3. Structure of Atom

1. The electronic configuration of cu (29) ____.
2. Elliptical orbits are introduced by____.
3. The sub shell of the orbital for $l = 1$ is ____.
4. ____is a group of wavelength.
5. Splitting of spectral lines due to magnetic field is called____.
6. Splitting of spectral lines in the presence of electric field is called____.
7. Stationary orbits are introduced by ____.
8. The electronic configuration of Cr is ____.
9. The elements in which outermost orbital's are completely filled are called____.
10. Short notation of electron configuration is ____.
11. Shape of s-orbital is ( )

    a) Spherical b) Dumbbell c) Square d) Double Dumbbell

12. Shape of p-orbital is ( )

    a) Double Dumbbell b) Dumbbell c) Spherical d) Circle

13. Shape of d-orbital is ( )

    a) Spherical b) Dumbbell c) Double Dumbbell d) Square

14. In VIBGYOR which color having higher wavelength or lower frequency_____ ( )

    a) Red b) Violet c) Indigo d) Orange

15. Which of the following is not correct? ( )

    a) $2p^6$ b) $3s^1$ c) $2d^3$ d) $4f^{12}$

16. Stable elements ( )

    a) Alkali b) Alkali Metals c) Inert d) None

17. The sub shell of the orbital for $l = 1$ is ( )

    a) s b) p c) d d) f

18. Short rotation of electron configuration is ( )

a) nl^x b) nl^n c) $l n^x$ d) $x n^l$

19. Splitting of spectral lines in the presence of electric field is called ( )

a) Zeeman Effect b) Stark Effect c) Dispersion d) Limen Effect

20. What is 'n' value for L shell ( )

a) 1 b) 2 c) 3 d) 4

Answers

1) $[(Ar)4s^1 3d^{10}]$ 2) Sommerfeld

3) p 4) Spectrum

5) Zeeman Effect 6) Stark Effect

7) Neil's Bohr 8) $[Ar]4s^1 3d^5$

9) Inert Gas 10) nl^x

11) a 12) b

13) c 14) a

15) c 16) c

17) b 18) a

19) b 20) b