

5. Chemical Bonding

1. Ionic compounds dissolved in ____ solvents.
2. Example for polar solvents is ____.
3. The shape of BeCl_2 is ____.
4. Electro positivity is also called as ____.
5. Valence bond theory was proposed by ____.
6. The shape of NH_3 is ____.
7. 1 nanometer = ____Meter
8. Examples of Triple bond molecules ____.
9. General electronic configuration of Noble gas ____.
10. The noble element which is exception of octet rule ____.

11. Which of the following elements is electronegative? ()
a) Sodium b) Oxygen c) Magnesium d) Calcium
12. An element ${}_{11}^{23}\text{X}$ forms an ionic compound with another element Y. then the charge on the ion formed by X- is ()
a) +1 b) +2 c) -1 d) -2
13. An element 'A' forms chloride ACl_4 . The no. of electrons in the valence shell of A. ()
a) 1 b) 2 c) 3 d) 4
14. Bond angle in methane is ()
a) 170° b) 104.5 c) 109.28° d) 180°
15. Shape of ammonia molecule ()
a) Linear b) Angular c) Pyramidal d) Tetrahedral

Answers

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|------------------------------------|---|
| 1) Polar | 2) H ₂ O |
| 3) Linear | 4) Metallic Character |
| 5) Linus Pauling | 6) Pyramidal |
| 7) 10 ⁻⁹ | 8) N ₂ , C ₂ H ₂ |
| 9) ns ² np ⁶ | 10) Helium |
| 11) b | 12) a |
| 13) d | 14) c |
| 15) c | |