we know $\alpha = \frac{\Delta L}{L_0 \Delta\theta} \Rightarrow \Delta\theta = \frac{\Delta L}{\alpha L_0} = \frac{5\times 10^{-5}}{10\times 10^{-6}\times 1} = 5°C$

5. (b) Since 102.2°$F \rightarrow$ 39°C and 98.6°$F \rightarrow$ 39°C

Hence ΔQ = m. s. ΔQ = 80 × 1000 × (39 − 37)

= 16 × 10· cal = 160 $kcal$.

6. (b) If the sheet is heated then both d and d will increase since the thermal expansion of isotropic solid is similar to true photographic enlargement.

7. (a) $W = JQ \Rightarrow \frac{1}{2} I\omega^2 = J(MS\Delta\theta) \Rightarrow \frac{1}{2}\left(\frac{2}{5}MR^2\right)\omega^2$

$= J(MS\Delta\theta) \Rightarrow \Delta\theta = \frac{R^2\omega^2}{5Js}$

8. (b) $W = JQ = J(mL) \Rightarrow P\times t = J(mL) \Rightarrow P = J\left(\frac{m}{t}\right)L$;

where $\frac{m}{t}$ = rate of melting of ice by chewing

$= \frac{60\ gm}{min} = \frac{1\ gm}{sec} \Rightarrow P$ = 4.2 × 1 × 80 = 336 W.

9. (c) $\frac{X-L}{U-L} = \frac{C}{100} \Rightarrow \frac{62-(10)}{110-(-10)} = \frac{C}{100}$ (C = 60°C)

10. (d) Heat released to convert x gm of steam at 100°C to water at 100°C is x × 540 $cals$.

If y gm of ice is converted from 0°C to water at 100°C it requires heat y × 80 + y × 1 × 100 = 180 y

$\therefore\ x\times 540 = 180y$ or $\frac{y}{x} = \frac{540}{180} = \frac{3}{1}$

11. (b) When length of the liquid column remains constant, then the level of liquid moves down with respect to the container, thus γ must be less than 3α.

Now we can write $V = V_{\cdot}(1+\gamma\Delta T)$

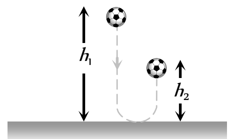
Since $V = Al = [A_{\cdot}(1+2\alpha\Delta T)]l_{\cdot} = V_{\cdot}(1+2\alpha\Delta T)$

Hence $V_{\cdot}(1+\gamma\Delta T) = V_{\cdot}(1+2\alpha\Delta T) \Rightarrow \gamma = 2\alpha$.

12. (c) By using

$\Delta\theta = 0.0023h = 0.0023\times 500 = 1.15°C \approx 1.16°C$

13. (b) According to energy conservation, change in potential energy of the ball, appears in the form of heat which raises the temperature of the ball.



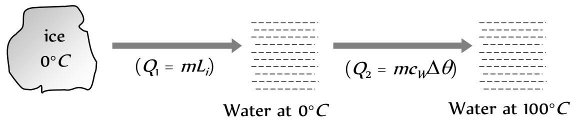
$i.e.$ $mg(h_1-h_2) = m.c.\Delta\theta$

$\Rightarrow \Delta\theta = \frac{g(h_1-h_2)}{c}$

$= \frac{10(10-5.4)}{460} = 0.1°C$

14. (c) $\theta_{mix} = \frac{\theta_w - \frac{L_i}{C_W}}{2} = \frac{100-\frac{80}{1}}{2} = 10°C$

15. (c) Ice (0°C) converts into water (100°C) in following two steps.



Total heat required

$Q = Q_1 + Q_2 = 1\times 80 + 1\times 1\times(100-0) = 180\ cal$

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